

Hydrocarbon Chapter 24

 \bigcap

Aliphatic Hydrocarbon

Organic compounds composed of only carbon and hydrogen are called hydrocarbons. Hydrocarbons are two types

- (1) *Aliphatic Hydrocarbon (Alkanes, Alkenes and Alkynes).*
- (2) *Aromatic Hydrocarbon (Arenes)*
- (1) **Sources of aliphatic hydrocarbon**

Mineral oil or crude oil, petroleum [Petra \rightarrow rock; oleum \rightarrow oil] is the dark colour oily liquid with offensive odour found at various depths in many regions below the surface of the earth. It is generally found under the rocks of earth's crust and often floats over salted water.

(2) **Composition**

(i) *Alkanes* **:** found 30 to 70% contain upto 40 carbon atom. Alkanes are mostly straight chain but some are branched chain isomers.

(ii) *Cycloalkanes* **:** Found 16 to 64% cycloalkanes present in petroleum are; cyclohexane, methyl cyclopentane etc. cycloalkanes rich oil is called asphaltic oil.

(iii) *Aromatic hydrocarbon* **:** found 8 to 15% compound present in petroleum are; Benzene, Toluene, Xylene, Naphthalene etc.

(iv) *Sulphur, nitrogen and oxygen compound* **:** Sulphur compound present to the extent of 6% include mercaptans [R-SH] and sulphides [R-S-R]. The unpleasant smell of petroleum is due to sulphur compounds. Nitrogenous compounds are pyridines, quinolines and pyrroles. Oxygen compounds present in petroleum are. Alcohols, Phenols and resins. Compounds like chlorophyll, haemin are also present in it.

(v) *Natural gas* **:** It is a mixture of Methane (80%), Ethane (13%), Propane (3%), Butane (1%), Vapours of low boiling pentanes and hexanes (0.5%) and Nitrogen (1.3%). L.P.G. Contain butanes and

pentanes and used as cooking gas. It is highly inflammable. This contain, methane, nitrogen and ethane.

(vi) *C.N.G.* **:** When natural gas compressed at very high pressure is called compressed natural gas (CNG). Natural gas has octane rating of 130 it consists, mainly of methane and may contain, small amount of ethane and propane.

(3) **Theories of origin of petroleum :** Theories must explain the following characteristics associated with petroleum,

Its association with brine (sodium chloride solution). The presence of nitrogen and sulphur compounds in it. The presence of chlorophyll and haemin in it. Its optically active nature. Three important theories are as follows.

(i) *Mendeleeff's carbide theory or inorganic theory*

- (ii) *Engler's theory or organic theory*
- (iii) *Modern theory*

(4) **Mining of petroleum :** Petroleum deposits occurs at varying depth at different places ranging from 500 to 15000 feet. This is brought to the surface by artificial drilling.

(5) **Petroleum refining :** Separation of useful fractions by fractional distillation is called petroleum refining.

Fraction			
	Boiling range (°C)	Approximate composition	Uses
Uncondensed gases	Upto room temperature	$C_1 - C_4$	Fuel gases: refrigerants; production of carbon black,
			hydrogen; synthesis of organic chemicals.
Crude naphtha on refractionation	$30 - 150^{\circ}$	$C_5 - C_{10}$	
yields,			
(i) Petroleum ether	$30 - 70^{\circ}$	$C_5 - C_6$	Solvent
(ii) Petrol or gasoline	$70 - 120^{\circ}$	$C_6 - C_8$	Motor fuel; drycleaning; petrol gas.
(iii) Benzene derivatives	$120 - 150^{\circ}$	$C_8 - C_{10}$	Solvent; drycleaning
Kerosene oil	$150 - 250^{\circ}$	$C_{11} - C_{16}$	Fuel; illuminant; oil gas
Heavy oil	$250 - 400^{\circ}$	$C_{15} - C_{18}$	As fuel for diesel engines; converted to gasoline by
			cracking.
Refractionation gives,			
(i) Gas oil, (ii) Fuel oil,			
(iii) Diesel oil			
Residual oil on fractionation by	Above 400°	$C_{17} - C_{40}$	
vacuum distillation gives,			
(i) Lubricating oil		$C_{17} - C_{20}$	Lubrication
(ii) Paraffin wax		$C_{20} - C_{30}$	Candles; boot polish; wax paper; etc
(iii) Vaseline		$C_{20} - C_{30}$	Toilets; ointments; lubrication.
(iv) Pitch		$C_{30} - C_{40}$	Paints, road surfacing
Petroleum coke			As fuel.
(on redistilling tar)			

Table : 24.1

(6) **Purification**

(i) *Treatment with concentrated sulphuric acid* **:** The gasoline or kerosene oil fraction is shaken with sulphuric acid to remove aromatic compounds like thiophene and other sulphur compound with impart offensive odour to gasoline and kerosene and also make them corrosive.

(ii) *Doctor sweetening process* **:**

 $2RSH$ + Na_2PbO_2 + $S \rightarrow$ $RSSR$ + PbS + 2 $NaOH$ CH_3 CH_3 CH_3

(iii) *Treatment with adsorbents* : Various fractions are passed over adsorbents like alumina, silica or clay etc, when the undesirable compounds get adsorbed.

(7) **Artificial method for manufacture of Petrol or gasoline**

(i) Cracking, (ii) Synthesis

(i) *Cracking* **:** It is a process in which high boiling fractions consisting of higher hydrocarbons are heated strongly to decompose them into lower hydrocarbons with low boiling points. Cracking is carried out in two different ways.

(a) *Liquid phase cracking :* In this process, the heavy oil or residual oil is cracked at a high temperature $(475 - 530^{\circ}C)$ under high pressure (7 to 70 atmospheric pressure). The high pressure keeps the reaction product in liquid state. The conversion is approximately 70% and the resulting petrol has the octane number in the range 65 to 70.

The cracking can be done in presence of some catalysts like silica, zinc oxide, titanium oxide, ferric oxide and alumina. The yields of petrol are generally high when catalyst is used.

(b) *Vapour phase cracking :* In this process, kerosene oil or gas oil is cracked in vapour phase. The temperature is kept $600 - 800^{\circ}$ C and the pressure is about 3.5 to 10.5 atmospheres. The cracking is facilitated by use of a suitable catalyst. The yields are about 70%.

(ii) *Synthesis* **:** Two methods are applicable for synthesis.

(a) *Bergius process :* This method was invented by Bergius in Germany during first world war.

\n
$$
\text{Coal } + H_2 \xrightarrow{\text{Fe}_2\text{O}_3} \text{Mix. Of hydrocarbons or crude oil}
$$
\n

\n\n
$$
\text{Which combine with lead to the following equation: } -\frac{F\phi_2 G_3}{450 - 500^\circ C}
$$
\n

\n\n
$$
Pb + Br - CH_2 - Ct
$$
\n

(b) *Fischer- tropsch process :* The overall yield of this method is slightly higher than Bergius process.

$$
H_2O + C \xrightarrow{} 1200^{\circ}C \xrightarrow{} CO + H_2
$$
Watergas

 $XCO+ yH_2 \longrightarrow \text{Mix. Of hydrogen} + H_2O$.

The best catalyst for this process is a mixture of cobalt (100 parts), thoria, (5 parts), magnesia (8 parts) and kieselguhr (200 parts).

Characteristics of hydrocarbons

(1) **Knocking :** The metallic sound during working of an internal combustion engine is termed as knocking.

"*The greater the compression greater will be efficiency of engine*." The fuel which has minimum knocking property is always preferred.

The tendency to knock falls off in the following order **:** Straight chain alkanes > branched chain alkanes > olefins > cyclo alkanes > aromatic hydrocarbons.

(2) **Octane number :** It is used for measuring the knocking character of fuel used in petrol engine. The octane number of a given sample may be defined as the percentage by volume of iso-octane present in a mixture of iso-octane and *n*-heptane which has the same knocking performance as the fuel itself.

$$
CH_3 - CH_2 - CH_2 - CH_2 - CH_2 - CH_2 - CH_3
$$

\n*n*-heptane; octane no. = 0
\n CH_3 CH_3
\n $CH_3 - C - CH_2 - C - CH_3$; Octane no. = 100

$$
CH_3 - C - CH_2 - C - CH_3
$$
; Octane no. = 100
CH₃

2, 2, 4-Trimethyl pentane or Iso-octane.

For example : a given sample has the knocking performance equivalent to a mixture containing 60% iso-octane and 40% heptane. The octane number of the gasoline is, therefore, 60.

Presence of following types of compounds increases the octane number of gasoline.

(i) In case of straight chain hydrocarbons octane number decreases with increase in the length of the chain.

(ii) Branching of chain increases the value of octane number

(iii) Introduction of double bond or triple bond increases the value of octane number.

(iv) Cyclic alkanes have relatively higher value of octane number.

(v) The octane number of aromatic hydrocarbons are exceptionally high

(vi) By adding gasoline additives (eg TEL)

(3) **Antiknock compounds :** To reduce the knocking property or to improve the octane number of a fuel certain chemicals are added to it. These are called *antiknock compounds*. One such compound, which is extensively used, is tetraethyl lead (TEL). TEL is used in the form of following mixture,

TEL = 63% , Ethylene bromide = 26% , Ethylene chloride = 9% and a dye $= 2\%$.

However, there is a disadvantage that the lead is deposited in the engine. To remove the free lead, the ethylene halides are added which combine with lead to form volatile lead halides.

$$
Pb + Br - CH_2 - CH_2 - Br \rightarrow PbBr_2 + CH_2 = CH_2
$$

Ethylenebromide *Volatile* Ethylene

 $H_2O + C \longrightarrow \frac{1200^{\circ}C}{C} + H_2$ manganese carbonyl (called as AK-33-X) is used in developed However, use of TEL in petrol is facing a serious problem of Lead pollution, to avoid this a new compound cyclopenta dienyl countries as antiknocking compound.

> (4) **Other methods of improving octane number of hydrocarbon.**

(i) *Isomerisation* [Reforming] **:** By passing an alkane over A/Cl_3 at $200^\circ C$.

$$
CH_3CH_2CH_2CH_2CH_3 \xrightarrow{\text{AlC4}_3 \qquad |}_{200^{\circ}C} CH_3CH_2CH_3 \xrightarrow{\text{AlC4}_3 \qquad |}_{1 \text{Sopentane}}
$$
\n(Octanenumber = 62) (Octanenumber = 90)

(ii) *Alkylation* **:**

CH_3	CH_3	CH_3	CH_3	
1				
1				
1				
2				
3				
4				
5				
6				
6				
7				
8				
9				
10				
11				
12				
13				
14				
15				
16				
17				
18				

(iii) *Aromatisation* **:**

$$
CH_3(CH_2)_5 CH_3 \xrightarrow{Pt1 A_2O_3
$$

\n $Heptane$
\n CH_3
\n $+$ 4 H_2
\n $+$ 4 $+$

The octane no. of petrol can thus be improved.

 By increasing the proportion of branched chain or cyclic alkanes.

 By addition of aromatic hydrocarbons Benzene, Toluene and Xylene (BTX).

By addition of methanol or ethanol.

 \bullet By additon of tetraethyl lead $(C_2H_5)_4$ *Pb*

(5) **Cetane number :** It is used for grading the diesel oils.

$$
CH_3 - (CH_2)_{14} - CH_3
$$
 Cetane \rightarrow cetane no. = 100 C_nH

The Methyl *naphthalenes* a diesel oil is the percentage of cetane (hexadecane) by volume in a mixture of cetane and α -methyl naphthalene which has the same ignition property as the fuel oil under consideration.

(6) **Flash point :** The lowest temperature at which an oil gives sufficient vapours to form an explosive mixture with air is referred to as flash point of the oil.

The flash point in India is fixed at $44^{\circ}C$, in France it is fixed at 35^o*C*, and in England at 22.8^o*C*. The flash point of an oil is usually determined by means of "*Abel's apparatus*".

Chemists have prepared some hydrocarbons with octane number even less than zero (e.g., *n*-nonane has octane number – 45) as well as hydrocarbon with octane number greater than 100 (e.g., 2, 2, 3 trimethyl-butane. has octane number of 124).

(7) **Petrochemicals :** All such chemicals which are derived from petroleum or natural gas called petrochemicals. Some chemicals which are obtained from petroleum are summarised in table :

Table : 24.2

"Alkanes are saturated hydrocarbon containing only carboncarbon single bond in their molecules."

Alkanes are less reactive so called paraffins; because under normal conditions alkanes do not react with acids, bases, oxidising agents and reducing agent.

General formula : C_nH_{2n+2}

Examples are CH_A , $C₂H₆$, $C₃H₈$,

(1) **General Methods of preparation**

)⁴ *Pb* (i) *By catalytic hydrogenation of alkenes and alkynes* (Sabatie and sanderen's reaction)

$$
C_nH_{2n^+}H_2 \xrightarrow{\text{N} \atop \text{hleat}} C_nH_{2n+2} \, ; \, \, C_nH_{2n-2} + 2H_2 \xrightarrow{\text{N} \atop \text{Alkene}} C_nH_{2n+2}
$$

 \Box Methane is not prepared by this method

(ii) *Birch reduction* **:**

$$
R-CH = CH_2 \xrightarrow{1. \text{ NaI} \text{ NH}_3} R-CH_2-CH_3
$$

(iii) *From alkyl halide*

(a) By reduction :
$$
RX + H_2 \xrightarrow{ZnI HCI} RH + HX
$$

(b) With hydrogen in presence of
$$
pt/pd
$$
:
 $RX + H_2 \xrightarrow{PdorPt} RH + HX$

(c) With
$$
HI
$$
 in presence of Red phosphorus :

 $RBr+2Hl \longrightarrow RH+HBr+1_{2}$
Purposeof Red Pis to remove I_{2} in the form of Pl_{3}

(iv) *By Zn-Cu couple* **:**

$$
2CH_3CH_2OH + Zn \xrightarrow{Cn} \xrightarrow{Cu} (CH_3CH_2O)_2 Zn + 2H
$$

2*incutboxide*

$$
RX + 2H \longrightarrow RH + HX
$$

(v) *Wurtz reaction* :

$$
R\overline{X} + \overline{2}\overline{X} = \overline{2\Delta a} + \overline{X}^1R' \xrightarrow{Dryether} R - R + 2NaX
$$

Alkyl halide

□ *R*−*Br* or *RI* preferred in this reaction. The net result in this reaction is the formation of even no. of carbon atoms in molecules.

(vi) *Frankland's reaction* **:**

$$
2RX + Zn \longrightarrow R - R + ZnX_2
$$

(vii) *Corey-house synthesis*

$$
CH_3 - CH_2 - Cl \xrightarrow{1.Li} (CH_3 - CH_2)_2
$$
LiCu $\xrightarrow{CH_3 - CH_2 - Cl} (CH_3 - CH_2 - CH_2 - CH_3$
 $CH_3 - CH_2 - CH_2 - CH_3$

□ Reaction is suitable for odd number of Alkanes.

(viii) *From Grignard reagent*

(a) *By action of acidic* '*H*' :

$RMgX$	$+ HOH$	$RH + Mg(OH)X$	$CH_3COCH_3 + 4H$	$\frac{2n-Hg}{HCl}$
Alkyl magnesium Water	Alkane	ACH_3	$ACBCH_3 + 4H$	HCH_3
halide	(Propanonel)			

(b) *By reaction with alkyl halide* :

$$
R - X + RMgX \longrightarrow R - R + MgX
$$

(ix) *From carboxylic acids*

(a) *Laboratory method* [Decarboxylation reaction or Duma reaction]

$$
R\text{ COONa+} \text{NaOH} \xrightarrow{\text{heat}} R \rightarrow R \rightarrow H + \text{Na}_2\text{CO}_3
$$

NaOH and *CaO* is in the ratio of 3 : 1. (Sodalime)

(b) *Kolbe's synthesis* :

$$
R-C-O^T N a^+ \xrightarrow{\text{Electrolysis}} R-C-O^T + N a^+
$$

At anode [Oxidation] :

$$
2R-C-O-2\epsilon \longrightarrow 2R-C-\stackrel{\circ}{O} \longrightarrow 2\stackrel{\bullet}{R}+2CO_2
$$

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 $2R \longrightarrow R - R$ (alkane)

At cathode [Reduction] :

$$
2Na^{+} + 2e^{-} \longrightarrow 2Na^{-} \longrightarrow 2NaOH + H_{2}(\uparrow)
$$

 \square Both ionic and free radical mechanism are involved in this reaction.

(c) *Reduction of carboxylic acid* :

$$
CH_3COOH + 6HI \xrightarrow{\text{Reduction}} CH_3CH_3 + 2H_2O + 3I_2
$$
\n
$$
\xrightarrow{\text{(ii) Density: Alkanes are lighter than water.}} CH_3CO + 3I_2
$$
\n
$$
\xrightarrow{\text{(iii) } Density: Alkanes are lighter than water.}
$$
\n
$$
\xrightarrow{\text{(iv) } Callibility: Insoluble in water.} CH_3CH_3 + 2H_2O + 3I_2
$$

(x) *By reduction of alcohols, aldehyde, ketones or acid derivatives*

$$
CH_3OH + 2Hl \xrightarrow{\text{Red } P} CH_4 + H_2O + I_2
$$
 (iv) **Boiling points and** (Methyl alcohol)
\n(Methyl alcohol)
\n(Methyl alcohol)
\n $150^{\circ}C$

2 2 Ethane 2 6 ¹⁵⁰ Red (Ethanal) Acetaldehyde *CH* ³*CHO* ⁴*HI ^C ^H ^H ^O* ²*^I ^C P ^o* boiling points. Molecular mass No.of branches ¹

$$
CH_3COCH_3 + 4 H I \xrightarrow{\text{Red}P} CH_3CH_2CH_3 + H_2O + 2I_2
$$
\n
\n_{Acotone} (Propano)
\n_{(Propanon})
\n
$$
CH_3COCH_3 + 4 H I \xrightarrow{\text{Red}P} CH_3CH_2CH_3 + H_2O + 2I_2
$$
\n
\n
$$
H.P.(K) \xrightarrow{\text{85.9}} 138 \xrightarrow{\text{138}} 143.3
$$

O

 $\frac{\text{Red}P}{200^{\circ}C}CH_3 - CH_3 + H_2O + HCl + 3I_2$ symmetrical strategy example of $200^{\circ}C$ Acetyl chloride $200\degree C$ Ethane
(Ethanoyl chloride) C_3 – C– C/+ 6 H/ $\frac{\text{Re}aP}{\text{Im}a}$ \rightarrow CH₃ – CH₃ + H₂O+ HC/ + 3/₂ symmetrical structure and || $CH_3 - C - C$ + 6 $HI - \frac{C+C}{200^{\circ}C} \rightarrow CH_3 - CH_3 + H_2O + HCl + 3I_2$ symmetrical structure and result $-C$ - C/+ 6 HI $\frac{\text{Red}P}{\text{QO}Q}$ \rightarrow CH₃ - CH₃+ H₂O+ HCI+3I₂ symmetrical structure and re

$$
CH_3-C-NH_2+6H1 \xrightarrow{\text{Red }P} CH_3-CH_3+H_2O+NH_3+3I_2
$$

\n
$$
CH_3-C-NH_2+6H1 \xrightarrow{Red P} CH_3-CH_3+H_2O+NH_3+3I_2
$$

\n
$$
C\ C
$$

□ Aldehyde and ketones when reduced with amalgamated zinc and conc. *HCl* also yield alkanes.

Clemmenson reduction :

$$
CH_3CHO + 4H \xrightarrow{Zn-Hg} CH_3 - CH_3 + H_2O
$$
\n
$$
\xrightarrow{\text{Actualdehyde}} \qquad \qquad HCl \qquad \qquad \xrightarrow{\text{Ethane}}
$$

$$
CH_3COCH_3 + 4H \xrightarrow{Zn-Hg} CH_3CH_2CH_3 + H_2O
$$

Accept
(Propano)
(Propano)
(Propano)

 $R - \chi + R M g \chi$
on heating on heating \Box Aldehydes and ketones ($>C = O$) can be reduced to hydrocarbon in presence of excess of hydrazine and sodium alkoxide on heating.

Wolff-kishner reduction :

$$
R\n\nC = O \xrightarrow{-H_2NNH_2} R\n\nR\n\n180oC-N_2\n\nR\n\n
$$
C = NNH_2 \xrightarrow{C_2H_5ONa} R\n\n180oC-N_2\n
$$
$$

(xi) *Hydroboration of alkenes*

(a) *On treatment with acetic acid*

$$
R-CH = CH_2 \xrightarrow{\qquad B_2H_6} (R-CH_2-CH_2)_3B \xrightarrow{\qquad CH_3COOH} (R-CH_2-CH_3)B \xrightarrow{\qquad CH_3COOH} (R-CH_2-CH_3)
$$
\n
$$
R-CH_2-CH_3 \xrightarrow{\qquad \text{Rikene}} R-CH_2-CH_3
$$
\n
$$
R-CH_2-H_3 \xrightarrow{\qquad \text{Rikane}}
$$

(b) *Coupling of alkyl boranes by means of silver nitrate*

$$
6[R-CH = CH_2] \xrightarrow{2B_2H_6} 2R-CH_2-CH_2-I_3B \xrightarrow{AgNO_3 25^oC} NaoH
$$

3[RCH_2CH_2-CH_2CH_2H_3]

O (2) **Physical Properties**

(i) *Physical state* **:** Alkanes are colourless, odourless and tasteless.

(iii) *Solubility* **:** Insoluble in water, soluble in organic solvents, solubility ∞ $\frac{1}{\text{Molecular mass}}$

 \rightarrow 2HI \rightarrow $CH_4 + H_2O + I_2$
(iv) **Boiling points and Melting points :** Melting points and

 \Box Melting points of even $>$ Odd no. of carbon atoms, this is because, the alkanes with even number of carbon atoms have more symmetrical structure and result in closer packing in the crystal structure as compared to alkanes with odd number of carbon atoms.

(i) *Substitution reactions of Alkanes*

 $CH_3CHO_4 + H \xrightarrow{Zn-Hg} CH_3 - CH_3 + H_2O$ (a) $Halogenation: R-H+X-X \longrightarrow R-X+HX$ The reactivity of halogen is : $F_2 > Cl_2 > Br_2 > l_2$

 \Box Fluorine can react in dark Cl_2 , Br_2 require light energy. I_2 doesnot show any reaction at room temperature, but on heating it shows iodination.

 \Box Iodination of methane is done in presence of oxidising agent such as $HNO₃$ / $HIO₃$ / HgO which neutralises HI .

Chlorination of methane :

$$
CH_4 + 2Cl - Cl \xrightarrow{-2HCl} CH_2 - Cl_2 \xrightarrow{-HCl} CH_2
$$

\n $CH_4 + 2Cl - Cl \xrightarrow{-2HCl} CH_2 - Cl_2 \xrightarrow{-HCl} CH_2$
\n $CH_4 + 2Cl - Cl \xrightarrow{-HCl} CH_2$
\n $CH_4 + 2Cl - Cl \xrightarrow{u \times light, Cl_2}$
\n $QHCl$
\n $QHCl$ <

(ii) *Reaction based on free radical mechanism*

(a) Nitration:
$$
R - H + HONO_2 \xrightarrow{High} R - NO_2 + H_2O
$$

\nAlkané
\nAltonale
\nNitrolkané
\nNitrolkané

Nitrating mixture : (i) $(Con HNO₃ + Con H₂SO₄)$ at $\left.\begin{array}{cc} H_2C \\ H_2 \end{array}\right)$ *C* 250° C

(ii) $(HNO₃$ vapourat 400[°] – 500[°]C).

(b) *Sulphonation* **:** Free radical mechanism $R - H + HOSO₃H \xrightarrow{\text{SO}_3} R - SO₃H + H₂O$ n -Heptgne **Ctan** un reaction

 Lower alkanes particularly methane, ethane, do not give this reaction.

(iii) *Oxidation*

(a) *Complete Oxidation or combustion* :

$$
C_nH_{2n+2}+\left(\frac{3n+1}{2}\right)O_2\longrightarrow nCO_2+(n+1)H_2O+Q
$$

$$
R-CH_2\underset{\sim}{\underbrace{\uparrow}}H-\frac{CH_2O_3+OH}{CH_2}+H
$$

 \Box This is exothermic reaction.

(b) *Incomplete combustion or oxidation*

$$
2CH_4 + 3O_2 \xrightarrow{\text{Burn}} 2CO + 4H_2O
$$

$$
CH_4 + O_2 \longrightarrow C + 2H_2O
$$

(c) *Catalytic Oxidation* : $CH_4 + [O] \xrightarrow{CU-100} CH_3OH$ $C_4 + [O] \frac{Cu - tube}{100 atm/200°C} CH_3OH$ (viii) HCN formation:

This is the industrial method for the manufacture of methyl alcohol.

 \Box Higher alkanes are oxidised to fatty acids in presence of manganese stearate.

$$
CH_3(CH_2)_nCH_3 \xrightarrow[100-160^\circ C]{} CH_3(CH_2)_n COOH
$$

\n $CH_3-CH_2-CH_3+SO_2$

(d) *Chemical oxidation* :

$$
(CH_3)_3 CH \xrightarrow{K M n O_4} (CH_3)_3. C. OH
$$

Isobutane\nTetiary buyl alcohol\n

(iv) *Thermal decomposition or cracking or pyrolysis or fragmentation*

CH_4	1000^0C	$0.2H_2$	Individual members of alkanes	
C_2H_6	500^0C	$CH_2 = CH_2 + H_2$		
C_2H_6	500^0C	$2H_2 = CH_2 + H_2$		
$Chane C_2O_3 + A_2O_3$	$Ethylene$	$Etlylene$		
C_3H_8	$0.2H_4 + CH_4$ or $C_3H_6 + H_2$	U		
$0.3H_8$	$0.2H_4 + CH_4$	$0.3H_6 + H_2$	U	U

*I*₂ □ This reaction is of great importance to petroleum industry. (v) *Isomerisation :*

 $CHCl_3 \xrightarrow{CHCl_4} CCl_4$ (vi) **Aromatisation :** butane

(a) *Reaction* with $CH₂ N₂$ (Diazo methane) : *Hexane*

$$
R-CH_2-H+CH_2N_2 \xrightarrow{hv} R-CH_2-CH_2-H
$$

(b) *Reaction* with **CHCl₃** / **NaOH** :

$$
O_2 \longrightarrow nCO_2 + (n+1)H_2O + Q
$$

$$
R-CH_2 \longrightarrow H \xrightarrow{CHCl_3 \cup H^-} R-CH_2-CHCl_2
$$

$$
CC_2 \longrightarrow C+Cl_2
$$

(c) *Reaction* with $CH_2 = C$: *O* $Z_2 = C$:

$$
R-CH_2-H\frac{CH_2=CA}{:CH_2-CO}R-CH_2-CH_3
$$

(viii) *HCN formation* **:**

$$
2CH_4 \xrightarrow{N_2 \text{ 1 electric arc}} 2HCN + 3H_2 \text{ or}
$$

$$
CH_4 + NH_3 \xrightarrow[700^{\circ}C]{} HCN + 3H_2
$$

(ix) *Chloro sulphonation*/*Reaction with* $SO_2 + Cl_2$

$$
CH_3-CH_2-CH_3+SO_2+Cl_2\longrightarrow
$$

$$
CH_3-CH_2-CH_2SO_2Cl+HCl
$$

 \Box This is used in the commercial formation of detergent.

(x) Action of steam :
$$
CH_4 + H_2O \xrightarrow{Ni/Al_2O_3} CO + 3H_2
$$

(1) **Methane :** Known as *marsh gas*.

 $\frac{500^{\circ}C}{2}$ $CH_2 = CH_2 + H_2$ (i) *Industrial method of preparation* **:** Mathane gas is obtained on a large scale from natural gas by liquefaction. It can also be obtained by the application of following methods,

(a) *From carbon monoxide* : A mixture of carbonmonoxide and hydrogen is passed over a catalyst containing nickel and carbon at 250 \degree *C* when methane is formed.

$$
CO + 3H_2 \xrightarrow{\text{Ni} + \text{C}} CH_4 + H_2O
$$
 made. It is a multistep conversion.

(b) *Bacterial decomposition of cellulose material present in sewage water* : This method is being used in England for production of methane.

$$
(C_6H_{10}O_5)_n + nH_2O \longrightarrow 3nCH_4 + 3nCO_2
$$

\n
$$
\xrightarrow{\text{[0]} C_6H_{10}O_5} CH_3^C, COOH \longrightarrow C
$$

\n
$$
\xrightarrow{\text{Cellululose}}
$$

\n
$$
\xrightarrow{\text{Cellululose}}
$$

(c) *Synthesis* : \Box By striking an electric arc between carbon electrodes in an atmosphere of hydrogen at 1200°C, methane is formed.

$$
C+2H_2 \xrightarrow{1200^{\circ}C} CH_4
$$

By passing a mixture of hydrogen sulphide and carbon disulphide vapour through red hot copper, methane is formed.

 $CS_2 + 2H_2S + 8Cu$ *High temperature* $\rightarrow CH_4 + 4Cu_2S$ contain double bond. These are also known as

(ii) *Physical properties*

(a) It is a colourless, odourless, tasteless and non-poisonous gas.

(b) It is lighter than air. Its density at NTP is 0.71 g/L.

(c) It is slightly soluble in water but is fairly soluble in ether, alcohol and acetone.

(d) Its melting point is $-182.5^{\circ}C$ and boiling point is \sqrt{a} $-161.5^{\circ}C$.

(iii) *Uses*

(a) In the manufacture of compounds like methyl alcohol, formaldehyde, methyl chloride, chloroform, carbon tetrachloride, etc.

(b) In the manufacture of hydrogen, used for making ammonia.

(c) In the preparation of carbon black which is used for making printing ink, black paints and as a filler in rubber vulcanisation.

(d) As a fuel and illuminant.

(2) **Ethane**

(i) *Methods of preparation*

(a) *Laboratory method of preparation :*

 C_2H_5OH Ethane

(b) *Industrial method of preparation :*

$$
CH_2 = CH_2 + H_2 \xrightarrow{\text{M} \atop \text{Ethylene}} CH_3 - CH_3
$$
\n
$$
\xrightarrow{\text{Ethylene}} \xrightarrow{\text{C} \atop \text{I} \atop \text{C} \atop \text{I}}
$$
\n
$$
\xrightarrow{\text{C} \atop \text{C} \atop \text{C}
$$

(ii) *Physical properties*

(a) It is a colourless, odourless, tasteless and non-poisonous

(b) It is very slightly soluble in water but fairly soluble in alcohol, acetone, ether, etc.

(c) Its density at NTP is 1.34 *g/L*

(d) It boils at -89° C. Its melting point is -172° C.

(iii) *Uses*

gas.

(a) As a fuel. (b) For making hexachloroethane which is an artificial camphor.

(3) **Interconversion of Alkanes**

Ascent of alkane series,

(i) *Methane to ethane* **:**

$$
CH_4 \xrightarrow{Cl_2} CH_3 CH \xrightarrow{\text{Wurtzreaction}} CH_3 - CH_3
$$
 formed by dehalogenation of it.
\nMethod 1. Determine the output of the output.

(ii) *Butane from ethane* **:**

$$
C_2H_6 \xrightarrow{C_2} C_2H_5Cl \xrightarrow{\text{Wurtz reaction}} C_2H_5 - C_2H_5
$$

Ethane
(excess) UV Ethyl chloride Heat with *Nai*n ether

Descent of alkane series : Use of decarboxylation reaction is

<i>Ethane to methane</i>		
$C_2H_6 \xrightarrow{C_2} C_2H_5Cl \xrightarrow{Aq. KOH} C_2H_5OH \xrightarrow{C_2} H_5OH \xrightarrow{C_2} H_5OH$	CH_3CHO	
Éthynalcohol	Actaldehyde	
(excess)	$CH_3COOH \xrightarrow{NaOH} CH_3 COONa \xrightarrow{NaOH/CoO} CH_4$	
$Cl \rightarrow CH_3COOH \xrightarrow{NaOH} CH_3 COONa \xrightarrow{NaOH/CoO} CH_4$		
Adeclicacid	heat	Methane
Higher $\xrightarrow{C_2} \rightarrow$ Alkyl $\xrightarrow{Aq} \rightarrow$ Alcohol $\xrightarrow{[O]} \rightarrow$ Aldehyde $\xrightarrow{[O]} \rightarrow$		
alkane \xrightarrow{UV} halide \xrightarrow{KOH}	Alcohol $\xrightarrow{[O]} \rightarrow$ Aldehyde $\xrightarrow{[O]} \rightarrow$	
Acid $\xrightarrow{NaOH} \rightarrow$ Sodium salt of $\xrightarrow{NaOH/CoO} \rightarrow$ Lower alkane		

Alkenes

These are the acyclic hydrocarbon in which carbon-carbon contain double bond. These are also known as olefins, because lower alkene react with halogens to form oily substances. General formula is C_nH_{2n} . Examples, C_2H_4 , C_3H_6 , C_4H_8 .

(1) **Preparation methods**

(i) *From Alkynes* **:**

$$
R-C \equiv C - R + H_2 \xrightarrow{\text{Lindlar's Catalogst}} R - C = C - R
$$

\n
$$
R \rightarrow C = C - R
$$

\n
$$
H
$$

\n
$$
H
$$

\n
$$
H
$$

\n
$$
H
$$

a Poison's catalyst such as $BASO₄$, CaCO₃ are used to stop the reaction after the formation of alkene.

(ii) From mono halides :
\n
$$
H H
$$

\n $|$
\n $R-C-C-H+Alc KOH$
\n $-HX \rightarrow R-C= C-H$

 $C_2 H_5 I + 2H \xrightarrow{Zn-Cu couple} C_2 H_6 + H$
 $C_2 H_5OH$ \xrightarrow{Ethane} $C_2 H_6 + H$ is formed in majority because of its *H X H* Alkene If we use alc. *NaOH* in place of *KOH* then trans product is formed in majority because of its stability. According to saytzeff

H

H

|

^o rule. (iii) *From dihalides* (a) *From Gem dihalides ^R – CH ⁼ CH – ^R* ² *ZnX*² *^R – CH + + CH – ^R X Zn X X Zn X*

 \Box If we take two different types of gemdihalides then we get three different types of alkenes .

 \Box Above reaction is used in the formation of symmetrical alkenes only.

(b) *From vicinal dihalides :*

$$
\begin{array}{ccc}\nH & H & H \\
\mid & \mid & \mid \\
R-C-C-H+Zndust \xrightarrow{\Lambda} & \mid & \mid \\
\mid & \mid & \downarrow \\
X & X & X\n\end{array}
$$

Alkene is not formed from 1, 3 dihalides. Cycloalkanes are

$$
\begin{array}{ccc}\nCH_2-CH_2-CH_2\longrightarrow\longrightarrow & CH_2 & +ZnX_2 & \text{ (iv)}\\
\downarrow & & \downarrow & & \downarrow\\ X & X & H_2C \longrightarrow CH_2 & & \text{ (v)}\\ \n\end{array}
$$

(iv) *By action of NaI on vicinal dihalide* **:**

Br Br
 $\begin{array}{c} | \\ | \\ \hline \end{array}$ $\begin{array}{c} | \\ | \\ \hline \end{array}$
 $\begin{array}{c} | \\ \hline \end{array}$ $\begin{array}{c} | \\ \hline \end{array}$ $\begin{array}{c} | \\ \hline \end{array}$ $C-C\leftarrow \frac{N\epsilon I}{C}$ \rightarrow $\left\langle C-C\leftarrow \longrightarrow \right\rangle$ $\left\langle C=C\leftarrow \right\rangle$ (vi) The 1 *Br* | | $-C\leftarrow \longrightarrow \rightarrow C-C\leftarrow \longrightarrow \rightarrow C=$ *I I* | | vic dihalide $\frac{a_0 e^{i\theta}}{i}$ unstable $\frac{a_1 e^{i\theta}}{i}$ alkene $CH_3CH_2OH \xrightarrow{H_2SO_4 \text{ or } H_3PO_4} CH_2 = CH_2 + H_2O$ (vii) These are lighter than water.
Ethyl alcohol 343 K Ethene (vi) *Kolbe's reaction* **:** \vert $+2H_2O \longrightarrow \vert$ \downarrow $+2CO_2 + H_2 + 2KOH$ (3) Chemical propertie *CH COOK* 2 *CH COOK* 2 **Ethene** 2 $2H_2O \frac{\text{Electrolyss}}{2}$ 2 $|$ $+ 2CO_2 + H_2 + 2KOH$ dipole moments are higher than those of a
(3) Chemical properties Potassiumsuccinate

(vii) *From esters* [Pyrolysis of ester] **:**

$$
CH_3-CO-O
$$

\n
$$
- - + - - +
$$

\n
$$
CH_2-CH_2
$$

\n
$$
CH_3-COOH
$$

\n
$$
CH_3-COOH
$$

\n
$$
CH_3-COOH
$$

\n
$$
CH_2=CH_2
$$

\n
$$
CH_2=CH_2
$$

(viii) **P***yrolysis of quaternary ammonium compounds* **:**

$$
(C_2H_5)_4 \stackrel{\dagger}{\text{Notav} \text{Mathylamonion}} \begin{array}{c}\n (C_2H_5)_3 \stackrel{\dagger}{\text{Univ}} \begin{array}{c}\n (C_2H_5)_4 \stackrel{\dagger}{\text{Univ}} \begin{array}{c}\n (C_2H_5)_5 \stackrel{\dagger}{\text{Univ}} \\
 (D_5H_5)_6 \stackrel{\dagger}{\text{Univ}}\n \end{array}\n \end{array}
$$

(ix) *Action of copper alkyl on vinyl chloride* **:**

$$
H_2C = CHCl \xrightarrow{CuR_2} H_2C = CHR
$$

\n
$$
R - C = C - R + H_2 \xrightarrow{1}
$$

\n
$$
R - C = C - R + H_2 \xrightarrow{1}
$$

(x) *By Grignard reagents* **:**

$$
Mg\left(\begin{array}{c}R\\+X-CH=CH_2\longrightarrow MgX_2+R-CH=CH_2\end{array}\right)
$$
 (iii) *Reduction of alkene via hydro*
X converted into alkane by protolysis

(xi) *The wittig reaction* **:**

$$
(Pf)_{3}P = CH_{2} + CH - R \longrightarrow (Pf)_{3}P = O + R - CH
$$

\n||
\nU\n
\nHydroboration: Alkene give a
\nwhich called hydroboration. In this

 Ph_3 *P* = *CH* – *R*+ *CH* – *R*——> $(Ph_3P = O + R - CH = CH - R$ compound || $(Ph)_{3}P = CH - R + CH - R -$

(xii) *From bromo ether* [Boord synthesis]

2 5 4 9 | | | *Br R CH CH R Zn R O CH CH Br R CH C H OH Zn*

(2) **Physical Properties**

(i) Alkenes are colourless and odourless.

(ii) These are insoluble in water and soluble in organic solvents.

(iii) Physical state

$$
C_1 - C_4 \longrightarrow \text{gas}
$$

\n
$$
C_4 - C_{16} \longrightarrow \text{liquid}
$$

\n
$$
C_{17} \longrightarrow \text{solid wax}
$$

\n
$$
C_{17} \longrightarrow \text{solid wax}
$$

\n
$$
C_{18} \longrightarrow \text{solid wax}
$$

 $C_2 - CH_2 - CH_2 \xrightarrow{Zndust} CH_2 + ZnX_2$ (iv) *B.P.* and *M.P.* decreases with increasing branches in

 $H_2 C$ *CH*₂ (v) The melting points of cis isomers are lower than trans isomers because cis isomer is less symmetrical than trans. Thus trans packs more tightly in the crystal lattice and hence has a higher melting point.

 $\frac{N_{\text{el}}}{\text{accelo}}$ $\sum_{-1/2} C \leftarrow C \leftarrow$ $\sum_{n=1}^{\infty} C \leftarrow$ (vi) The boiling points of cis isomers are higher than trans isomers because cis-alkenes has greater polarity (Dipole moment) than trans one.

(vii) These are lighter than water.

*CH*₂ dipole moments are higher than those of alkanes. (viii) **Dipole moment** : Alkenes are weakly polar. The, π electron's of the double bond. Can be easily polarized. Therefore, their

$CO₂ + H₂ + 2KOH$ (3) Chemical properties

CH (i) **F***rancis experiment* **:** According to Francis electrophile first attacks on olefinic bond.

Glasswood 450°	CH_3 – COOH	CH_2 = $CH_2 + Br - Br \xrightarrow{CCl_4} \rightarrow CH_2 -$						
$Eq. N_2$	$CH_2 = CH_2$	Br	Br					
Py ammonium compounds:	$NaCl$	Br	Br					
C_2H_5) ₃ $N + C_2H_4 + H_2O$	$ $	$ $	$ $					
$Ref. Amine)$	Etane	Br	Br	Br	Br	Br	Br	Cl
on vinyl chloride:	HH	HH	HH					
$2C = CHR$	$R - C = C - R + H_2 \xrightarrow{Ni} R - C - C - R$	$ $						

(iii) *Reduction of alkene via hydroboration* **:** Alkene can be

$$
RCH = CH_2 \xrightarrow{H-BH_2} (R-CH_2-CH_2)_3 B
$$

\n
$$
R \xrightarrow{H^+ + H_2O} R-CH_2-CH_3
$$

\n
$$
\xrightarrow{H^+ + H_2O} R-CH_2-CH_3
$$

H H

2 which called hydroboration. In this reaction formed trialkylborane, $CH₂$ **Hydroboration** : Alkene give addition reaction with diborane $R \longrightarrow (Ph, P = O + R - CH = CH - R$ compound *O* Which is very important and used for synthesis of different organic compound

$$
3R-CH = CH_2 + BH_3 \longrightarrow
$$

\n
$$
Br \qquad (R- CH_2 - CH_2)_3 B \text{ Trialkyl borane}
$$

\n
$$
O-C_2H_5 \qquad \qquad \downarrow NaOH/H_2O_2 \qquad \downarrow H/H_2O_2
$$

\n
$$
R-CH_2-CH_3 \qquad R-CH_2-CH_2OH \qquad R-CH_2-CH_3
$$

The overall result of the above reaction appears to be antimarkownikoff's addition of water to a double bond.

(iv) *By treatment with* $AgNO₃ + NaOH$ *: This reaction gives* coupling

$$
CH_3
$$

6CH₃ - CH₂ - CH₂ - C = CH₂ $\xrightarrow{B_2H_6}$

$$
2[CH_3-(CH_2)_2-C-CH_2]_3B \xrightarrow{\text{Agl NO}_3\text{NaOH}}
$$
\n
$$
CH_3
$$
\n
$$
CH_3-C-C-H+CH_3-C-C-H
$$
\n
$$
CH_3
$$
\n
$$
CH
$$

$$
CH_3 - CH_2 - CH_3
$$

\n
\n $CH_3 - CH_2 - CH_2 - CH_2 - CH_2 - CH_2 - CH_3$
\n
\n $CH_3 - CH_2 - CH_2 - CH_2 - CH_2 - CH_2 - CH_3$
\n
\n $CH_3 - CH_2 + H_2 + H_2 + H_2$
\n
\n $CH_3 - CH_2 - CH_2 - CH_2 - CH_2 - CH_2 - CH_3$
\n
\n $CH_3 - CH_2 - CH_2 - CH_2 - CH_2 - CH_2$
\n
\n $CH_3 - CH_2 - CH_2 - CH_2 - CH_2 - CH_2$

(v) *Birch reduction :* This reaction is believed to proceed via anionic free radical mechanism.

 $+e^{-}$

$$
R-CH = CH_2 \xrightarrow[\text{the}]{Na} R-CH-CH_2 \xrightarrow[\text{the}]{K} R-CH-CH_3 \xrightarrow[\text{the}]{K} R-CH-CH_2-CH_3 \xrightarrow[\text{the}]{C} H_2 + H^+ HSO_4 \xrightarrow[\text{the}]{C} CH_2 \xrightarrow[\text{the}]{K} R-CH-CH_3 \xrightarrow[\text{the}]{K} R-CH_2-CH_3 \xrightarrow[\text{the}]{K} R-CH_2-H_2 HSO_4 \xrightarrow[\text{the}]{K} CH_2 + H^+ HSO_4 \xrightarrow
$$

(vi) *Halogenation*

$$
CH_3CH = CH_2 + Cl_2 \xrightarrow{500^{\circ}C} CICH_2 - CH = CH_2 + HCl
$$
 (x) **Reaction with nitrosyl chloride**
\n^{Allylchloride}
\nor 3-Chloro-1-propene

 \Box If NBS [N-bromo succinimide] is a reagent used for the specific purpose of brominating alkenes at the allylic position.

$$
CH_3CH=CH_2 + \Big|_{CH_2-CO} \longrightarrow N-Br \longrightarrow
$$

Provene NBS

$$
CH2-CH=CH2+
$$

\n
$$
CH2-CO
$$
\n
$$
CH2-CO
$$
\n
$$
NH2-CO
$$

Allyl bromide Succinimide
 I In presence of polar medium alkene form vicinal dihalide with halogen.

$$
H H H
$$
\n
$$
R - C = C - H + X - X
$$
\n
$$
H H
$$
\n
$$
H H
$$
\n
$$
V
$$
\n
$$
C
$$
\n
$$
V
$$
\n $$

Reactivity of halogen is $F_2 > C l_2 > Br_2 > l_2$ (vii) *Reaction with HX* [Hydrohalogenation]

$$
\angle C = C \left\langle + HX \longrightarrow \begin{matrix} H & & & & \mathcal{N} \\ & \downarrow & & & \mathcal{N} \\ & \downarrow & & & \mathcal{N} \\ & \downarrow & & & H & H \\ & & \downarrow & & & \mathcal{N} \\ & & & & \mathcal{N} \end{matrix} \right\rangle
$$
With acidic *KMnO₄*:

According to markownikoff's rule and kharasch effect.

$$
CH_3-CH = CH_2 + HBr \longrightarrow CH_3-CH_3-C-H
$$
\n
$$
CH_3-CH = CH_2 + HBr \longrightarrow CH_3-C-H
$$
\n
$$
CH_3 \longrightarrow CH_3-C-H
$$
\n
$$
CH_3 \longrightarrow CH_3
$$
\n
$$
CH_3 \longrightarrow CH_3
$$
\n
$$
CH_3 \longrightarrow CH_3
$$
\n
$$
CH_3 \longrightarrow H_3
$$

According to Anti Markownikoff rule (Based on F.R.M.) CH_3 – $CH = CH_2 + HBr \xrightarrow{\text{Peroxide}}$

$$
HH \t H \t H \t H
$$

\n
$$
CH_3 - C - C - H + CH_3 - C - C - H
$$

\n
$$
H \t H \t H
$$

\n
$$
Br \t H \t H \t H
$$

\n
$$
(minor)
$$

\n
$$
(major)
$$

CH (viii) *Reaction with hypohalous acids* **:**

$$
CH_2 - CH_2 - CH_3
$$

\n $CH_2 = CH_2 + H \overline{O} \overline{C} \overline{C} \longrightarrow CH_2OH \cdot CH_2Cl$
\n Et hylenechlorohydin
\n Et lylenechlorohydin

H **I** In case of unsymmetrical alkenes markownikoff rule is followed.

(ix) *Reaction with sulphuric acid* **:**

$$
CH_2 = CH_2 + H^+ HSO_4 \longrightarrow CH_3 CH_2 HSO_4
$$

Etthylene

$$
CH_3 CH_2 HSO_4 \longrightarrow CH_2 = CH_2 + H_2 SO_4
$$

 $\frac{N_a}{N_a}$ $R - CH - CH_3 \xrightarrow{Et - O - H_1} R - CH_2 - CH_3$
 \Box This reaction is used in the seperation of alkene from a gaseous mixture of alkanes and alkenes.

$$
C = 0
$$

$$
C = 0
$$

Tillden reagent)

 \Box If hydrogen is attached to the carbon atom of product, the product changes to more stable oxime.

H glycol | | | | [] | | 4 *H H H HO OH H O H OH R C C H H R C C OH Alk KMnO*

$$
\begin{array}{ccc}\n\downarrow & & H & H & O \\
\downarrow & | & || & || & || \\
\text{and kharasch effect.} & & & R - C = C - H + [Q] \xrightarrow{\text{acidic}} R - C - O - H + CQ_2 + H_2O\n\end{array}
$$

H H (xii) *Hydroxylation*

H Br H , | (a) Using per oxy acid : Trans(racemic) 2-Butene 3 3 \sim 3 \sim 3 | *H C OH* | | | $H - C \xrightarrow{\text{H2O}_2, \text{HCO}_3} H - C - OH$ CH_3
 \downarrow
 H_2O_2 , HCOOH \rightarrow H_1C_1OH *CH HO C H CH CH H C* \rightarrow *C* \rightarrow *H*₂ O_2 , HCOOH \rightarrow *H* \rightarrow *C* \rightarrow *OH* $\begin{array}{ccc}\n\frac{H_2O_2, HCOOH}{or \ HCO_3H} & H-C-OH \\
\hline\n\end{array}$ $- C$ $H O - C - H$

$$
R \sim
$$

H R Trans

(b) Hydroxylation by
$$
OsO_4
$$
: $\begin{array}{ccc}\n C & \text{(Anti markownikoff's rule)} \\
 \downarrow \text{C} & \text{C} \\
 C & \text{A} & \text{A} \\
 \end{array}$ \n
\n $\begin{array}{ccc}\n R & \text{SUS} & \text{CUS} \\
 \downarrow \text{CUS} & \text{DUS} & \text{CUS} \\
 \end{array}$ \n
\n $\begin{array}{ccc}\n \downarrow & \text{CUS} \\
 \downarrow & \text{CUS} & \text{CUS} \\
 \end{array}$ \n
\n $\begin{array}{ccc}\n \downarrow & \text{CUS} \\
 \downarrow & \text{CUS} & \text{CUS} \\
 \end{array}$ \n
\n $\begin{array}{ccc}\n \downarrow & \text{CUS} \\
 \downarrow & \text{CUS} & \text{CUS} \\
 \end{array}$

H HO OH H R

 \Box If per benzoic acid or peroxy acetic acid is used then oxirane are formed. (eh)

$$
R-CH = CH - R - \frac{C_6H_5CO_3H}{\sigma CH_3CO_3H} \rightarrow R - CH - CH - R - \frac{H_2O}{\sigma CH_3CO_3H} \rightarrow R - CH - CH - R
$$
\n
$$
CH = CH - CH - CH - \frac{C_6H_5CO_3H}{\sigma CH_3CO_3H} \rightarrow R - CH - CH - R
$$
\n
$$
R - CH - CH - R
$$
\n
$$
R - CH - CH - R
$$
\n
$$
M_2C = O + H \longrightarrow [H_2C = O]
$$
\n
$$
\frac{R - CH - CH_2}{\sigma CH_3CO_3H} \rightarrow R - \frac{C_6}{\sigma CH_3CO_3H} \rightarrow R - \frac{C_7}{\sigma CH_3CO_3H} \rightarrow R - \frac{C_7}{\sigma CH_3CO_3H} \rightarrow R - \frac{C_7}{\sigma CH_3CO_3H}
$$
\n
$$
M = \frac{C_7}{\sigma CH_3CO_3H} \rightarrow R - \frac{C_7}{\sigma CH_3CO_3H}
$$

(xiii) *Combustion* :
$$
C_nH_{2n} + \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$

They burn with luminous flame and form explosive mixture with air or oxygen.

(xiv) *Ozonolysis*

 $Qzonide$
 Q Application of ozonolysis : This process is quite useful to locate the position of double bond in an alkene molecule. The double bond is obtained by Joining the carbon atoms. of the two carbonyl compounds.

(xv) *Oxy – mercuration demercuration* : With mercuric acetate (in THF), followed by reduction with *NaBH*₄ / *NaOH* is also an example of hydration of alkene according to markownikoff's rule.

$$
(CH_3)_3 C - CH = CH_2 + (CH_3 COO_2 Hg \longrightarrow CH_3 - CH_2 - Hg \longrightarrow CH_3 + (CH_3COO_2 Hg \longrightarrow CH_3)
$$
\n
$$
(CH_3)_3 C - CH - CH_2 - Hg \longrightarrow CH_3 + (CH_3) C - CH - CH_3
$$
\n
$$
(CH_3)_3 C - CH - CH_2 - Hg \longrightarrow CH_3 + (CH_3) C - CH - CH_3
$$
\n
$$
CH_3 - CH_2 - Hg \longrightarrow CH_3 + (CH_3) C - CH - CH_3
$$
\n
$$
CH_3 - CH_2 - Hg \longrightarrow CH_3 + (CH_3) C - CH - CH_3
$$
\n
$$
CH_2 = CH_2 + HO - NO_2 \longrightarrow CH_2 O + C_2
$$
\n
$$
CH_2 = CH_2 + HO - NO_2 \longrightarrow CH_2 O + C_2
$$
\n
$$
CH_3 \longrightarrow CH_2 - CH_2 + HO - O - O - C
$$
\n
$$
CH_3 \longrightarrow CH_2 - CH_2 + HO - O - O - C
$$
\n
$$
CH_3 \longrightarrow CH_2 - CH_2 + HO - O - O - C
$$
\n
$$
CH_3 \longrightarrow CH_2 - CH_2 + HO - O - O - C
$$
\n
$$
CH_3 \longrightarrow CH_2 - CH_2 + HO - O - O - C
$$
\n
$$
CH_3 \longrightarrow CH_2 - CH_2 + HO - O - O - C
$$

(xvi) *Epoxidation*

(a) By
$$
O_2
$$
 | $Ag : CH_2 = CH_2 + \frac{1}{2}O_2$ $\xrightarrow{Ag} CH_2 - CH_2$
\n $\xrightarrow{CH_2 = CH_2 + CH_3COCl \longrightarrow CH_2^{\text{R} + \text{C}}}$
\n $\xrightarrow{CH_2 \text{C} + \text{C} + \text{C} \longrightarrow CH_2^{\text{R} + \text{C}}}$
\n $\xrightarrow{CH_2 \text{C} + \text{C} + \text{C} \longrightarrow CH_2^{\text{R} + \text{C}}}$
\n $\xrightarrow{CH_2 \text{C} + \text{C} + \text{C} \longrightarrow CH_2^{\text{R} + \text{C}}}$

(b) Epoxidation by performic acid or perbenzoic acid :

$$
CH_2 = CH_2
$$

\n $CH_2 = CH_2$ \n CH_2

O

$$
CH_3 - CH = CH_2 \xrightarrow{H-C-O-O-H} CH_3 - CH - CH_2
$$
 carbo

(xvii) *Hydroboration*

$$
3R-CH = CH_2 + BH_3 \longrightarrow (R-CH_2-CH_2)_3B \xrightarrow{\text{Trialkyl borane}} (1) \text{ General methods of preparation}
$$
\n
$$
R-CH_2-CH_2-OH + BOH_3 \qquad \qquad CH_2Br-CH_2Br \qquad \qquad \text{alc KOH or NaNH}_2
$$
\n
$$
CH_2Br-CH_2Br \qquad \qquad \text{alc KOH or NaNH}_2
$$

$$
R-CH = CH_2 + CO + H_2 \xrightarrow{CotH(CO)_4} R-C-C-H
$$

\n
$$
H H
$$
\n
$$
C-C
$$
\n
$$
H H
$$
\n
$$
C=O
$$
\n
$$
H
$$
\n
$$
C=O
$$

$$
\Box
$$
 If $CO + H_2O$ is taken then respective acid is formed.
\n
$$
H = CH = CH - R - \frac{C_6H_5CO_3H}{C_6C_6H} \rightarrow R - CH - CH - R - \frac{-H_2O}{C_6C_6C_6H} \rightarrow R - CH = CH_2 + CO + H_2O - \frac{C_6H_1CO_4}{C_6C_6H} \rightarrow R - CH_2 - CH_2
$$

$$
R - CH - CH - R
$$
\n
$$
H_2C = 0 + H \longrightarrow [H_2C = \mathcal{O}H \longleftarrow H_2 \mathcal{O} - \mathcal{O}H_1
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3n}{2}O_2 \longrightarrow nCO_2 + nH_2O
$$
\n
$$
+ \frac{3
$$

 If in polymerisation zeigler- natta catalyst $[(R)_{3} A1 + TiCI_{4}]$ is used then polymerisation is known as zeiglernatta polymerisation.

(xxi) Isomerisation :
\n
$$
CH_3 - CH_2 - CH_2 - CH = CH_2
$$
 $\xrightarrow{AICl_3}$
\n $CH_3 - CH_2 - CH = CH - CH_3$

The mechanism proceeds via carbocation.

$$
CH_{2} = CH_{2} + HO - NO_{2} \longrightarrow CH_{2}OH.CH_{2}NO_{2}
$$
\n
$$
CH_{2} = CH_{2} + HO - NO_{2} \longrightarrow CH_{2}OH.CH_{2}NO_{2}
$$
\n
$$
(xxiii) Addition of Acetyl chloride :\n
$$
CH_{2} = CH_{2} + CH_{3}COCl \longrightarrow CH_{2}CICH_{2}COCH_{3}
$$
\n
$$
CH_{2} = CH_{2} + CH_{3}COCl \longrightarrow CH_{2}CICH_{2}COCH_{3}
$$
\n
$$
CH_{2} = CH_{2} + CH_{3}COCl \longrightarrow CH_{2}ClCH_{2}COCH_{3}
$$
\n
$$
CH_{2} = CH_{2} + CH_{3}COCl \longrightarrow CH_{2}ClCH_{2}COCH_{3}
$$
\n
$$
CH_{2} = CH_{2} + CH_{3}COCl \longrightarrow CH_{2}ClCH_{2}COCH_{3}
$$
\n
$$
CH_{2} = CH_{2} + CH_{3}COCl \longrightarrow CH_{2}ClCH_{2}COCH_{3}
$$
$$

 $Z_2 = CH_2$ $\begin{array}{ccc}\nC - 0 - 0 - H & O \\
\downarrow \\
\downarrow \\
C + 2 - C + 1 & O\n\end{array}$ poisonous mustard gas (War gas); (vi) For making ethylene-oxygen flame. (i) For the manufacture of polythene – a plastic material; (ii) For artificial ripening of fruits; (iii) As a general anaesthetic; (iv) As a starting material for a large number of compounds such as glycol, ethyl halides, ethyl alcohol, ethylene oxide, etc; (v) For making flame.

Alkynes

 \sum_{3} 2 \sum_{1} 2 \sum_{2} \sum_{1} \sum_{3} \sum_{1} \sum_{2} \sum_{1} \sum_{1} \sum_{1} \sum_{1} \sum_{1} \sum_{1} \sum_{1} \sum_{1} CH_3 – $CH = CH_2$ $\xrightarrow{H-C-O-O-H} CH_3$ – $CH-CH_2$ carbon triple bond are called alkynes. General formula is C_nH_{2n-2} . *O* These are the acyclic hydrocarbons which contain carbon-

(Anti markownikoff's rule)

COOH

(more stable)
\n
$$
\begin{array}{ccc}\n&\text{(more stable)}\\
&\text{(1)}\\
&\text{- } C = C - + N u^- &\longrightarrow & -C - C - \\
&\text{(alkyl carbonion)}\\
&\text{(less stable)}\n\end{array}
$$

(i) *Acidity of alkynes* **:** Acetylene and other terminal alkynes (1- alkynes) are weakly acidic in character

Ex.
$$
CH \equiv CH + \text{NaNH}_2 \longrightarrow H - C \equiv \overline{C} \text{Na}^+ + \frac{1}{2} H_2
$$

(Monosodium acetylide)

The acetylenic hydrogen of alkynes can be replaced by copper (I) and silver (I) ions. They react with ammonical solutions of cuprous chloride and silver nitrate to form the corresponding copper and silver alkynides.

$$
CH \equiv CH + 2[Cl(NH_3)_2]Cl \longrightarrow Cu - C \equiv C - Cu + 2NH_4Cl + 2NH_3
$$

Dicopper acetylide (Red ppt)

 $CH \equiv CH + 2[Aq(NH_3)_2]NO_3 \longrightarrow AgC \equiv C - Ag + 2NH_4NO_3 + 2NH_3$ Disilver acetylide (white ppt)

This reaction can be used to distinguish between 2-alkynes and 1-alkynes. 1-alkynes will give this test while 2-alkynes, will not give this test.

$$
CH_3 - C \equiv CH + 2[Ag(NH_3)_2] NO_3 \longrightarrow CH_3 - C \equiv C - Ag
$$
\n
$$
CH_3 - C \equiv C - CH_3 + 2[Ag(NH_3)_2] NO_3 \longrightarrow No \text{ reaction}
$$
\n
$$
CH_3 - C \equiv C - CH_3 + 2[Ag(NH_3)_2] NO_3 \longrightarrow No \text{ reaction}
$$

 $CH_3 - C \equiv C - CH_3$
Explanation for the acidic character : It explained by *SP* \Box $CH_3 - C \equiv CH + CH_3 - Mg - X \longrightarrow$ hybridisation. We know that an electron in *s*-orbital is more tightly $CH_3 - C \equiv C - Mg - X + CH_4 \xrightarrow{R-X} CH_3 - C \equiv C - R + MgX_2$ (50%) as compared to $s\beta^2$ (33%) or $s\beta^3$ (25%), due to large sheld than in a *p* -orbital. In *sp* hybridisation *^s* -character is more character the carbon atom is quite electronegative.

(ii) *Reaction with formaldehyde*

$$
HC \equiv CH + 2CH_2O \longrightarrow CH_2 - C \equiv C - CH_2 \xrightarrow{\text{Lii NH}_3} \rightarrow
$$

\n
$$
OH \qquad \qquad \downarrow
$$

\n
$$
CH_2 - CH = CH - CH_2OH \quad \text{[Trans-product]}
$$

\n
$$
OH
$$

(4) **Chemical properties of acetylene**

 In reaction with gem dihalide, Alc. *KOH* is not used for elimination in 2nd step.

 \Box In reaction with vicinal dihalide, if the reactant is 2-butylene chloride then product is 2-butyne as major product.

Preparation of higher alkynes (by metal acetylide)

□ Acetylene gives salt with *NaNH*₂ or *AgNO*₃ (ammonical) which react with alkyl halide to give higher alkyne.

$$
\Box
$$
 2CH = CH \longrightarrow $Na \rightarrow C$ = C - $Na \longrightarrow$ $CH_3 - C$ = C - CH₃ $CH_3 - C$ = C - CH₃ + 2[A]
\n $CH_3 - C$ = CH + CH₃ - Mg - X \longrightarrow CH_{3} *Explanation for the acidic*
\n $CH_3 - C$ = C - Mg - X + CH₄ \longrightarrow $CH_3 - C$ = C - R+ MgX₂ \longrightarrow $CH_3 - C$ = C - Mg - X + CH₄ \longrightarrow $CH_3 - C$ = C - R+ MgX₂ \longrightarrow (50%) as compared to sp^2 (33%)

(2) **Physical properties**

(i) Acetylene is a colourless gas. It has a garlic odour. The odour is due to presence of impurities of phosphorous and hydrogen sulphide. However, pure acetylene has pleasant odour.

(ii) It is insoluble in water but highly soluble in acetone and alcohol. Acetylene is transported under high pressure in acetone soaked on porous material packed in steel cylinders.

(iii) Its boiling point is $-84^\circ C$.

(iv) It is lighter than air. It is somewhat poisonous in nature.

(v) It burns with luminous flame and forms explosive mixture with air.

(3) **Chemical reactivity of alkynes** : $C = C$ is less reactive than the carbon-carbon double bond towards electrophilic addition reaction. This is because in alkyne carbon has more S-character so more strongly will be the attraction for π electrons. Alkyne also undergo nucleophilic addition with electron rich reagents. Ex. Addition of water, cyanide, carboxylic acid, alcohols. Nucleophilic addition can be explained on the basis that alkynes form vinylic carbanion which is more stable than alkyl carbanion formed by alkene

$$
PU
$$
\n
$$
-C \equiv C - + N\bar{U} \longrightarrow -C = C -
$$
\n
$$
Vinylic \text{ carbonion}
$$
\n
$$
CH \equiv CH
$$
\n
$$
CH \equiv CH
$$

$$
C = C - C
$$

Degree of C ^{atalysts at Uranon $\overrightarrow{c_{is}}$ The number of degree of} unsaturation in a hydrocarbon is given by

 $\frac{2n_1 + 2 - n_2}{2}$, Where n_1 is the number of carbon atoms; n_2

is the number of hydrogen atoms.

For example in C_6H_{12} , the degree of unsaturation is

$$
=\frac{2\times 6+2-12}{2}=1
$$

Tests of unsaturation

(a) *Baeyer's reagent* : It is 1% **KMnO**₄ solution containing sodium carbonate. It has pink colour. An aqueous solution of the compound, a few drops of Baeyer's reagent are added, the pink colour of the solution disappears. The decolourisation of pink colour indicates the presence of unsaturation in the compound.

$$
\Box
$$
 Alkene without any hydrogen atom on the carbon
forming the double bond

$$
\uparrow R
$$

$$
\uparrow R
$$

$$
\uparrow R
$$

don't show this test.

Oxidative–Hydroboration : Alkynes react with BH_3 (in THF) $\frac{61}{3}$ $\frac{376}{6}$ bromine solution in carbon tetrachioride are added to it, the colour of bromine disappears. It indicates the presence of (b) *Bromine- carbon tetrachloride test* : The compound is dissolved in carbon tetrachloride or chloroform and then a few drops of 5% bromine solution in carbon tetrachloride are added to it, the colour of bromine disappears. unsaturation.

> *R R*

$$
C=C
$$

(i) Acetylene is used as an illuminant.

(ii) It is used for the production of oxy-acetylene flame. The temperature of the flame is above $3000^\circ C$. Is is employed for cutting

(iv) It is used as a general anaesthetic under the name naracylene.

(v) Acetylene has synthetic applications. It serves as a starting material for the manufacture of a large variety of substances.

(vi) On electrical decomposition acetylene produces finely divided carbon and hydrogen. Hydrogen is used in airships. $C_2H_2 \longrightarrow 2C + H_2$

and finally converted into carbonyl compounds.

3*CH*₃ - *C* = *CH*
$$
\xrightarrow{BH_3 + THF}
$$
 (CH₃ - *CH* = *CH*)₃ *B* $\xrightarrow{H_2O_2}$
\n $CH_3 - CH = CHOH$ $\xrightarrow{Tautomeries}$ CH_3CH_2CHO \xrightarrow{R} \xrightarrow{R}
\nOr $\xrightarrow{H_2SO_4}$ $CH_3 - CH_3 - C - CH_3$ (Acetone) (5) *Uses*

Thus it is useful for preparing aldehyde from terminal alkyne.

Reduction of Alkyne **:** Alkynes add on hydrogen in presence of suitable catalysts like finely divided *Ni, Pd.*

$$
CH \equiv CH + H_2 \xrightarrow{Ni} CH_2 = CH_2 \xrightarrow{Ni} CH_3 - CH_3
$$
 and welling of metals.
(iii) Acetylene is used for artificial ripening of fruits.

If the triple bond is not present at the end of the carbon chain of the molecule, the alkene formed may be cis and trans depending upon the choice of reducing agents.

With *Nal NH*₃ or *Lil NH*₃ in (liquid ammonia) trans alkene is almost an exclusive product while catalytic reduction at alkyne affords mainly cis alkenes.

(6) **Interconversion**

(i) **Conversion of ethane into ethene :** (Alkane into alkene)

$$
CH_3 - CH_3 \xrightarrow{\text{B}_2} C_2 H_5 \text{Br} \xrightarrow{\text{Alc}} CH_2 = CH_2
$$

\nEthan (a) $CH_3 - CH_3 \xrightarrow{\text{B}_2} C_1 H_5 \text{Br}$
\n $CH_3 - CH_3 \xrightarrow{\text{B}_2} C_4$

(ii) **Ethene into ethane :** (Alkene into alkane)

$$
CH_2 = CH_2 \xrightarrow{H_2} \xrightarrow{H_2} CH_3 - CH_3
$$

Ethene
$$
N_1 300^{\circ}C
$$
Ethane

(iii) **Ethane into ethyne (acetylene) :** *i.e.,* alkane into alkyne

$$
{}_{51}^{2}C H_{3} \xrightarrow{Bf_{2}} CH_{3}CH_{2}Br \xrightarrow{A/C} CH_{2} = CH_{2} \xrightarrow{Bf_{2}} CH_{3}
$$
 forms a
Ethané
 ${}_{\text{Ethane}}$

$$
1, 2-Dibromoethane
$$
 or *Neiv*
$$
V = \frac{1}{2}
$$

$$
CH \equiv CH \xrightarrow{H_2} CH_2 \rightarrow CH_2 = CH_2 \xrightarrow{H_2} CH_3 - CH_3
$$

\n $CH \equiv CH \xrightarrow{H_2} CH_2 \rightarrow CH_2 \rightarrow CH_3$
\n $CH \equiv CH \xrightarrow{H_2} CH_2 = CH_2 \xrightarrow{H_2} CH_3 - CH_3$
\n $CH \equiv CH \xrightarrow{H_2} CH_2 = CH_2 \xrightarrow{H_2} CH_3 - CH_3$
\nThe remaining gaseous mixture is passed through concentrated

- (v) **Ethene into propene : Ascending in alkene series**
- $[H]$ $[$ $]$ $CH_2 = CH_2 \xrightarrow{H \rightarrow} CH_3 CH_2 I \xrightarrow{KCN} CH_3 CH_2 CM \xrightarrow{[H]} \text{Reduction}$

Ethene **Propanentitrile Reduction**

(Ethyl cyanide)

$$
CH_3CH_2CH_2NH_2 \xrightarrow{HNO_2} CH_3CH_2CH_2OH
$$

\n1-Propanol
\n $CH_3CH_2CH_2NH_2 \xrightarrow{HNO_2} CH_3CH_2OH$
\n 1 -Propanol
\nThe methane or ethane is left by

$$
CH_3CH = CH_2 \xleftarrow[KOH]{\text{ACH}} CH_3CH_2CH_2Br \xleftarrow{PB_5}
$$

Propene

- or $\begin{equation} CH_2 = CH_2 \xrightarrow{Hl} CH_3 CH_2 I \xrightarrow{L(CH_3)_2 \text{CL}} CH_3 CH_2 CH_3 \xrightarrow{E} CH_3 \xrightarrow{CH_3CH_3} CH_3 \xrightarrow{E} CH_3 CH$ $CH_2 = CH_2 \xrightarrow{HI} CH_3 CH_2 I \xrightarrow{Li(CH_3)_2 Cu} CH_3 CH_2 CH_3$
Ethene lodoethane Propane
- $\frac{P_1}{P_1}$ $CH_3CH = CH_2$
Propene $\frac{2}{\mu} \rightarrow CH_3CH_2CH_2CH \frac{AIC}{KOH} \rightarrow CH_3CH = CH_2$

Propene Propene $\frac{Cl_2}{h\nu}$ > CH₃ CH₂ CH₂ CI $\frac{Alc}{h\sqrt{c}}$ > CH₃ CH = CH₂
 $\frac{Alc}{h\sqrt{c}}$ + Characterscape KOH
- or $CH_2 = CH_2$ $\xrightarrow{H1} CH_3CH_2I \xrightarrow{CH_3I1 Na} CH_3CH_2CH_3$ Propane **Propane**

$$
\frac{Cl_2}{h\nu} \rightarrow CH_3CH_2CH_2Cl \frac{Alc}{KOH} \rightarrow CH_3CH = CH_2
$$

¹-Chloro propane ^{KOH} Propene

(vi) **Propene into ethene :** Descending an alkene series

$$
CH_3-CH = CH_2 \xrightarrow{\begin{array}{c} O_3/H_2O \longrightarrow \\ \text{Fropene} \end{array}} CH_3CH_2 \xrightarrow{\begin{array}{c} O_3/H_2O \longrightarrow \\ \text{Ethanal} \end{array}} CH_3CHO \xrightarrow{\begin{array}{c} \text{LH} \\ \text{L4}H\text{H}_4 \end{array}} CH_2 = CH_2
$$
\n
$$
CH_3CH_2OH \xrightarrow{\begin{array}{c} H_2SO_4 \longrightarrow \\ \text{Ethanol} \end{array}} CH_2 = CH_2
$$

(vii) **Acetylene into propyne (methyl acetylene) :** (Ascent)

$$
CH \equiv CH \xrightarrow{\text{Na}} CH \equiv C\text{Na} \xrightarrow{\text{CH}_3} CH \equiv C - CH_3
$$

Monosodium
acetyliche Propyne

(viii) **Propyne into acetylene :** (Descent)

$$
CH_3C = CH
$$
—
\n $^{CH_3C} = CH$ —
\n Pcopyne
\n ^{PCH}CHO — $^{PCL} \rightarrow CH$ CHCl — $^{Alc} \rightarrow CH = CH$
\n ^{CH}CHO — $^{PCL} \rightarrow CH$ CHCl — $^{Alc} \rightarrow CH = CH$

$$
CH_3CHO \xrightarrow{PC \xi} CH_3CHCl_2 \xrightarrow{Alc} CH \equiv CH
$$
\n
\nAcetaldehyde
\nChloride
\nchloride

$$
(ix)
$$
 1-Butyne into 2-pentyne : (Ascent)

 $CH_3CH_2C \equiv CH \xrightarrow{\textit{NaNH}_2} CH_3CH_2C \equiv C-Na \xrightarrow{\textit{CH}_3I} \rightarrow$ $1 - B$ utyne

$$
CH_3CH_2-C \equiv CCH_3
$$

2-Perityne

(x) **1-Butyne into 2-pentanone :** (Not more than three steps)

$$
CH_3CH_2C \equiv CH \xrightarrow{\text{N}a\text{N}H_2} CH_3CH_3CH_2C \equiv CNa \xrightarrow{\text{CH}_3} \rightarrow
$$

1-Butyne

$$
CH_3CH_2C \equiv CCH_3 \xrightarrow{H_2O, H_2SQ_4} CH_3CH_2CH_2CCH_3
$$
\n
$$
CH_3CH_2C \equiv CCH_3 \xrightarrow{H_2O, H_2SQ_4} CH_3CH_2CH_2CCH_3
$$
\n
$$
H_3C \equiv CCH_3 \xrightarrow{H_3Q_4} CH_3CH_2CH_2CCH_3
$$

Separation of alkane, alkene and alkyne

 $CH_2 = CH_2 \xrightarrow{H_2} CH_3 - CH_3$ CH_3 *CH*₃ $\frac{2}{\text{Ethene}}$ $\frac{CCl_4}{CCl_4}$ recovered by decomposition of the precipitate with an acid. $CH_3 - CH_3 \xrightarrow{Bf_2} CH_3 CH_2 Br \xrightarrow{A/c} CH_2 = CH_2 \xrightarrow{Bf_2}$ forms a red precipitate. It is filtered. The alkyne or acetylene is chloride solution. The alkyne (acetylene) reacts with $Cu₂Cl₂$ and

$$
CH_2Br - CH_2Br - \frac{Alc KOH}{\text{or NaNH}_2} \rightarrow CH \equiv CH
$$
\n
$$
C_2H_2 + Cu_2Cl_2 + 2NH_4OH \rightarrow C_2Cu_2 + 2NH_4Cl + 2H_2O
$$
\n
$$
C_2H_2 + Cu_2Cl_2 + 2NH_4OH \rightarrow C_2Cu_2 + 2NH_4Cl + 2H_2O
$$
\n
$$
C_2Cu_2 + 2HNO_3 \rightarrow C_2H_2 + Cu_2(NO_3)_2
$$
\n
$$
C_2Cu_2 + 2HNO_3 \rightarrow C_2H_2 + Cu_2(NO_3)_2
$$

 $\overrightarrow{\text{Reduction}}$ heated at 170 \degree C to regenerate ethene. H_2 *SO*₄. Alkene is absorbed. The Hydrogen sulphate derivatives is

$$
C_2H_4 + H_2SO_4 \longrightarrow C_2H_5HSO_4 \xrightarrow[170^{\circ}C]{} C_2H_4 + H_2SO_4
$$

The methane or ethane is left behind unreacted.

Distinction between alkanes, Alkenes and Alkynes

Property	Alkane (Ethane)	Alkene (Ethene)	Alkyne (Ethyne)
Molecular formula	$C_nH_{2n+2}(C_2H_6)$	$C_nH_{2n}(C,H_4)$	$C_nH_{2n-2}(C_2H_2)$
Nature	Saturated	Unsaturated	Unsaturated
	Single bond between carbon atoms. Each carbon atom is sp^3 -hybridized $C - C$ Bond length 1.34 Å Bond energy: 83 Kcal mol ⁻¹	Double bond between two carbon atoms. Both carbon atoms are sp^2 -hybridized $C = C$ 146 Kcal mol ⁻¹	Triple bond between two carbon atoms both carbon atoms are sp-hybridized $-C \equiv C -$ 1.20 Å 200 Kcal mol ⁻¹
	Burns with nonluminous flame	Burns with luminous flame	
Burning	$C_2H_6+7/2O_2 \rightarrow 2CO_2+3H_2O$	$C_2H_4+3O_2 \rightarrow 2CO_2+2H_2O$	Burns with smoky flame $C_2H_2+5/2O_2 \rightarrow 2CO_2+H_2O$
Reaction with H_2		Forms alkane	Forms alkene and alkane
		$C_nH_{2n} + H_2 \xrightarrow{Ni} C_nH_{2n+2}$ 300° C Alkane $C_2H_4 + H_2 \rightarrow C_2H_6$	$C_nH_{2n} + H_2 \xrightarrow[300^{\circ}C]{} C_nH_{2n+2}$ Alkane $C_nH_{2n-2} + H_2 \xrightarrow{Ni} C_nH_{2n}$ 300°C Alkene
Reation with conc. H_2SO_4 and hydrolysis		Addition $C_2H_4+H_2SO_4 \rightarrow C_2H_5HSO_4$ $H_2O \rightarrow C_2H_5OH$ Alcohol	Addition $C_2H_2 \rightarrow CH_3CH(HSO_4)_2 \xrightarrow{H_2O}$ CH ₃ CHO Aldehyde
Br_2/CCl_4	$\overline{}$	Decolourises Dibromo derivative, $C_2H_4 + Br_2 \rightarrow C_2H_4Br_2$	Decolourises Tetrabromo derivative, $C_2H_2Br_4$
Baeyer's reagent (Alk. KMnO ₄	$\overline{}$	Decolourises Glycol is formed CH ₂ CH ₂ OH $ + H_2O+O \rightarrow $ CH ₂ CH ₂ OH	Decolourises Oxalic acid is formed CH COOH \parallel +40-+ CH COOH
Ammonical Cu_2Cl_2	\equiv	\equiv	Red precipitate CH CCU + Cu_2Cl_2 + 2 NH ₄ OH \rightarrow CH CCU (Red) $+ 2NH_4Cl + 2H_2O$
Ammonical silver nitrate		$\overline{}$	White precipitate CH $C - Ag$ $ $ + 2AgNO ₃ + 2NH ₄ OH \rightarrow CH $C - Ag$ $+ 2NH_4Cl + 2H_2O$

Table : 24.3

Cycloalkane

(1) **Methods of preparation**

(i) *From dihalogen compounds* (Freund reaction)**:**

(ii) *From alkenes* **:**

$$
CH_3-CH=CH_2+CH_2I_2 \xrightarrow{Zn-Cu \text{ alloy}} CH_3-CH-CH_2
$$

Propene
Graphs
Corbocyclic compounds with double bon

(iii) *From Aromatic compounds*

(i) First two members are gases, next three members are liquids and higher ones are solids.

(ii) They are insoluble in water but soluble in alcohol and ether.

(iii) Their boiling points show a gradual increase with increase of molecular mass. Their boiling points are higher than those of isomeric alkenes or corresponding alkanes.

(iv) Their density increase gradually with increase of molecular mass.

(3) **Chemical properties :** Cycloalkanes behave both like alkenes and alkanes in their chemical properties. All cycloalkanes undergo substitution reaction with halogen in the presence of light (like alkane). All cycloalkane (lower members) undergo addition reaction (ex. Addition of H_2 , HX , X_2). Further the tendency of forming addition compounds decreases with increase in size of ring cyclopropane > Cyclobutane > Cyclopentane. Relative ring opening of ring is explained by Baeyer strain theory.

(i) *Addition in spiro cycloalkane* **:** If two cycloalkane fused with one another then addition take place in small ring

Because small ring is more unstable than large ring *Spiro compound*

Higher cycloalkanes do not give addition due to more stability.

(ii) *Free radical substitution with Cl***²**

(iii) *Addition reaction*

Methyl cyclopropane called cycloalkenes. Some of the common cycloalkenes are $\hat{C}H_2$ Carbocyclic compounds with double bonds in the ring are

Cycloalkenes can be easily obtained by Diels-Alder reaction. Cyclobutene Cyclopentene Cyclohexene 1, 4-Cyclohexadiene

These compounds undergo the electrophilic addition reactions which are characteristic of alkenes, while the ring remains intact. Cycloalkenes decolourise the purple colour of dilute cold **KMnO**₄ or

red colour of bromine in carbon tetrachloride.

Dienes^{xene)} These are hydrocarbon with two carbon-carbon double bonds. *CH²*

Dienes are of three types (1) **Conjugated dienes :** Double bonds are seperated by one

single bond. Ex : $CH_2 = CH - CH = CH_2 (1, 3 \text{-} butadiene)$

CH (2) **Cumulative dienes** : Double bonds are adjacent to each

other.

Ex : $CH_2 = C = CH_2$ Propadiene [allene]

(3) **Isolated or Non-conjugated** : Double bonds are separated by more than one single bond.

Ex : $CH_2 = CH - CH_2 - CH = CH_2 (1, 4 \text{ pentadiene})$

The general formula is C_nH_{2n-2} . The predominant member of this class is 1, 3-butadiene.

(1) **Method of preparation**

(i) From acetylene :
\n
$$
2HC \equiv CH - \frac{C\mu_2 C_2}{NH_4Cl} + IC \equiv C - CH = CH_2
$$
\n
$$
CH_2 = CH - CH = CH_2
$$
\n
$$
CH_2 = CH - CH = CH_2
$$
\n
$$
CH_2 = CH - CH = CH_2
$$
\n
$$
CH_2 = CH - CH = CH_2
$$
\n
$$
CH_2 = CH - CH = CH_2
$$
\n
$$
H_3 = 0
$$
\n
$$
H_4 = 0
$$
\n
$$
H_5 = 0
$$
\n
$$
H_6 = 0
$$
\n
$$
H_7 = 0
$$
\n
$$
H_8 = 0
$$
\n
$$
H_9 = 0
$$
\n
$$
H_9 = 0
$$
\n
$$
H_1 = 0
$$
\n
$$
H_2 = 0
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\n
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H_3 = 0
$$
\n
$$
H_4 = 0
$$
\n
$$
H_5 = 0
$$
\n
$$
H_7 = 0
$$
\n
$$
H_8 = 0
$$
\n
$$
H_9 = 0
$$

(ii) *From* **1***,* **4***-dichlorobutane* : 1, 3-Butadiene $\overline{CH_2}$ CH_2CH_2 $\overline{CH_2}$ \longrightarrow $\overline{CH_2}$ $=$ $CH - CH = CH_2$ \longrightarrow \overline{C} \longrightarrow \overline{C} \longrightarrow \overline{C} \longrightarrow \overline{C} 1,4-Dichlorobutane $_2$ \cup $_2$ \cup $_2$ \cup $_2$ \longrightarrow \cup $_2$ \to \cup $_1$ \to \cup | | *Cl Cl*

(iii) *From* **1,4-***butanediol* : 1, 3-Butadiene 2 2 2 4 *CH CH CH CH CH CH CH CH ^H SO* $CH_2CH_2CH_2CH_2$ CH_2 | | *OH OH* (iv) *From butane* :

1, 3-Butadiene *C* $CH_3CH_2CH_2CH_3 \xrightarrow{\text{Catalyst}} CH_2 = CH - CH = CH_2$

n-Butane $600^{\circ}C$

1,3-Butadiene (v) Oranglusia

(*Cr2O³* used as catalyst.) (v) *From cyclohexene* :

$$
\bigcup_{1,3 \text{ - } CH_2} = CH - CH = CH_2 + CH_2 = CH_2
$$
\n
$$
\bigcup_{\text{the one}}^{O_3} \bigcup_{\text{the one}}^{O_3
$$

(2) **Physical property :** 1,3-butadiene is a gas.

(3) **Chemical properties**

(i) *Addition of halogens* **:**

$$
CH_2 = CHCH = CH_2 + Br_2
$$
\n
$$
CH_2 = CHCH = CH_2 + Br_2
$$
\n
$$
CH_2 = CHCH = CH_2 + Br_2
$$
\n
$$
CH_2
$$
\n

1,4-Dibromo-2-butene (1, 4-Addition) predominates (70%) in an ionising solvent (acetic acid)

(ii) *Addition of halogen acids* : *CH² =CH–CH=CH² +HBr CH3CHBrCH=CH² (1, 2-Addition) 3-Bromo-1-butene (Major yield at low temp.) CH3–CH=CH–CH2Br (1, 4-Addition) 1-Bromo-2-butene*

(iii) *Addition of water* : *(Major yield at high temp.) CH² =CH–CH=CH² +H2O CH3CHOHCH=CH² But-3-en-2-ol*

(iv) *Polymerisation* : 1, 3-Butadiene *CH3CH=CHCH2OH But –2-en-1-ol*

Diels-alder reaction :

Stability of conjugated dienes : It is explained on the basis of

⁴ The four π electrons of 1, 3-butadiene are delocalised over all $CH_2 = CH - CH = CH_2$ the four atoms. This delocalisation of the π electrons makes the notecule more stable. molecule more stable.

CH2–O–CH – CH–O–CH²

(1) **Source of Arenes**

Source of arenes is coal. It contains benzene, xylene, naphthalene etc. Arenes are obtained by destructive distillation of coal.

(2) **Distillation of coal**

COA

Table : 24.4

 \Box The residue left after fractional distillation of coal-tar is called pitch.

(4) **Isolation of benzene**

Light oil $\frac{\text{cold }H_2SO_4}{\text{red}}$ Basic impurities removed $\frac{NaOH}{\text{red}}$ have $4n\pi$ electrons Phenolsremoved $\xrightarrow{\text{distillation}}$ Benzene(255 - 257 K)
 Folving
 Folving Toluane(383 K)

General characteristics of arenes

(1) All arenes have general formula $[C_n H_{2n} - 6y]$. Where *y* is number of benzene rings and *n* is not less than 6.

(2) Arenes are cyclic and planar. They undergo substitution rather than addition reactions.

(3) **Aromaticity or aromatic character :** The characteristic behaviour of aromatic compounds is called aromaticity. Aromaticity is due to extensive delocalisation of π -electrons in planar ring system. Huckel (1931) explained aromaticity on the basis of following rule.

Huckel rule **:** For aromaticity the molecule must be planar, cyclic system having delocalised $(4n+2)\pi$ electrons where *n* is an integer equal to 0, 1, 2, 3,------.

Thus, the aromatic compounds have delocalised electron cloud of 2,6,10 or 14 π electron

Naphthalene 10π electrons

Similarly cyclolpentadienyl anion or tropylium ion are also aromatic because of containing 6π electrons ($n=1$).

Anthracene 14π electrons

electrons (*y* action compounds also that $(n \theta x)$ electrons (φ atto \hbar)($n=0$) Tropyllium ion 6 π Cyclopropenyl Cyclopentadienyl anion 6π

Molecules do not satisfy huckel rule are not aromatic.

 $\frac{\text{cold } H_2 \text{SG}_4}{\text{8}}$ Basicimpurities removed $\frac{N^2OH}{\text{8}}$ have $4n\pi$ electrons are called antiaromatic compounds and this 8π electrons
(4) **Antiaromaticity** : Planar cyclic conjugated species, less stable than the corresponding acyclic unsaturated species are called antiaromatic. Molecular orbital calculations have shown that such compounds have $4n\pi$ electrons. In fact such cyclic compounds which characteristic is called antiaromaticity.

Toluene(383 K) antiaromatic compound because it has $4n\pi$ electrons ($n = 1$) and it is enolsremoved—^{distillation}→Benzene(255 - 257*K*) *Example :* 1,3-Cyclobutadiene, It is extremely unstable
[Acidic impurities] **Toluene(383** *K***)** antigrametic compound because it has $A \rho \pi$ electrons $(\rho - 1)$ and it is less stable than 1,3 butadiene by about 83.6 *KJ mol*–1 .

$$
n = 4
$$

Thus, cyclobutanediene shows two equivalent contributing structures and it has $n = 1$.

Benzene is the first member of arenes. It was first discovered by Faraday (1825) from whale oil. Mitscherllich (1833) obtained it by distillating benzoic acid with lime. Hofmann (1845) obtained it from coal tar, which is still a commercial source of benzene.

(1) **Structure of benzene :** Benzene has a special structure, which is although unsaturated even then it generally behave as a saturated compound.

(i) *Kekule's structure* **:** According to Kekule, in benzene 6 carbon atoms placed at corner of hexagon and bonded with hydrogen and double bond present at alternate position.

(a) *Evidence in favour of Kekule's structure*

• Benzene combines with 3 molecules of hydrogen or three molecules of chlorine. It also combines with 3 molecules of ozone to form triozonide. These reactions confirm the presence of three double bonds.

 Studies on magnetic rotation and molecular refraction show the presence of three double bonds and a conjugated system.

 The synthesis of benzene from three molecule of acetylene also favour's Kekule's structure.

Benzene gives cyclohexane by reduction with hydrogen.

Unusual stability of benzene.

 According to Kekule, two ortho disubstituted products are possible. But in practice only one ortho disubstituted product is known.

 Heat of hydrogenation of benzene is 49.8 *kcal/mole,* whereas theoretical value of heat of hydrogenation of benzene is 85.8 *kcal/mole.* It means resonance energy is 36 *kcal/mole.*

 \bullet $C-C$ bond length in benzene are equal, (although it contains 3 double bonds and 3 single bonds) and are 1.39 Å.

Kekule explained this objection by proposing that double bonds in benzene ring were continuously oscillating between two adjacent positions.

(ii) *From benzene derivatives* Sodium benzoate

(a) *From phenol* :

OH
\n
$$
\begin{array}{ccc}\n\text{OH} \\
\hline\n\text{Phenol} \\
\text{(b) From chlorobenzene :} \\
\end{array}
$$
\n
$$
\begin{array}{ccc}\n\text{distill} \\
\hline\n\text{Benzene} \\
\end{array}
$$
\n
$$
\begin{array}{ccc}\n\text{Benzene} \\
\end{array}
$$

heat

Benzene

$$
\begin{array}{c}\nCl \\
\hline\n\end{array}\n\qquad + 2H \xrightarrow{\text{Ni-Al alloy}} \begin{array}{c}\n\hline\n\end{array}\n\qquad + HCl
$$
\nChlorobenzene

(c) *By first preparing grignard reagent of chlorobenzene and then hydrolysed*

$$
C_6H_5Cl \xrightarrow{Mg} C_6H_5MgCl \xrightarrow{H_2O} C_6H_6+Mg
$$
\nChlorobenzene

\nChlorobenzene

\nChlorobenzene

\nChlorobenzene

\nChlorobenzene

\nChlorobenzene

\nChlorobenzene

\nChlorobenzene

\nChlorobenzene

\nChloride

\nAdition of hydrogen: Benzene

(d) *From benzene sulphonic acid* :

(e) *From benzene diazonium chloride* :

 \Box Cyclic polymerisation takes place in this reaction.

(g) Aromatisation:
$$
C_6H_{14} \xrightarrow{C_2O_3/A_2O_3} C_6H_6 + 4H_2
$$

 $\xrightarrow{n-Hexane} 500^\circ C$ Benzine

(3) **Properties of benzene**

(i) *Physical properties*

(a) Benzene is a colourless, mobile and volatile liquid. It's boiling point is 80°*C* and freezing point is 5.5°*C*. It has characteristic odour.

(b) It is highly inflammable and burns with sooty flame.

(c) It is lighter than water. It's specific gravity at 20°C is 0.8788.

(d) It is immiscible with water but miscible with organic solvents such as alcohol and ether.

(e) Benzene itself is a good solvent. Fats, resins, rubber, etc. dissolve in it.

(f) It is a non-polar compound and its dipole moment is zero.

(g) It is an extremely poisonous substance. Inhalation of vapours or absorption through skin has a toxic effect.

(ii) *Chemical properties* : Due to the presence of π electron clouds above and below the plane benzene ring, the ring serves as a source of electrons and is easily attacked by electrophiles (Electron loving reagents). Hence electrophilic substitution reaction are the characteristic reactions of aromatic compounds.

Substitution reactions in benzene are prefered rather than addition are due to the fact that in the former reactions resonance stabilised benzene ring system is retained while the addition reactions lead to the destruction of benzene ring. Principal reactions of benzene can be studied under three heads,

- (a) Addition reactions
- (b) Substitution reactions

(c) Oxidation reactions

OH (a) *Addition reactions* : In which benzene behaves like

Addition of hydrogen : Benzene reacts with hydrogen in the presence of nickel (or platinum) as catalyst at 150°C under pressure to form cyclohexane.

Unimolecular : Mostly uncommon in aromatic substitution, there is only one example which obtain in benzene diazonium dichloride.

Bimolecular :

Example : hexadienyl anion)

Elimination-addition mechanism (Benzyne mechanism)

 Electrophilic substitution reaction : Benzene undergoes this reaction because it is an electron rich system due to delocalized π electrons.

 Free radical aromatic substitution : The aromatic substitution reactions which follow free radical mechanisms are very few and have limited synthetic value. But some typical example of these reactions are:

The mechanism of chlorination of behalene at high temperature is similar to that of the free radical aliphatic substitution

$$
Cl_2 \longrightarrow Cl + Cl \text{ (Chain initiation)}
$$

$$
C_6H_6 + Cl \longrightarrow C_6H_5 + HCl \text{ (}H\text{-} \text{ abstraction)}
$$

 $C_{\epsilon}H_{\epsilon} + Cl_2 \longrightarrow C_{\epsilon}H_{\epsilon}Cl + Cl$ (Chain propagation)

(c) *Oxidation* :
$$
2C_6H_6 + 15O_2 \longrightarrow 12CO_2 + 6H_2O
$$

H 6530 *kJ/mole*

When vapours of benzene and air are passed over vanadium pentoxide at 450 – 500°*C*, maleic anhydride is obtained.

$$
C_6H_6+9[O] \xrightarrow[450-500^\circ C] \xrightarrow[450-500^\circ
$$

□ Strong oxidising agents converts benzene slowly into *CO*₂ and water on heating.

(d) *Reduction* :

(iii) *Uses* **:** (a) In dry cleaning (b) As a motor fuel when mixed with petrol. (c) As a solvent. (d) In the manufacture of gammexane (As insecticide). (e) In the preparation of nitrobenzene, chlorobenzene, benzene sulphonic acid, aniline, styrene, etc. Many of these are employed for making dyes, drugs, plastics, insecticides, etc.

Directive effect in substituted benzene derivatives

(1) **Directive effect in mono substituted benzene derivatives :** The substituent already present on the benzene ring directs the

incoming substituent to occupy ortho $(2 \text{ or } 6)$, meta $(3 \text{ or } 5)$ or para (4) position. This direction depends on the nature of the first substituent and is called *directive or the orientation effect***.**

benzene ring towards further substitution. These effects are called The substituent already present can increase or decrease the rate of further substitution, *i.e.,* it either activates or deactivates the *activity effects***.**

> There are two types of substituents which produce directive effect are,

> (i) Those which direct the incoming group to ortho- and parapositions simultaneously (Neglecting meta all together).

> (ii) Those which direct the incoming group to meta-position only (Neglecting ortho- and para-positions all together).

The above mechanism is followed when *S* is Para product *-OH*, - NH₂, - Cl, - Br, - I, - OR, - NR₂, - NHCOR etc. Para attack
above mechanism

In methyl or alkyl group, the $+I$ effect of the methyl group or alkyl group initiates the resonance effect.

Thus, methyl or alkyl group directs all electrophiles to ortho and para positions.

Theory of meta directing group : The substituent, *S* withdraws electrons from ortho and para positions. Thus, *m-*position becomes a point of relatively high electron density and further substitution by electrophile occurs at meta position. For example, $-NO₂$ group is a meta directing (Electron withdrawing). Its mechanism can be explained as :

All meta-directing groups have either a partial positive charge or a full positive charge on the atom directly attached to the ring.

(2) **Directive effect in disubstituted benzene**

(i) If the directive effects of two substituents reinforce, then a single product is formed.

Example :

Thus, both (CH_3, NO_2) direct further substitution to the same position (Ortho with respect to $CH₃$).

(ii) If the directing effect of two groups oppose each other strongly activating groups win over deactivating or weakly activating group. The sequence of directing power is

 $-MH_2 > -OH > -OCH_3 - > NHCOCH_3 > -C_6H_5 > CH_3 >$ meta directors

(iii) There is normally little substitution when the two groups are meta to each other. Aromatic rings with three adjacent substituents are generally prepared by same other routes.

Toluene, methyl benzene or phenyl methane

Toluene is the simplest homolouge of benzene. It was first obtained by dry distillation of tolubalsam and hence named toluene. It is commercially known as tolual.

(1) **Methods of preparation**

(i) *From benzene* [Friedel-craft's reaction] **:**

$$
\bigodot + CH_3Cl \xrightarrow{AICI_3} \bigodot + HCl
$$

 \Box Afroushelide employed may undergo an isomeric change

$$
C_6H_6 + \frac{CICH_2CH_2CH_3 \xrightarrow{A/CI_3} \xrightarrow{A/CI_3} C_6H_5CH \begin{matrix} CH_3 \\ CH_3 \\ CH_4 \end{matrix}}{C_{H_3} + HCl}
$$
\n
$$
\xrightarrow[\text{Isopropylbenzen(65-70%)}]{\text{Lopropylbenzen(65-70%)}}
$$

*CH*³

 \Box Catalysts can be used in place of anhydrous A/CI_3 are,

 $AICI_3 > SbCl_3 > SnCl_4 > BF_3 > ZnCl_2 > HgCl_2$ (ii) *Wurtz fitting reaction* **:**

$$
\left\langle \bigcirc \right\rangle \rightarrow Br+2Na+BrCH_3 \xrightarrow{\text{Ether}} \left\langle \bigcirc \right\rangle \rightarrow CH_3+2Nabr
$$

(iii) *Decarboxylation* **:** Bromobenzene Methyl bromide Toluene

$$
C_6H_4 \n\nCOONa + NaOH \n\n $C_6H_5CH_3 + Na_2CO_3$ \n $C_7O_7 \n\nSolution\n\nSolution\n\nSolution\n\n $C_8H_5CH_3 + Na_2CO_3$ \n $C_9O_3$$
$$

(iv) From crossol :
\n
$$
CH_3
$$

\n CH_3
\n CH_2
\n CH_3
\n CH_2
\n CH_3
\n CH_2

(v) *From toluene sulphonic acid* **:**

Phenyl magnesium bromide
(Viii) **Commercial preparation**

From coal tar : The main source of commercial production of toluene is the light oil fraction of coal-tar. The light oil fraction is washed with conc. H_2SO_4 to remove the bases, then with $NaOH$ to remove acidic substances and finally with water. It is subjected to fractional distillation. The vapours collected between $80 - 110^{\circ}C$ is 90% benzol which contains $70 - 80$ % benzene and $14 - 24$ % toluene. 90% benzol is again distilled and the portion distilling between $108 - 110^{\circ}C$ is collected as toluene.

(2) **Physical properties** *n*-Heptane

(i) It is a colourless mobile liquid having characteristic aromatic odour.

(ii) It is lighter than water (*sp. gr.* 0.867 at 20° *C*).

(iii) It is insoluble in water but miscible with alcohol and ether in all proportions.

(iv) Its vapours are inflammable. It boils at 110°*C* and freezes at –96°*C*.

(v) It is a good solvent for many organic compounds.

(vi) It is a weak polar compound having dipole moment 0.4*D*.

(3) **Chemical properties :** Toluene shows the behaviour of both an alipatic and an aromatic compound.

(i) *Electrophilic substitution reactions* **:** Aromatic character (More reactive than benzene) due to electron releasing nature of methyl group.

 \Box *E*⁺ may be \overrightarrow{C} , $N\overrightarrow{O_2}$, SO_3 *H* etc.

(ii*) Reactions of side chain*

(a) *Side chain halogenation* :

Benzyl chloride on hydrolysis with aqueons reaustic soda forms benzyl alcohol.

$$
C_6H_5CH_2Cl + NaOH \longrightarrow C_6H_5CH_2OH + NaCl
$$

(Benzylchloride)

Benzal chloride on hydrolysis forms benzaldehyde.

$$
C_6H_5CHCl_2 + 2NaOH \longrightarrow C_6H_5CH(OH)_{2} + 2NaCl
$$

\n
$$
C_6H_5CHO + H_2O
$$

Benzo trichloride on hydrolysis forms benzoic acid. \mathcal{C}_6 H_5 CCl_3 $+$ 3 $NaOH$ \longrightarrow \mathcal{C}_6 H_5 $\mathcal{O}(H)_3$ $+$ 3 $NaCl$ (Benzotrichloride)

$$
C_6H_5\overset{\star}{C}OOH+H_2O
$$

 \overline{C}

(b) *Oxidation* :

 \bullet With hot acidic KMnO₄:

$$
\begin{array}{ccc}\nCH_3 & & & COOH \\
\hline\n\end{array}
$$

With acidic manganese or chromyl chloride (Etardsreaction) : Toluene Benzoic acid

CH₃
\n
$$
+2[O]
$$
 $\xrightarrow{CrO_2C_2}$ \xrightarrow{CHO}
\nToluene
\nBenzaldehyde

 \Box All alkyl benzenes on oxidation with hot acidic $KMnO_4$ or

 N_a , Cr_2O_7 form benzoic acid. The length of the side chain does not matter.

(c) *Hydrogenation* : (d) *Combustion* : $C_6H_5CH_3 + 9O_2$ – (e) *Ozonolysis* : *O* $O \rightarrow O$ *CH C C O* $O \neq O$ *C CH* $\mathcal{C}H\diagdown_{\mathcal{O}}$ *O O* Triozonide Zn *HOH HC H C* $HC \ge$ *CH C H* $+3O_3$ -Toluene *Na / liquid NH³ – C2H5OH* Birch reduction *R* exane $+3H_2$ *R* Alkyl benzene

$$
CH3-C=O
$$

\n
$$
H-C=O
$$

\n
$$
CHO
$$

\n
$$
H-C=O
$$

\n
$$
CHO
$$

(i) In the manufacture of benzyl chloride, benzal chloride, benzyl alcohol, benzaldehyde, benzoic acid, saccharin, etc.

(ii) In the manufacture of trinitrotoluene (TNT), a highly explosive substance.

- (iii) As an industrial solvent and in drycleaning.
- (iv) As a petrol substitute.

(4) **Uses**

(v) In the manufacture of certain dyes and drugs.

Properties: It is pale yellow crystalline solid (M.P. = 81 °C).

Uses : \bullet It is used as an explosive in shells, bombs and torpedoes under the name trotyl.

 When mixed with 80% ammonium nitrate it forms the explosive **amatol**.

 TNT is also used as a mixture of aluminium nitrate, alumina and charcoal under the name **ammonal**.

T.N.B. (Tri-nitro benzene)

Preparation :

Properties and uses: It is colourless solid $(M.P. = \frac{T.N.B.}{122^{\circ}C}$. It is more explosive than T.N.T. and used for making explosive.

Xylenes (Dimethyl benzene) *C***6***H***4(***CH***3)²**

These are produced along with benzene, toluene and ethylbenzene when aromatisation of $C_6 - C_8$ fraction of petroleum naphtha is done. The xylenes are isolated from the resulting mixture (BTX) by fractional distillation.

These can be prepared by Wurtz – Fittig reaction. A mixture of bromotoluene and methylbromide is treated with sodium in dry ethereal solution to form the desired xylene.

These can also be obtained by Friedel – craft's synthesis, *p*-Bromotoluene *p*-Xylene

m-Xylene can be obtained from mesitylene.

Xylenes are colourless liquids having characteristic odour. The boiling points of three isomers are,

o-Xylene=144°*C*; *m*-Xylene=139°*C*; *p*-Xylene=138°*C*.

Xylenes undergo electrophilic substitution reactions in the same manner as toluene. Upon oxidation with $K M nO_4$ or $K_2 Cr_2O_7$, Xylenes form corresponding dicarboxylic acids.

Xylenes are used in the manufacture of lacquers and as solvent for rubber. *o*-Xylene is used for the manufacture of phthalic anhydride. *COOH* Terephthalic acid

Ethyl benzene $(C_6H_5C_2H_5)$

It can be prepared by the following reactions,

(1) **By Wurtz-Fittig reaction :**

 $C_6H_5Br + 2Na + BrC_2H_5 \longrightarrow C_6H_5C_2H_5 + 2NaBr$

(2) **By Friedel-craft's reaction :**

$$
C_6H_5H + BrC_2H_5 \xrightarrow{A/C_3} C_6H_5C_2H_5 + HBr
$$

(3) **By catalytic reduction of styrene :**

$$
C_6H_5CH = CH_2 + H_2 \longrightarrow C_6H_5CH_2CH_3
$$

(4) **By alkyl benzene synthesis :**

$$
C_6H_5H + H_2C = CH_2 \xrightarrow{A/C_3, HCl} C_6H_5CH_2CH_3
$$

It undergoes electrophilic substitution reactions in the same way as toluene. When oxidised with dil. $HNO₃$ or alkaline $KMD₄$ or chromic acid it forms benzoic acid.

$$
C_6H_5C_2H_5 \xrightarrow{\qquad [O]\qquad} C_6H_5COOH
$$

Styrene (*C6H5CH=CH2***)**

It is present in storax balsam and coal-tar in traces.

(i) *Dehydrogenation of side chain of ethylbenzene* **:**

(ii) *Decarboxylation of cinnamic acid* **:** This is the laboratory preparation and involves heating of cinnamic acid with a small amount of quinol.

$$
C_6H_5CH = CHCOOH \xrightarrow{Quinol} C_6H_5CH = CH_2 + CO_2
$$

(iii) *Dehydration of* **1***-phenyl ethanol with H***2***SO***⁴ :** $C_6 H_5$ CHOHCH₃ $\frac{H_2SO_4}{-H_2O}$ \rightarrow $C_6 H_5CH = CH_2$

(iv) *Dehydration of* **2***-phenyl ethanol with ZnCl***² :** $C_6 H_5 CH_2 CH_2 OH \xrightarrow{ZnCl_2, \text{heat}} C_6 H_5 CH = CH_2$ also important since many of the substitute).

(v) *Dehydrohalogenation of* **1***-phenyl-***1***-chloro ethane* **:** On heating with alcoholic potassium hydroxide, a molecule of hydrogen chloride is eliminated by the chloroderivative.

$$
C_6H_5CHCICH_3 \xrightarrow{\text{Alc. KOH}} C_6H_5CH = CH_2
$$

(2) **Properties :** It is a colourless liquid, boiling point 145°*C*. On keeping, it gradually changes into a solid polymer called **metastyrene**. The polymerisation is rapid in sunlight or when treated with sodium. It shows properties of benzene ring (Electrophilic substitution) and unsaturated side chain (Electrophilic addition). However, the side chain double bond is more susceptible to electrophilic attack as compared to benzene ring.

At lower temperature and pressure, it reacts with hydrogen to produce ethylbenzene and at higher temperature and pressure, it is converted into ethyl cyclohexane.

With bromine, it gives the dibromide.

Styrene.
Halogen acids add to the side chain. Styrene dibromide

$$
C_6H_5CH = CH_2 + HX \longrightarrow C_6H_5CHXCH_3
$$

Preparation of ring substituted styrenes is not done by direct halogenation but through indirect route.

When oxidised under drastic conditions, the side chain is completely oxidised to a carboxyl group.

In presence of peroxides, styrene undergoes free radical polymerisation resulting in the formation of polystyrene – an industrially important plastic.

$$
nC_6H_5CH = CH_2 \xrightarrow{\text{Peroxide}} \left[-\frac{CH - CH_2}{C_6H_5}\right]
$$

Co-polymers of styrene with butadiene and other substances are also important since many of them are industrially useful products such as *SBR* (A rubber substitute).

Bi-phenyl $(C_6H_5 - C_6H_5)$

 $C_6 H_5$ CHClCH₃ $\frac{\text{Alc. } KOH}{\text{Heat}}$ $C_6 H_5$ CH = CH₂ hydrocarbon in which two benzene rings are directly linked to each other. It occurs in coal-tar. It is the simplest example of an aromatic other.

(1) **Methods of formation**

(i) *Fittig reaction* **:** It consists heating of an ethereal solution of bromobenzene with metallic sodium.

$$
\left\langle \bigcirc \right\rangle Br + 2Na + Br \left\langle \bigcirc \right\rangle \longrightarrow \left\langle \bigcirc \right\rangle + 2NaBr
$$

(ii) *Ullmann biaryl synthesis* **:** Iodobenzene, on heating with copper in a sealed tube, forms biphenyl. The reaction is facilitated if a strong electron withdrawing group is present in ortho or para position.

$$
\left\langle \bigodot \right\rangle I+2Cu+I-\left\langle \bigodot \right\rangle \longrightarrow \left\langle \bigodot \right\rangle \left\langle \bigodot \right\rangle +2CuI
$$

(iii) *Grignard reaction* **:** Phenyl magnesium bromide reacts with bromo benzene in presence of $CoCl₂$.

 $C_6H_5CH = CH_2 + HX \longrightarrow C_6H_5CHXCH_3$ orientating effect. But, it has been experimentally shown that presence (2) **Properties :** It is a colourless solid, melting point 71°*C*. It undergoes usual electrophilic substitution reactions. Since aryl groups are electron withdrawing , they should have deactivating and *m*of one benzene ring activates the other for electrophilic substitution and directs the incoming group to *o-* and *p-* positions. It has been shown that monosubstitution in the bi-phenyl results in the formation of para isomer as the major product.

Another special feature of the biphenyl is the behaviour towards second substitution in a monosubstituted biphenyl. The second substituent invariably enters the unsubstituted ring in the ortho and para position no matter what is the nature of substituent already present.

 Octane number may be less than zero (*e.g.*, *n*-Nonane has an octane number-45) and higher than 100 (*e.g.*, Triptane or 2, 3, 3- Trimethylbutane has an octane number of 124).

To avoid lead pollution, a new compound cyclopentadienyl manganese carbonyl

$$
\Rightarrow -Mn \leftarrow \begin{matrix} CO \\ CO \\ CO \end{matrix}
$$

(called as AK-33-X) is used a Contiknock now a days in developed countries (unleaded pertol).

Acetylene has a garlic odour when impure due to impurities of phosphine and hydrogen sulphide.

 Fluorination is a violent reaction and can be controlled by diluting fluorine with nitrogen.

The relative acidic character of water, alcohols acetylene, ammonia, ethylene and ethane follows the order :

 $H_2O > ROH > HC \equiv CH > CH_3 > CH_2 = CH_2 > CH_3 - CH_3$.

Obviously, the basic character of their conjugate bases follows the reverse order, *i.e.*,

 $CH_3CH_2^- > CH_2 = CH^- > NH_2^- > HC \equiv C^- > RO^- > HO^-$.

Z Wilkinson's catalyst : (Triphenylphosphine) rhodium, (*PPh*₃)₃ *RhCl* is called wilkinson's catalyst. It reduces alkenes and alkynes while other common functional groups such as $C=O$, NO_2 and $C=$ *N* remain unaffected,

 \blacktriangleright The order of reactivity of primary (1°), secondary (2°) and tertiary (3^o) hydrogens in alkanes follows the sequence : $3^{\circ} > 2^{\circ} >$ 1 o .