

Chapter 28 Carboxylic acids and Their derivatives

Carboxylic Acids

Carboxylic acids are the compounds containing the carboxyl

functional group $\begin{vmatrix} -C & OH \\ \parallel \\ O \\ \end{vmatrix}$

The carboxyl group is made up of carbonyl (>C=O) and hydroxyl (–OH) group.

Classification

(1) Carboxylic acids are classified as monocarboxylic acids, dicarboxylic acids, tricarboxylic acids etc. depending on the number of -COOH groups present in the molecule.

 $\begin{array}{c} CH_3COOH & CH_2COOH & CH_2COOH \\ & & & \\ & CH_2COOH & CHCOOH \\ & & & & \\ & & & \\ & & & \\ & & & \\ & & & \\ & & & \\ & & & \\ & & & \\ & & &$

Monocarboxylic acid Dicarboxylic acid Tricarboxylic acid

(2) Monocarboxylic acids of aliphatic series are commonly known as fatty acids such as palmitic acid ($C_{15}H_{31}COOH$) and stearic acid ($C_{17}H_{35}COOH$).

(3) The general formula for monocarboxylic acids is $C_n H_{2n+1} COOH$ or $C_n H_{2n} O_2$. Where n = number of carbon atoms.

(4) The carboxylic acids may be aliphatic or aromatic depending upon whether -COOH group is attached to aliphatic alkyl chain or aryl group respectively.

Methods of preparation of monocarboxylic acid

(1) By oxidation of alcohols, aldehydes and ketones

$$\begin{array}{c} RCH_2OH \xrightarrow{[O]} RCHO \xrightarrow{[O]} RCOOH \\ alcohol & K_2Cr_2O_7 & K_2Cr_2O_7 \\ \end{array} \xrightarrow{(O)} Carboxylic acid \\ Carboxylic acid$$

□ Aldehyde can be oxidized to carboxylic acid with mild oxidising agents such as ammonical silver nitrate solution $[Ag_2O \text{ or } Ag(NH_3)_2^+OH^-]$

□ Methanoic acid can not be prepared by oxidation method.

 \Box Ketones can be oxidized under drastic conditions using strong oxidising agent like $K_2 Cr_2 O_7$.

☐ Methyl ketones can also be converted to carboxylic acid through the haloform reaction.

$$\begin{array}{c} R- \begin{array}{c} C- CH_3 + 3I_2 + 3NaOH \xrightarrow{\Lambda} \\ \parallel \\ O \\ R- \begin{array}{c} C- OH + CHI_3 + 3NaI + 3H_2O \\ \parallel \\ O \\ \end{array} \end{array}$$

(2) By Hydrolysis of nitriles, ester, anhydrides and acid chloride

(i) Hydrolysis of nitriles

$$R - C \equiv N + HOH \xrightarrow{HCI}_{\text{or NaOH}} \left[R - C \xrightarrow{OH}_{NH} \right] \xrightarrow{\text{Rearrangement}}_{R - C \xleftarrow{O}_{NH_2}} \frac{H_2O}{HCI} RCOOH + NH_4CI$$

(ii) Hydrolysis of Esters

$$\begin{array}{c} RCOOR + HOH \xrightarrow{HCI} RCOOH + ROH \\ Ester & OH^{-} \\ Acid \\ \end{array} \xrightarrow{RCOOH + ROH \\ Alcohol \\ \end{array}$$

(iii) Hydrolysis of Anhydrides

$$CH_{3} - C \rightarrow O + HOH \xrightarrow{H^{+} / OH^{-}} 2CH_{3}COOH$$

$$CH_{3} - C \rightarrow O + HOH \xrightarrow{H^{+} / OH^{-}} 2CH_{3}COOH$$

Ethanoicanhydride

(iv) Hydrolysis of acid chloride and nitro alkane

$$R-CH_2-NO_2 \xrightarrow{85\%H_2SO_4} RCOOH$$

(v) Hydrolysis of Trihalogen :

$$R - C \stackrel{X}{\longrightarrow} X + 3 \text{ NaOH} \rightarrow \left[\begin{array}{c} R - C \stackrel{OH}{\longrightarrow} \\ OH \\ OH \end{array} \right] \stackrel{-H_2O}{\longrightarrow} \\ R - C \stackrel{O}{\longleftarrow} OH^+ 3 \text{ NaX}$$

(3) From Grignard Reagent

$$O = \overleftarrow{C} = O + \overrightarrow{RMgX} \xrightarrow{\text{Dry ether}} R - \overrightarrow{C} - OMgX$$

arbon dioxide $\overrightarrow{Grignard reagent} \xrightarrow{H^+ \mid H_2O} RCOOH + MdOH$

(4) From Alkene or Hydro-carboxy-addition (koch reaction)

$$CH_2 = CH_2 + CO + H_2O \xrightarrow[500-1000 atm]{H_3PO_4} CH_3CH_2COOH$$

(5) Special methods

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(i) Carboxylation of sodium alkoxide

$$\begin{array}{c} \textit{RONa+ CO} \rightarrow \textit{RCOONa} \xrightarrow{\textit{HCI}} \textit{RCOOH} \\ \textit{Sod.alkoxide} & \textit{Sod.salt} & \textit{Acid} \end{array}$$

(ii) Action of heat on dicarboxylic acid

$$\begin{array}{c} R-CH \xrightarrow{COOH} \xrightarrow{-CO_2} \\ COOH \xrightarrow{} heat \xrightarrow{} R-CH_2COOH \\ heat \xrightarrow{} Monocarboxylic acid \end{array}$$

(iii) From acetoacetic ester

$$\begin{array}{c} CH_{3}CO \mid CHRCO \mid OC_{2}H_{5} & \xrightarrow{} Hydrolysis \\ OH \mid H & OH \mid H \end{array} \xrightarrow{} Hydrolysis + RCH_{2}COOH + C_{2}H_{5}OH \end{array}$$

(iv) Oxidation of alkene and alkyne

$$\begin{array}{c} RCH = CHR \\ Alkene \\ KMnO_{4} \end{array} \xrightarrow{[O]} RCOOH + RCOOH \\ \hline \end{array}$$

$$R - \underset{\text{Alkyne}}{C \equiv C - R} \xrightarrow{(i)O_3} R - COOH + RCOOH$$

(v) The Arndt-Eistert synthesis

$$\begin{array}{ccc} R-C-CI + CH_2N_2 \rightarrow R-C-CHN_2 \xrightarrow{H_2O} \\ \parallel \\ O & O \\ \end{array} \xrightarrow{H_2O} \\ Ag_2O \\ R-CH_2 - COOH \end{array}$$

(vi) From acid amides

$$\begin{array}{c} RCONH_2 + H_2O \xrightarrow[]{\text{Acid}} RCOOH + NH_3\\ \text{Amide} \end{array} \xrightarrow[]{\text{Amide}} HNO_2 \xrightarrow[]{\text{Acid}} RCOOH + N_2 + H_2O\\ \text{Amide} \text{Nitrous acid} \end{array}$$

Physical properties of monocarboxylic acids

(1) **Physical state :** The first three members (upto 3 carbon atoms) are colourless, pungent smelling liquids. The next six members are oily liquids having unpleasant smell. The higher members are colourless and odourless waxy solids.

(2) **Solubility :** The lower members of the aliphatic carboxylic acid family (upto C_4) are highly soluble in water. The solubility decreases with the increase in the size of the alkyl group. All

carboxylic acids are soluble in alcohol, ether and benzene etc.

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 \Box The solubility of lower members of carboxylic acids is due to the formation of hydrogen bonds between the – *COOH* group and water molecules.

□ Acetic acid exists in the solution in dimer form due to intermolecular hydrogen bonding. The observed molecular mass of acetic acid is 120 instead of 60.

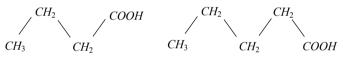
(3) Melting point

(i) The melting points of carboxylic acids donot vary smoothly from one member to another.

(ii) The melting point of the acids having even number of carbon atoms are higher than those containing an odd number immediately above and below them.

(iii) The acids with even number of carbon atoms have the – COOH group and the terminal – CH_3 group on the opposite side of the carbon chain.

(iv) In the case of odd numbers, the two groups lie on the same side of the chain.

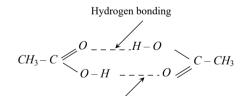


the two terminal groups lie on the opposite sides of the chain

the two terminal groups lie on the same side of the chain

When the terminal groups lie on the opposite sides the molecules fit into each other more closely. More effective packing of the molecule in the lattice. Therefore, results into higher melting point.

(4) **Boiling point :** Boiling point of carboxylic acids increase regularly with increase of molecular mass. Boiling points of carboxylic acids are higher than those of alcohols of same molecular mass. This is due to intermolecular hydrogen bonding between two acid molecules.



Hydrogen bonding Acetic acid dimer

Acidic nature of monocarboxylic acids

(1) Cause of acidic nature

(i) A molecule of carboxylic acid can be represented as a resonance hybrid of the following structures.

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(ii) Due to electron deficiency on oxygen atom of the hydroxyl group (Structure II), their is a displacement of electron pair of O-H bond toward the oxygen atom. This facilitate the release of hydrogen as proton (H^+).

$$R - C \stackrel{O \oplus}{=} O \leftarrow H \leftrightarrow R - C \stackrel{O}{\ll} O \leftrightarrow R - C \stackrel{O}{\ll} O = R - C \stackrel{O}{\ll} O \stackrel{1.27 A^{\circ}}{1.27 A^{\circ}}$$
Resonancefyrbrid

(iii) The resulting carboxylate ion also stabilized by resonance (As negative charge is dispersed on both the oxygen atom). This enhance the stability of carboxylate anion and make it weaker base or strong acid.

(2) Effect of substituent on acidic nature

(i) An electron withdrawing substituent (-I effect) stabilizes the anion by dispersing the negative charge and therefore increases the acidity.

(ii) An electron releasing substituent (+ I effect) stabilizes negative charge on the anion resulting in the decrease of stability and thus decreased the acidity of acid.

Electron with drawing nature of halogen : F > Cl > Br > I

Thus, the acidic strength decreases in the order :

 $FCH_2COOH > CICH_2COOH > BrCH_2COOH > ICH_2COOH$

similarly :

$$CCl_3COOH > CHCl_2COOH > CH_2CICOOH > CH_3COOH$$

(iii) Inductive effect is stronger at α -position than β -position similarly at β -position it is more stronger than at γ -position

Example:

$$CH_{3} - CH_{2} - CH_{-}COOH > CH_{3} - CH_{-}CH_{2} - COOH$$

$$| CI = CI$$

$$> CH_{2} - CH_{2} - CH_{2} - COOH$$

$$| CI = CI$$

(iv) Relative acid strength in different compounds

 $RCOOH > HOH > ROH > HC \equiv CH > NH_3 > RH$

□ Greater the value of K_a or lesser the value of pK_a stronger is the acid, i.e. $pK_a = -\log K_a$

 \Box Acidic nature (K_a) α 1/molecular weight

$$HCOOH > CH_3COOH > C_2H_5COOH$$

 K_2 Value 17.7×10^{-5} 1.75×10^{-5} 1.3×10^{-5}

The formic acid is strongest of all fatty acids.

□ Acetic acid is less weak acid than sulphuric acid due to less degree of ionisation.

Chemical properties of monocarboxylic acids

(1) Reaction involving removal of proton from -OH group

(i) *Action with blue litmus* : All carboxylic acids turn blue litmus red.

(ii) Reaction with metals

$$2CH_{3}COOH + 2Na \rightarrow 2CH_{3}COONa+ H_{2}$$
Sodium acetate

$$2CH_{3}COOH + Zn \rightarrow (CH_{3}COO)_{2}Zn + H_{2}$$
Zinc acetate
(iii) Action with alkalies

$$CH_{3}COOH + NaOH \rightarrow CH_{3}COONa+ H_{2}O$$
Aceticacid
Sodium acetate
(iv) Action with carbonates and bicarbonates

$$2CH_{3}COOH + Na_{2}CO_{3} \rightarrow 2CH_{3}COONa+ CO_{2} + H_{2}O$$
Sod. acetate

$$CH_{3}COOH + NaHCO_{3} \rightarrow CH_{3}COONa+ CO_{2} + H_{2}O$$

Sod.acetate Reaction of carboxylic acid with aqueous sodium carbonates solution produces bricks effervescence. However most phenols do not produce affervescence. Therefore, this reaction may be used to

produce effervescence. Therefore, this reaction may be used to distinguish between carboxylic acids and phenols. (2) Reaction involving replacement of -OH group

(i) Formation of acid chloride

$$CH_3COOH+PCI_5 \rightarrow 3CH_3COCI+POCI_3 + HCI$$

Aceticacid Acetylchloride
 $3CH_3COOH+PCI_3 \rightarrow 3CH_3COCI+H_3PO_3$
Aceticacid Acetylchloride
 $CH_3COOH+SOCI_2 \rightarrow CH_3COCI+SO_2 + HCI$
Aceticacid Acetylchloride
(ii) Formation of esters (Esterification)
 $CH_3COOH+H = OC_2H_5 = Conc.H_2SO_4 = CH_3COOC_2H_5 + H_2O$
 $CH_3COOC_2H_5 + H_2O$

Ethyl acetate (Fruity smelling)

(a) The reaction is shifted to the right by using excess of alcohol or removal of water by distillation.

(b) The reactivity of alcohol towards esterification.

tert-alcohol < *sec*-alcohol < *pri*-alcohol < methyl alcohol

(c) The acidic strength of carboxylic acid plays only a minor role.

$$R_3CCOOH < R_2CHCOOH < RCH_2COOH < CH_3COOH < HCOOH$$

When **methanol** is taken in place of **ethanol**. then reaction is called **trans esterification**.

(iv) Formation of amides

$$\begin{array}{c} CH_3 COOH+ \ NH_3 \xrightarrow{\text{heat}} CH_3 COONH_4 \xrightarrow{\Delta} \\ \text{Aceticacid} & \text{Amm. acetate} \end{array}$$

(v) Formation of acid anhydrides

(vi) Reaction with organo-metallic reagents

$$\begin{array}{c} R \ CH_2 \ MgBr + R \ COOH \longrightarrow R \ CH_3 + R \ COOMgBr \\ Alkane \end{array}$$

(3) Reaction involving carbonyl (>C = O) group:



$$\begin{array}{c} \textit{Reduction}: \ R-C-OH \xrightarrow{\text{LiA}/H_4} R-CH_2 - OH \\ \parallel \\ O \end{array}$$

Carboxylic acid are difficult to reduce either by catalytic hydrogenation or Na/C_2H_5OH

(4) Reaction involving attack of carboxylic group (-*COOH*)

(i) **Decarboxylation** :
$$R - \overset{\bigcup}{C} - OH \xrightarrow{(-CO_2)} R - H$$

When anhydrous alkali salt of fatty acid is heated with sodalime then :

 \Box When sodium formate is heated with sodalime H_2 is evolved. (Exception)

$$HCOONa + NaOH \longrightarrow H_2 + Na_2CO_3$$

(ii) Heating of calcium salts

$$(RCOO)_2 Ca \xrightarrow{\text{heat}} RCOR CaCO_3$$
Sodium salt

(iii) Electrolysis : (Kolbe's synthesis)

$$RCOONa \Rightarrow RCOO^{-} + Na^{+}$$

At anode
$$2RCOO^- \rightarrow R - R + 2CO_2 + 2e^-$$

At cathode
$$2Na^+ + 2e^- \rightarrow 2Na \xrightarrow{2H_2O} 2NaOH + H_2$$

 $2CH_3COOK + 2H_2O$ Electrolysis Potassiumacetate

$$\begin{array}{c} \mathcal{CH}_3 - \mathcal{CH}_3 + 2\mathcal{CO}_2 + 2\mathcal{KOH} + \mathcal{H}_2 \\ \text{Ethane} \end{array}$$

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(iv) Formation of Alkyl halide (Hunsdiecker's reaction)

$$\begin{array}{c} CH_3 COOAg+Br_2 \xrightarrow{\text{heat}} CH_3 Br + AgBr + CO_2\\ \text{Silver acetate} \end{array} \xrightarrow{CCl_4} Methyl bromide \end{array}$$

□ In Hunsdiecker reaction, one carbon atom less alkyl halide is formed from acid salt.

(v) Formation of amines (Schmidt reaction)

$$\begin{array}{ccc} RCOOH+ & N_3H & \xrightarrow{H_2SO_4(conc)} & RNH_2 + CO_2 + N_2 \\ & & & & \\ Acid & & & \\ Hydrazoic & & & \\ acid & & & \\ amine & & \\ \end{array}$$

In Schmidt reaction, one carbon less product is formed. (vi) *Complete reduction*

$$\begin{array}{c} CH_3 COOH + 6HI \xrightarrow{P} CH_3 CH_3 + 2H_2 O + 3I_2 \\ \text{Aceticacid} & \text{Ethane} \end{array}$$

In the above reaction, the -COOH group is reduced to a CH_3

(5) Reaction involving hydrogen of α-carbon Halogenation

(i) In presence of U.V. light

group.

$$\begin{array}{c} H \\ - C \\$$

(ii) In presence of Red P and diffused light [Hell Volhardzelinsky reaction] Carboxylic acid having an α -hydrogen react with Cl_2 or Br_2 in the presence of a small amount of red phosphorus to give chloro acetic acid. The reaction is known as **Hell Volhard-zelinsky reaction**.

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$$\begin{array}{c} CH_{3}COOH \xrightarrow{Cl_{2}, \text{red } P_{4}} \rightarrow CICH_{2}COOH \xrightarrow{Cl_{2}, \text{red } P_{4}} \rightarrow \\ \text{Aceticacid} \xrightarrow{-HCl} \text{Chloro aceticacid} \xrightarrow{-HCl} Cl_{2}COOH \\ \hline Cl_{2}CHCOOH \xrightarrow{Cl_{2}, \text{red } P_{4}} \rightarrow Cl_{3}CCOOH \\ \hline Dichloro aceticacid \xrightarrow{-HCl} \text{Trichloro aceticacid} \end{array}$$

Individual members of monocarboxylic acids

Formic Acid or Methanoic acid (HCOOH)

Formic acid is the first member of monocarboxylic acids series. It occurs in the sting of bees, wasps, red ants, stinging nettles. and fruits. In traces it is present in perspiration, urine, blood and in caterpillar's.

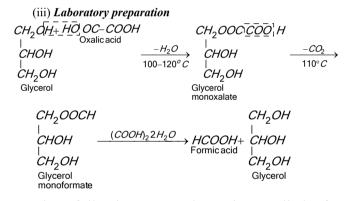
(1) Methods of preparation

$$CH_3OH + O_2 \xrightarrow{Pt} HCOOH + H_2O$$
Formic acid

(ii) *Hydrolysis of hydrocyanic acid :* Formic acid is formed by the hydrolysis of *HCN* with acids or alkalies.

$$HCN + 2H_2O \longrightarrow HCOOH + NH_3;$$

$$HCN + H_2O \xrightarrow{NaOH} HCOONa + NH_3$$



The following procedure is applied for obtaining anhydrous formic acid.

$$2HCOOH + PbCO_3 \rightarrow (HCOO)_2 Pb + CO_2 + H_2O;$$

Lead formate
$$(HCOO)_2 Pb + H_2S \rightarrow PbS + 2HCOOH$$

ppt. Formic acid

(iv) **Industrial preparation** : Formic acid is prepared on industrial scale by heating sodium hydroxide with carbon monoxide at $210^{\circ}C$ under a pressure of about 10 atmospheres.

$$CO+ NaOH \xrightarrow{\Delta} HCOONa$$

$$210^{\circ}C, 10 \text{ atm} \text{ Sodium formate}$$

Sodium formate thus formed is distilled with sodium hydrogen sulphate, when anhydrous formic acid distils over.

 $HCOONa + NaHSO_4 \rightarrow HCOOH + Na_2SO_4$

(2) Physical properties

(i) It is a colourless pungent smelling liquid.

(ii) It melts at $8.4^{\circ}C$ and boils at $100.5^{\circ}C$.

(iii) It is miscible with water, alcohol and ether. It forms azeotropic mixture with water.

(iv) It is strongly corrosive and cause blisters on skin.

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(v) It exists in aqueous solution as a dimer involving hydrogen bonding.

(3) Uses : Formic acid is used.

(i) In the laboratory for preparation of carbon monoxide.

(ii) In the preservation of fruits.

(iii) In textile dyeing and finishing.

(iv) In leather tanning.

(v) As coagulating agent for rubber latex.

(vi) As an antiseptic and in the treatment of gout.

(vii) In the manufacture of plastics, water proofing compounds.

(viii) In electroplating to give proper deposit of metals.

(ix) In the preparation of nickel formate which is used as a catalyst in the hydrogenation of oils.

(x) As a reducing agent.

(xi) In the manufacture of oxalic acid.

Acetic Acid (Ethanoic Acid) (CH₃COOH)

Acetic acid is the oldest known fatty acid. It is the chief constituent of vinegar and hence its name (Latin acetum = vinegar)

(1) Preparation

(i) By oxidation of acetaldehyde (Laboratory-preparation)

 $CH_{3}CHO \xrightarrow{Na_{2}Cr_{2}O_{7}} H_{3}COOH$

(ii) By hydrolysis of methyl cyanide with acid

 $CH_3CN + 2H_2O \xrightarrow{HCI} CH_3COOH + NH_3$

(iii) By Grignard reagent

$$CH_{3}MgBr + CO_{2} \rightarrow CH_{3} - \overset{\parallel}{C} - OMgBr \xrightarrow{H_{2}O/H^{+}}$$

$$\begin{pmatrix} O \\ CH_{3} - \overset{\parallel}{C} - OH \end{pmatrix}$$

(iv) By hydrolysis of acetyl chloride, acetic anhydride or acetamide and ester

(a)
$$CH_3COOC_2H_5 + H_2O \xrightarrow{H_2SO_4(\text{conc.})}$$

Ester

$$CH_3COOH + C_2H_5OH$$

(b)
$$CH_3COCI + H_2O \xrightarrow{\text{dil.} HCI} CH_3COOH + HCI$$

acetylchloride

(c)
$$(CH_3CO)_2O + H_2O \xrightarrow{\text{dil.}HCI} 2CH_3COOH$$

(v) Manufacture of acetic acid

(a) From ethyl alcohol (Quick vinegar process) : Vinegar is 6-10% aqueous solution of acetic acid. It is obtained by fermentation of liquors containing 12 to 15% ethyl alcohol. Fermentation is done by *Bacterium Mycoderma aceti* in presence of air at 30-35°C. The process is termed **acetous fermentation**.

$$\begin{array}{c} CH_3CH_2OH + O_2 & \xrightarrow{Mycodermaaceti} & CH_3COOH + H_2O\\ \hline \text{Ethyl alcohol} & & \text{Bacteria} & \text{Aceticacid} \end{array}$$

It is a slow process and takes about 8 to 10 days for completion. In this process, the following precautions are necessary: • The concentration of the ethyl alcohol should not be more than 15%, otherwise the bacteria becomes inactive.

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• The supply of air should be regulated. With less air the oxidation takes place only upto acetaldehyde stage while with excess of air, the acid is oxidised to CO_2 and water.

• The flow of alcohol is so regulated that temperature does not exceed 35°*C*, which is the optimum temperature for bacterial growth.

Acetic acid can be obtained from vinegar with the help of lime. The calcium acetate crystallised from the solution is distilled with concentrated sulphuric acid when pure acetic acid distils over.

(b) From acetylene : Acetylene is first converted into acetaldehyde by passing through 40% sulphuric acid at $60^{\circ}C$ in presence of 1% HgSO₄ (catalyst).

$$CH = CH + H_2O \xrightarrow{H_2SO_4(\text{dil.})} CH_3CHO$$
Acetylene $HgSO_4$ Acetaldehyde

The acetaldehyde is oxidised to acetic acid by passing a mixture of acetaldehyde vapour and air over manganous acetate at $70^{\circ}C$.

$$2CH_3CHO + O_2 \xrightarrow{\text{Manganousacetate}} 2CH_3COOH$$

$$70^{\circ}C$$

□ Acetylene required for this purpose is obtained by action of water on calcium carbide.

$$CaC_2 + 2H_2O \rightarrow Ca(OH)_2 + C_2H_2$$

The yield is very good and the strength of acid prepared is 97%. The method is also quite cheap.

(c) By the action of CO on methyl alcohol : Methyl alcohol and carbon monoxide react together under a pressure of 30 atmospheres and $200^{\circ}C$ in presence of a catalyst cobalt octacarbonyl, $Co_2(CO)_8$ to form acetic acid.

$$CH_{3}OH + CO \xrightarrow{Co_{2}(CO)_{8}} CH_{3}COOH$$

Methyl alcohol $30 atm 200^{\circ}C$ Aceticacid

(2) Physical properties

(i) At ordinary temperature, acetic acid is a colourless, corrosive liquid with a sharp pungent odour of vinegar. It has a sour taste.

(ii) Below $16.5^{\circ}C$, it solidifies as an icy mass, hence it is named glacial acetic acid.

(iii) It boils at 118°C. The high boiling point of acetic acid in comparison to alkanes, alkyl halides or alcohols of nearly same molecular masses is due to more stronger hydrogen bonding between acid molecules. This also explains dimer formation of acetic acid in vapour state.

(iv) It is miscible with water, alcohol and ether in all proportions.

(v) It is good solvent for phosphorus, sulphur, iodine and many organic compounds.

(3) Uses : It is used,

(i) As a solvent and a laboratory reagent.

(ii) As vinegar for table purpose and for manufacturing pickles.

(iii) In coagulation of rubber latex.

(iv) For making various organic compounds such as acetone, acetic anhydride, acetyl chloride, acetamide and esters.

(v) For making various useful metallic acetates, such as:

(a) Basic copper acetate which is used for making green paints.(b) *Al*, *Fe* and *Cr* acetates which are used as mordants in dyeing.

(c) Lead tetra-acetate which is a good oxidising agent.

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(d) Basic lead acetate which is used in the manufacture of white lead.

(e) Aluminium acetate which is used in the manufacture of water-proof fabrics.

(f) Alkali acetates which are used as diuretics.

| Property | Formic acid | Acetic acid |
|--|--|---|
| 1. Acidic nature, | | |
| (i) With electro-positive | Forms salts, Hydrogen is evolved. | Forms salts. Hydrogen is evolved. |
| metals | $HCOOH + Na \rightarrow HCOONa + \frac{1}{2}H_2$ | $CH_3COOH + Na \rightarrow CH_3COONa + \frac{1}{2}H_2$ |
| (ii) With bases | Forms salts. | Forms salts. |
| | $HCOOH + NaOH \rightarrow HCOONa + H_2O$ | $CH_3COOH + NaOH \rightarrow CH_3COONa + H_2O$ |
| (iii) With carbonates and | Forms salts. Carbon dioxide is evolved. | Forms salts. Carbon dioxide is evolved. |
| bicarbonates | $HCOOH + NaHCO_3 \rightarrow HCOONa + H_2O + CO_2$ | $CH_3COOH + NaHCO_3 \rightarrow$ |
| | | CH ₃ COONa+ H ₂ O+ CO ₂ |
| 2. Ester formation | Forms esters when treated with alcohols. | Forms esters when treated with alcohols. |
| | $HCOOH + C_2H_5OH \rightarrow HCOOC_2H_5 + H_2O$ | $CH_{3}COOH + C_{2}H_{5}OH \xrightarrow{H_{2}SO_{4}(conc)} \rightarrow$ |
| | | $CH_3COOC_2H_5 + H_2O$ |
| 3. Reaction with <i>PCl</i> ₅ | Forms formyl chloride which decomposes into CO and | |
| 5. Reaction with $F C I_5$ | HCl. | Forms acetyl chloride which is a stable compound. $CH_3COOH + PCI_5 \rightarrow$ |
| | $HCOOH + PCl_5 \rightarrow HCOCI(HCI + CO) + POCl_3 + HCI$ | CH ₃ COCI+ POC ₂ + HCl |
| | | |
| 4. Heating of ammonium | Forms formamide. | Forms acetamide. |
| salt | $HCOONH_4 \rightarrow HCONH_2 + H_2O$ | $CH_3COONH_4 \rightarrow CH_3CONH_2 + H_2O$ |
| 5. Heating alone | it decomposes into CO_2 and H_2 | Unaffected |
| | $HCOOH \rightarrow CO_2 + H_2$ | |
| 6. Heating with conc. | Decomposed into CO and H_2O | Unaffected |
| H_2SO_4 | $HCOOH \xrightarrow{Conc}_{H_2SO_4} CO + H_2O$ | |
| 7. Reaction with Cl_2 in presence of red P | Unaffected | Forms mono, <i>di</i> or trichloro acetic acids. |
| 8. Action of heat on salts, | | |
| (i) Calcium salt | Forms formaldehyde. | Forms acetone. |
| | $(HCOO)_2 Ca \rightarrow HCHO + CaCO_3$ | $(CH_3COO)_2Ca \rightarrow CH_3COCH_3 + CaCO_3$ |
| (ii) Sodium salt | Forms sodium oxalate. | Unaffected. |
| | $2HCOONa \xrightarrow{heat} H_2$ | |
| (iii) Sodium salt with soda- | Forms sodium carbonate and H_2 . | Forms sodium carbonate and methane. |
| lime | $HCOONa + NaOH \underline{CaO} > Na_2CO_3 + H_2$ | $CH_3COONa + NaOH \xrightarrow{CaO} CH_4 + Na_2CO_3$ |
| | | |
| 9. Electrolysis of sodium or potassium salt | It evolves hydrogen. | It forms ethane. |
| 10. On heating with P_2O_5 | Unaffected | Forms acetic anhydride. |
| | | $2CH_3COOH \xrightarrow{P_2O_5} (CH_3CO_{12}O + H_2O)$ |
| 11. Reducing nature, | | |
| (i) Tollen's reagent | Gives silver mirror or black precipitate. | Unaffected. |
| | $HCOOH + Ag_2O \rightarrow 2Ag + CO_2 + H_2O$ | |
| (ii) Fehling's solution | Gives red precipitate | Unaffected. |
| | | |
| | $HCOOH + 2CuO \rightarrow Cu_2O + CO_2 + H_2O$ | |

Table : 28.1 Comparison of Formic Acid and Acetic Acid

| | $HgCl_2 \rightarrow Hg_2Cl_2 \rightarrow 2Hg$ | |
|---|--|------------------|
| (iv) Acidified <i>KMnO</i> ₄ | Decolourises | Unaffected. |
| 12. Acid (neutral solution)+ NaHSO₃ + Sodiumnitroprusside. | Greenish blue colour. | Unaffected. |
| 13. Acid (neutral solution)+ neutral ferric chloride | Red colour which changes to brown ppt. on heating. | Wine red colour. |

Interconversions

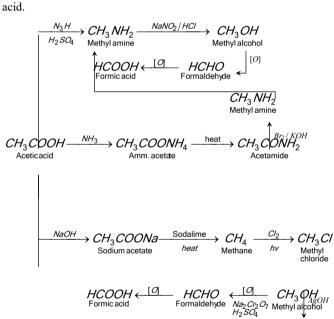
(1) Ascent of series : Conversion of formic acid into acetic acid.

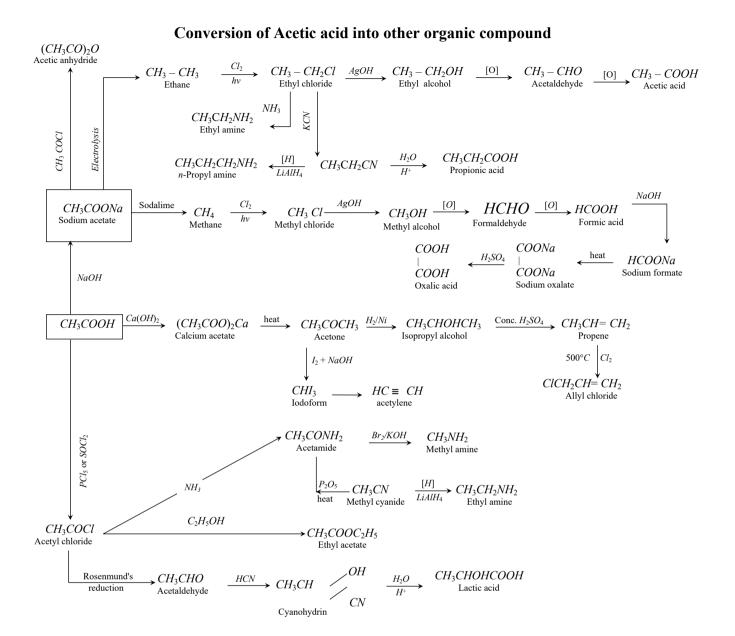
(i) $HCOOH \xrightarrow{Ca(OH)_2} (HCOO)_2 Ca \xrightarrow{heat} HCHO$ Formic acid Calcium formate Formaldehyde $CH_3 CHO \leftarrow [0] CH_3 CH_2 OH \leftarrow \frac{H_2 O}{H^+} CH_3 CH_2 OMgBr$ Acetaldehyde Ethyl alcohol H^+ Addition product $\downarrow [0] CH_3 COOH$ Aceticacid (ii) $HCHO \xrightarrow{H_2/Ni} CH_3 OH \xrightarrow{HI} CH_3 I$ Formaldehyde $CH_3 COOH \leftarrow H_2 O$ $Aceticacid HI CH_3 OH \xrightarrow{HI} CH_3 I$ Methyl alcohol Methyl iodide $CH_3 COOH \leftarrow \frac{H_2 O}{H^+} CH_3 CN$ $Aceticacid HI CH_3 CN$

Arndt-Eistert homologation : This is a convenient method of converting an acid, *RCOOH* to *RCH*₂*COOH*.

 $\begin{array}{c} RCOOH \xrightarrow{SOC_{2}} RCOCI \xrightarrow{CH_{2}N_{2}} RCOCHN_{2} \\ & EiOH \mid Ag_{2}O \\ RCH_{2}COOH \xleftarrow{Hydrolysis} RCH_{2}COOC_{2}H_{5} \end{array}$

(2) Descent of series : Conversion of acetic acid into formic





Dicarboxylic acids

F

The acids containing two carboxylic groups are called dicarboxylic acids.

The saturated dicarboxylic acid are represented by the general formula $C_n H_{2n}(COOH)_2$ where n = 0, 1, 2, 3 etc.

$$HO - C - (CH_2)_n - C - OH \text{ or } HOOC(CH_2)_n COOH$$

According to IUPAC system, the suffix-dioic acid is added to the name of parent alkane, *i.e.* Alkane dioxic acid.

| Table : 2 | 8.2 |
|-----------|-----|
|-----------|-----|

| Formula | Common name | IUPAC name |
|--|---------------|-----------------------|
| НООССООН | Oxalic acid | Ethanedioic acid |
| HOOCCH ₂ COOH | Malonic acid | 1-3 Propanedioic acid |
| HOOCCH ₂ CH ₂ COOH | Succinic acid | 1,4-Butanedioic acid |
| HOOC(CH ₂) ₃ COOH | Glutaric acid | 1,5-Pentanedioic acid |
| HOOC(CH ₂) ₄ COOH | Adipic acid | 1,6-Hexanedioic acid |

Oxalic Acid or Ethanedioic Acid COOH / or (COOH)₂ or (C₂H₂O₄) COOH

Oxalic acid is first member of dicarboxylic series.

It occurs as potassium hydrogen oxalate in the wood sorel, rhubarb and other plants of oxalis group and as calcium oxalate in plants of rumex family.

It is found in the form of calcium oxalate in stony deposits in kidneys and bladdar in human body.

Oxalic acid present in tomatoes.

(1) Methods of Preparation

(i) By oxidation of ethylene glycol with acidified potassium dichromate

$$\begin{array}{c} CH_2OH \\ | & +4[O] \xrightarrow{K_2Cr_2O_7} & COOH \\ CH_2OH & H_2SO_4 & COOH \\ Glycol & & COOH \end{array}$$

(ii) By hydrolysis of cyanogen with conc. hydrochloric acid :

$$\begin{array}{c} CN \\ | +4H_2O \xrightarrow{2(HCI)} | +2NH_4CI \\ CN \end{array}$$

(iii) By heating sodium or potassium in a current of carbon dioxide at $360^{\circ}C$

$$2Na+2CO_2 \xrightarrow{heat} COONa$$

 $COONa$
Sodium oxalate

(iv) Laboratory preparation

$$C_{12}H_{22}O_{11} + 18[O] \xrightarrow{HNO_3} 6 | + 5H_2O$$
Sucrose $COOH$
Ovalicacid

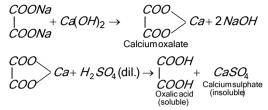
(v) Industrial method

$$\begin{array}{c} 2 \text{HCOONa} \xrightarrow{360^{\circ} \text{C}} & \begin{array}{c} \text{COONa} \\ & & \\ \text{Sod.formate} & & \\ & \\ & \\ & & \\ & & \\ & & \\ & & \\ & & \\ & & \\ & & \\ & & \\ & & \\ & \\ & &$$

Sodium formate is obtained by passing carbon monoxide over fine powdered of sodium hydroxide.

$$CO+ NaOH \xrightarrow{200^{\circ}C} HCOONa$$

The sodium oxalate thus formed is dissolved in water and calcium hydroxide is added. The precipitate of calcium oxalate is formed which is separated by filtration. It is decomposed with calculated quantity of dilute sulphuric acid.



(2) Physical Properties

(i) It is a colourless crystalline solid. It consists of two molecules of water as water of crystallisation.

(ii) The hydrated form has the melting point $101.5^{\circ}C$ while the anhydrous form melts at $190^{\circ}C$.

(iii) It is soluble in water and alcohol but insoluble in ether.

(iv) It is poisonous in nature. It affects the central nervous system.

(3) Chemical Properties

(i) Action of heat : It becomes anhydrous.

$$(COOH)_2 2H_2 O \xrightarrow{100-105^{\circ}C} (COOH)_2 + 2H_2 O$$
Hydrated oxalic
acid oxalic acid

(a) At 200°C, (COOH)₂ \longrightarrow HCOOH+ CO₂

On further heating, formic acid also decomposes. $HCOOH \rightarrow CO_2 + H_2$

(b) Heating with conc. H_2SO_4

$$\begin{array}{c} COOH \\ | & \xrightarrow{H_2SO_4} CO+CO_2+H_2O \\ COOH & \xrightarrow{(conc)} \end{array}$$

(ii) Acidic nature

Salt formation

$$\begin{array}{ccc} COOH \\ | & + KOH \rightarrow & | & KOH \\ COOH \\ Oxalic acid & Acid pot. oxalate \\ \hline COOH \\ \hline COONa \\ \hline \end{array} \xrightarrow{KOH} & COOK \\ Pot. oxalate \\ \hline COONa \\ \hline \end{array}$$

$$| + 2NaHCO_3 \rightarrow | + 2CO_2 + 2H_2O$$

$$COOH$$

$$COONa$$
Sod. oxalate

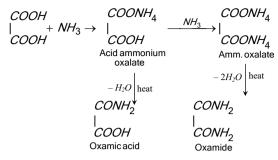
$$\begin{array}{c} COOH & COONa \\ | & + Na_2CO_3 \rightarrow | & + H_2O + CO_2 \\ COOH & COONa \end{array}$$

(iii) Esterification

(iv) Reaction with PCl₅:

$$\begin{array}{c} COOH & COCI \\ | + 2PCI_5 \rightarrow | + 2POCI_3 + 2HCI_5 \\ COOH & COCI \\ Oxalyl \\ chloride \end{array}$$

(v) Reaction with ammonia



(vi) **Oxidation** : When oxalic acid is warmed with acidified $KMnO_{4}$.

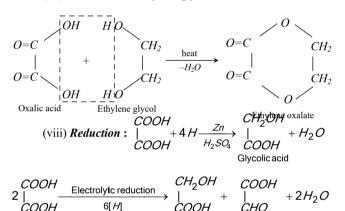
$$2KMnO_{4} + 3H_{2}SO_{4} \rightarrow K_{2}SO_{4} + 2MnSO_{4} + 3H_{2}O + 5[O]$$

$$\begin{bmatrix} COOH \\ I \\ COOH \\ +[O] \rightarrow 2CO_{2} + H_{2}O \end{bmatrix} \times 5$$

$$2KMnO_{4} + 3H_{2}SO_{4} + 5 \xrightarrow[OOH]{COOH}{Oxalic acid}} \rightarrow \underbrace{K_{2}SO_{4} + 2MnSO_{4}}_{Oxalic acid} + 10CO_{2} + 8H_{2}O$$

 \Box Oxalic acid decolourises the acidic KMnO₄ solution.

(vii) Reaction with ethylene glycol



Glycolic acid Glyoxalic acid

(ix) **Reaction with Glycerol :** At $100^{\circ} - 110^{\circ}C$, formic acid is formed. At 260°, allyl alcohol is formed.

(4) Uses : Oxalic acid (Polyprotic acid) is used,

(i) In the manufacture of carbon monoxide, formic acid and allyl alcohol.

(ii) As a laboratory reagent and as a standard substance in volumetric analysis.

(iii) In the form of antimony salt as a mordant in dyeing and calico printing.

(iv) In the manufacture of inks.

 $\left(v\right)$ For removing ink stains and rust stains and for bleaching straw, wood and leather.

(vi) In the form of ferrous potassium oxalate as developer in photography.

(5) Analytical test

(i) The aqueous solution turns blue litmus red.

(ii) The aqueous solution evolves efferve scences with $NaHCO_3$. (iii) The neutral solution gives a white precipitate with calcium chloride solution. It is insoluble in acetic acid.

$$\begin{array}{c} \mathcal{H}_2 C_2 \mathcal{O}_4 \xrightarrow{\mathcal{NH}_4 \mathcal{OH}} (\mathcal{NH}_4)_2 C_2 \mathcal{O}_4 \xrightarrow{\mathcal{CaC}_2} \mathcal{CaC}_2 \mathcal{O}_4 \\ \text{Oxalicacid} \xrightarrow{\mathcal{M}_4 \mathcal{OH}} (\mathcal{NH}_4)_2 \mathcal{C}_2 \mathcal{O}_4 \xrightarrow{\mathcal{CaC}_2} \mathcal{O}_4 \end{array}$$

(iv) Oxalic acid decolourises hot potassium permanganate solution having dilute sulphuric acid.

(v) With hot conc. H_2SO_4 , it evolves carbon monoxide which burns with blue flame.

Malonic Acid or Propane-1,3-Dioic Acid

$$CH_2 \xrightarrow{COOH} Or CH_2(COOH)_2 \text{ or } (C_3H_4O_4)$$

The acid occurs as calcium salt in sugar beet. It was so named because it was first obtained from malic acid (hydroxy succinic acid) by oxidation.

(1) Methods of Preparation : From acetic acid

$$\begin{array}{c} CH_{3}COOH \xrightarrow{Cl_{2}} CH_{2}CICOOH \xrightarrow{KCN(Aq)} \\ Aceticacid & P & Chloroacetcacid \\ \hline \\ CH_{2}CNCOOH \xrightarrow{H_{2}O/H^{+}} CH_{2} \swarrow COOH \\ Cyano aceticacid & COOH \end{array}$$

Malonic acid

(2) Physical Properties

(i) It is a white crystalline solid.

(ii) It's melting point is $135^{\circ}C$.

(iii) It is soluble in water and alcohol but sparingly soluble in

(3) Chemical Properties

(i) Action of heat

ether.

(a) Heating at $150^{\circ}C$:

 $CH_2(COOH)_2 \rightarrow CH_3COOH + CO_2$

(b) Heating with
$$P_2O_5$$
:
 $\begin{bmatrix} \overline{H} & \overline{OH} \\ \uparrow & \uparrow & -\uparrow \end{bmatrix}$
 $O = \underbrace{C - C - C}_{-L} = O \xrightarrow{P_2O_5}_{-L} O = \underbrace{C = C = C}_{+2H_2O} O = C = C = C = O + 2H_2O$
heat Carbon suboxide

[OH H](ii) **Reaction with aldehyde :** With aldehydes, α - β unsaturated acids are formed.

$$RCH = O + H_2 C \xrightarrow{COOH} \xrightarrow{Pyridine}_{heat}$$

$$RCH = CHCOOH + H_2 O + CO_2$$

$$a - \beta \text{ unsaturated acid}$$

(4) Uses : Its diethyl ester (malonic ester) is a valuable synthetic reagent for preparation of a variety of carboxylic acids.

$$| \qquad \text{or } (CH_2)_2(COOH)_2 \text{ or } (C_4H_6O_4)$$

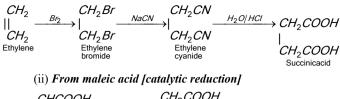
$$CH_2-COOH$$

It was first obtained by the distillation of yellow fossil, resin, amber and hence its name (Latin, *Succinum* = amber).

It is also formed in small amount during the fermentation of sugar.

(1) Methods of Preparation

(i) From ethylene



$$|| + H_2 \xrightarrow{Ni} |$$

$$CHCOOH \xrightarrow{H_2COOH} H_2 \xrightarrow{Ni} |$$

□ This is an industrial method.

(iii) Reduction of tartaric acid or malic acid

 $\begin{array}{c} CHOHCOOH \\ | \\ CHOHCOOH \\ Tartaric acid \end{array} \xrightarrow{P} \begin{array}{c} CH_2COOH \\ | \\ CH_2COOH \\ Succinicacid \end{array} \xrightarrow{HI} \begin{array}{c} CHOHCOOH \\ | \\ CH_2COOH \\ Malic acid \end{array}$

(2) Physical properties

(i) It is a white crystalline solid. It melts at $188^{\circ}C$

(ii) It is less soluble in water. It is comparatively more soluble in alcohol.

(3) **Chemical Properties :** Succinic acid gives the usual reactions of dicarboxylic acid, some important reactions are :

(i) Action of heat : At 300°C CH_2COOH 300°*C* СН₂СООН CH₂CO Succinicacid Succinicanhvdride (ii) With ammonia CH₂COOH CH3COONH NH_3 heat Ċн₂соон ĊH₂COONH₄ H_2O Ammonium succinate CH₂CC heat $-NH_3$ CH,CONH, CH2CO Succinamide Succinimide

(iii) Reaction with Br₂

ł

 $\begin{array}{c} CH_2 - CO \\ | \\ CH_2 - CO \\ \text{Succinimide} \end{array} \qquad NH + Br_2 \xrightarrow{NaOH} O^{\circ}C \\ | \\ O^{\circ}C \\ N^{\circ}C \\$

(iv) Reaction with ethylene glycol

$$HOOC - (CH_2)_2 - CO[\overline{PH} + \overline{H}]OCH_2 - CH_2O[\overline{H} + \overline{H}]OCH_2 - CH_2O[\overline{H} + \overline{H}]OCH_2)_2 - COOH + \dots + CH_2O$$

$$HOOC - (CH_2)_2 - CO - [-OCH_2 - CH_2O - OC] - CH_2O - OC$$

$$= (CH_2)_2 - CO[\overline{PH} + H_2O] - OC$$

 $-(CH_2)_2 - CO_{-}]_n - OH + H_2O_{-}$ Polyester

When sodium or potassium salt in aqueous solution is electrolysed, ethylene is obtained at anode.

(4) Uses : It finds use in volumetric analysis, medicine and in the manufacture of dyes, perfumes and polyester resins.

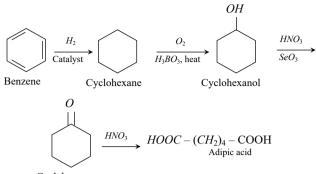
Adipic Acid or Hexane-1,6 –Dioic Acid

$$CH_{2}-CH_{2}-COOH
| or (CH_{2})_{4}(COOH)_{2} or (C_{6}H_{10}O_{4})
CH_{2}-CH_{2}-COOH$$

It was first obtained by the oxidation of fats (Latin, adeps = fat.)

(1) Methods of Preparation

(i) From benzene (In industries)



IFIShanan Russtrial method.

(ii) From tetrahydrofuran (THF)

$$\begin{array}{c} CH_2 - CH_2 \\ | & | \\ CH_2 & CH_2 \\ \hline O \\ THF \end{array} + 2CO + HOH \rightarrow HOOC - (CH_2)_4 - COOH \\ Adipic acid \\ Adipic acid \\ \end{array}$$

(2) Physical Properties

(i) It is a white crystalline solid. Its melting point is $150^{\circ}C$.

(ii) It is fairly soluble in alcohol and ether but less soluble in

 H_2

water.

(3) Chemical Properties

It shows all the general reaction of dicarboxylic acids.

(i) Action of heat

HOOC(CH₂)₄ COOH
$$\stackrel{\text{heat}}{300^{\circ}C}$$
 $\stackrel{H_2C}{\underset{H_2C}{\overset{|}}}$ $C = O + CO_2 + H_2O$

(ii) Formation of Nylon-66 [Reaction with hexa methylene diamine]

(4) Uses : It is used in the manufacture of several polymers. Unsaturated Acids : When the double bond presents in the carbon chain of an acid is called unsaturated acid.

Example:
$$CH_2 = CH - COOH + H - C - COOH$$

Acrylicacid

Acrylic Acid or Prop-2-Enoic Acid $CH_2 = CH - COOH$ or $(C_3H_4O_2)$ (1) Methods of Preparation (i) From allyl alcohol

$$\begin{array}{cccc} CH_2 & CH_2Br & CH_2Br & CH_2\\ \parallel & & & \\ CH & & & \\ -CH & & & \\ -CH_2OH & & \\ CH_2OH & & \\ \end{array} \xrightarrow[O]{} HBr & & \\ -HNO_3 \\ -HNO_3 \\ -HNO_3 \\ -HNO_3 \\ -HBr & \\ -CHBr & \\ -CHBr \\ -HBr & \\ -HBr \\$$

(ii) By oxidation of acrolein

$$CH_2 = CHCHO + [O] \xrightarrow{AgNO_3} CH_2 = CHCOOH$$

(iii) *From propionic acid* : $CH_3CH_2COOH \xrightarrow{Br_2/P} HVZ$ reaction

 $\begin{array}{c} CH_3CHBrCOOH \xrightarrow{Alc.KOH} CH_2 = CHCOOH \\ \alpha\text{-Bromopropionic acid} \end{array}$

(iv) By heating β-hydroxy propionic acid

 $\begin{array}{c} CH_2 - CH_2 - COOH \xrightarrow{ZnCl_2} CH_2 = CH - COOH \\ | \\ OH \\ \beta \text{-hydroxy propionic acid} \end{array}$

- (v) From vinyl cyanide
- $\begin{array}{c} HC \equiv CH + HCN \xrightarrow{Cu_2Cl_2/HCl} CH_2 = CH CN \\ \text{Acetylene} & 90^{\circ}C & \text{Vinyl cyanide} \end{array}$

$$\xrightarrow{H^{+}/H_{2}O} CH_{2} = CH - COOH$$

(vi) From ethylene cyanohydrin

$$\begin{array}{c} CH_2 - CH_2 \xrightarrow{+HCN} CH_2 - CH_2 - CN \xrightarrow{\text{Conc.}H_2SQ_4} \\ & & \\ & & \\ OH \\ CH \\ Ethylenecyanohydrin \\ Ethyleneoxide \end{array}$$

$$CH_2 = CH - CN \xrightarrow{H^+/H_2O} CH_2 = CHCOOH$$

Vinyl cyanide
(acrylonitrile)

Industrial method : This is a new method of its manufacture.

$$CH \equiv CH + CO + H_2O \xrightarrow{N(CO)_4} CH_2 = CHCOOH$$

(2) Physical Properties

 \Box It is colourless pungent smelling liquid. Its boiling point is 141°*C*.

 \Box It is miscible with water, alcohol and ether.

□ It shows properties of an alkene as well as of an acid.

(3) Chemical Properties

(i) With nascent hydrogen (Na and C_2H_5OH)

$$CH_2 = CHCOOH + 2[H] \longrightarrow CH_3CH_2COOH$$

(ii) *With halogens and halogen acids* : Markownikoff's rule is not followed.

$$CH_{2} = CHCOOH + Br_{2} \xrightarrow{CCl_{4}} CH_{2}Br - CHBrCOOH$$

$$\alpha,\beta\text{-Dibromopropionic acid}$$

$$CH_{2} = CHCOOH + HBr \rightarrow BrCH_{2} - CH_{2}COOH$$

$$\beta\text{-Bromopropionic acid}$$

(iii) **Oxidation** : In presence of dilute alkaline
$$KMnO_4$$
.
 $CH_2 = CHCOOH + [O] + H_2O \rightarrow CH_2OHCHOHCOOH$
Glycericacid

□ On vigorous oxidation, oxalic acid is formed. (iv) *Salt formation*

$$CH_2 = CHCOOH + KOH \rightarrow CH_2 = CHCOOK^+ + H_2O$$
$$2CH_2 = CHCOOH + Na_2CO_3 \rightarrow$$

 $2CH_2 = CHCOONa^+ + H_2O + CO_2$ Sodium acrylate

(v) *Ester formation*

$$CH_{2} = CHCOOH + HOC_{2}H_{5} \xrightarrow{Conc.H_{2}SO_{4}} \rightarrow \\ -H_{2}O \\ CH_{2} = CH - COOC_{2}H_{5} \\ Ethyl acrylate$$

(vi) With PCl₅

$$CH_2 = CHCOOH + PCl_5 \rightarrow CH_2 = CH - COCI$$

Acryl chloride

(4) Uses : Its ester are used for making plastics such as Lucite and plexiglass.

Unsaturated dicarboxylic acids

The molecular formula of the simplest unsaturated dicarboxylic acid is HOOC.CH = CH.COOH This formula, however represents two chemical compounds, maleic acid and fumaric acid, which are geometrical isomers.

$$\begin{array}{ccc} H-C-COOH & HOOC-C-H \\ \parallel \\ H-C-COOH & H-C-COOH \\ Cis form (Maleicacid) & Transform (Fumaricacid) \end{array}$$

(1) Methods of Preparation of Maleic Acid

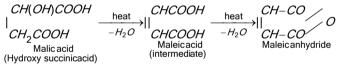
(i) By catalytic oxidation of 2-butene or benzene

$$CH-CH_{3} + 30_{2} \xrightarrow{V_{2}O_{5}} CHCOOH + 2H_{2}O$$

$$CH-CH_{3} + 30_{2} \xrightarrow{V_{2}O_{5}} || + 2H_{2}O$$

$$CHCOOH$$
Maleicacid
$$C_{6}H_{6} + \frac{9}{2}O_{2} \xrightarrow{V_{2}O_{5}} || CH-CO \longrightarrow O \xrightarrow{H_{2}O/H^{+}} || CHCOOH$$
Maleicanhydride
(ii) From malic acid :

(ii) From malic acid :



$$\begin{array}{c} \xrightarrow{NaOH} & CH-COONa \\ \xrightarrow{boil} & H^+/H_2O \\ CH-COONa \\ Sodium salt \\ \end{array} \begin{array}{c} CH-COOH \\ CH-COOH \\ Maleicacid \end{array}$$

(2) Methods of Preparation of Fumaric Acid

(i) From maleic acid :

_

$$\begin{array}{ccc} H-C-COOH & HOOC-C-H \\ \parallel & & \parallel \\ H-C-COOH & \text{boil} & H-C-COOH \\ Maleicacid & & H-C-COOH \end{array}$$

(ii) By oxidation of furfural with sodium chlorate

$$\begin{array}{c} HC \longrightarrow CH \\ \parallel & \parallel & +4[O] \longrightarrow HOOC-C-H \\ HC & C-CHO \end{array} \rightarrow \begin{array}{c} HOOC-C-H \\ \parallel & \parallel & +CO_2 \\ H-C-COOH \end{array}$$

iii) By heating malic acid at about 150°C for long time

$$CH(OH)COOH \xrightarrow{\text{heat}} HOOC-C-H$$

 $HOOC-C-H$
 $HOOC-C-H$
 $H-C-COOH$
Malicacid

(iv) *By heating bromosuccinic acid with alcoholic potash* : By heating bromosuccinic acid with alcoholic potash.

$$\begin{array}{c} CH_2COOH \\ | \\ CH.(Br)COOH \end{array} \xrightarrow{Alc. KOH} HOOC-C-H \\ || \\ H-C-COOH \end{array} + KBr + H_2O$$

(3) Physical Properties

(i) Both are colourless crystalline solids. Both are soluble in water.

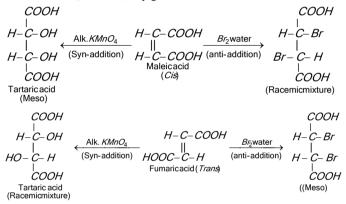
(ii) The melting point of maleic acid $(130.5^{\circ}C)$ is lower than the melting point of fumaric acid (287°*C*).

(4) Chemical Properties

Chemically, both the acids give the reactions of alkenes and dibasic acids except that the maleic acid on heating forms an anhydride while fumaric acid does not give anhydride.

$$\begin{array}{c} CHCOOH \\ || \\ CHCOOH \\ Maleicacid \end{array} \xrightarrow{heat} \begin{array}{c} CHCO \\ || \\ CHCO \\ Maleicanhydride \end{array} O+ H_2O$$

Both form succinic acid on reduction with sodium amalgam. They undergo addition reactions with bromine, hydrobromic acid, water, etc. and form salts, esters and acid chlorides as usual. With alkaline $KMnO_4$ solution, they get oxidised to tartaric acid.



Higher fatty acids

Palmitic, stearic and **oleic acids** are found in natural fats and oils as glyceryl esters.

They have derived their names from the natural source from which they are prepared by hydrolysis with alkali.

| Name of acids | Source | Molecular formula |
|---------------|------------------------------|---|
| Palmitic acid | Palm oil | СН ₃ (СН ₂) ₁₄ СООН |
| Stearic acid | Stear (meaning tallow) | СН ₃ (СН ₂) ₁₆ СООН |
| Oleic acid | Olive oil. | $CH_3(CH_2)_7 CH = CH(CH_2)_7 COOH$ |

Palmitic and stearic acids are waxy colourless solids with

melting points $64^{\circ}C$ and $72^{\circ}C$, respectively. They are insoluble in water but soluble in ethanol and ether. They find use in the manufacture of soaps and candles. Soaps contain sodium or potassium salts of these higher fatty acids.

Oleic acid has low melting point, *i.e.*, 16°C. It is insoluble in water but soluble in alcohol and ether. Besides the reactions of acids, it also gives reactions of alkenes. Two aldehydes are formed on ozonolysis.

$$CH_{3}(CH_{2})_{7} CH = CH(CH_{2})_{7} COOH \xrightarrow{(i)O_{3}}_{(i)Zn+H_{2}O}$$
$$CH_{3}(CH_{2})_{7} CHO + HOOO(CH_{2})_{7} CHO$$

It is used for making soaps, lubricants and detergents.

(1) **Difference between oils and fats :** Oils and fats belong to the same chemical group, yet they are different in their physical state.

(i) Oils are liquids at ordinary temperature (below $20^{\circ}C$) while fats are semi solids or solids (their melting points are more than $20^{\circ}C$). A substance may be classed as fat in one season and oil in another season or the same glyceride may be solid at a hill station and liquid in plains. Thus, this distinction is not well founded as the physical state depends on climate and weather.

(ii) The difference in oils and fats is actually dependent on the nature of monocarboxylic acid present in the glyceride. Oils contain large proportion of the glycerides of lower carboxylic acids, (*e.g.*, butyric acid, caprylic acid and caproic acid) and unsaturated fatty acids, (*e.g.*, oleic, linoleic and linolenic acids) while fats contain a large proportion of glycerides of higher saturated carboxylic acids, (*e.g.*, palmitic, stearic acids).

Lard (fat of hogs) is a solid fat and its composition in terms of fatty acids produced on hydrolysis is approximately 32% palmitic acid, 18% stearic acid, 45% oleic acid and 5% linolenic acid. Olive oil on the other hand, contains 84% oleic acid, 4% linoleic acid, 9% palmitic acid and 3% stearic acid.

(2) Physical Properties of oils and Fats

(i) Fats are solids, whereas oils are liquids.

(ii) They are insoluble in water but soluble in ether, chloroform and benzene.

(iii) They have less specific gravity than water and consequently float on the surface when mixed with it.

(iv) Pure fats and oils are colourless, odourless and tasteless but natural fats and oils possess a characteristic odour due to presence of other substances.

(v) They have specific melting points, specific gravity and refractive index hence they can be identified by these oil constants.

(vi) Animal fats contain cholesterol, an unsaturated alcohol, whereas vegetable fats contains phytosterol.

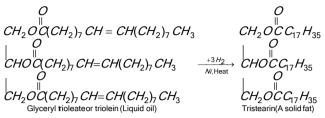
(3) **Chemical Properties :** They give reactions of carbon-carbon double bonds and ester groups.

(i) *Hydrolysis*

(a) By superheated steam

(c) *Enzyme hydrolysis* : Enzyme like lipase, when added to an emulsion of fat in water, hydrolyses it into acid and glycerol in about two or three days.

(ii) *Hydrogenation* : In the presence of finally divided nickel, at low pressure the hydrogenation process is called hardening of oils.



(iii) Hydrogenolysis [Reduction to alcohol]

$$\begin{array}{c} O \\ CH_2 - O - \overset{||}{C} - C_1 C_{17} H_{35} \\ | O \\ CH - O - \overset{||}{C} - C_{17} H_{35} \xrightarrow{6H_2} CH_2 OH \\ | O \\ | O \\ CH_2 - O - \overset{||}{C} - C_{17} H_{35} \\ Tristearin \end{array} \xrightarrow{CH_2 OH} \begin{array}{c} CH_2 OH \\ CH_2 OH \\ CH_2 OH \end{array} \xrightarrow{Octadecylalcohol} CH_2 OH \\ CH_2 OH \end{array}$$

(iv) **Drying**: Certain oils, containing glycerides of unsaturated fatty acids having two or three double bonds have the tendency of slowly absorbing oxygen from atmosphere and undergoing polymerisation to form hard transparent coating. This process is known as drying and such oils are called drying oils. Unsaturated oils such as linseed oil are, therefore, used as medium of paints and varnishes.

(v) **Rancidification :** On long storage in contact with air and moisture, oils and fats develop unpleasant smell. The process is known as rancidification. It is believed that rancidification occurs due to hydrolysis-oxidation.

(4) Analysis of oils and fats

(i) Acid value : It indicates the amount of free acid present in the oil or fat. It is defined as the number of milligrams of KOH required to neutralize the free acid present in one gram of the oil or fat. It is determined by dissolving a weighed amount of oil or fat in alcohol and titrating it against a standard solution of KOH using phenolphthalein as an indicator.

(ii) *Saponification value*: It is a measure of fatty acids present as esters in oils and fats. It is defined as the number of milligrams of *KOH* required to saponify one gram of the oil or fat or number of milligrams of *KOH* required to neutralize the free acids resulting from the hydrolysis of one gram of an oil or fat. It is determined by refluxing a Saponification number of fat or oil

 $=\frac{168,000}{M}$, Where M= molecular mass

(iii) *Iodine value*: Iodine value of a fat or oil is a measure of its degree of unsaturation. *It is defined as the number of grams of iodine taken up by 100 grams of fat or oil for saturation.* For a saturated acid glyceride, the iodine value is zero. Thus, the iodine value for a fat is low whereas for oil, it is high. As iodine does not react readily, in actual practice, iodine monochloride is used. Iodine monochloride is known as Wij's reagent.

(iv) **Reichert-Meissl value, (R/M value) :** It indicates the amount of steam volatile fatty acids present in the oil or fat. It is defined as the number of millilitres of 0.1 N KOH solution required to neutralize the distillate of 5 grams of hydrolysed fat. It is determined by hydrolysing a known weighed amount (5 grams) of the fat with alkali solution and the mixture is acidified with dilute sulphuric acid and steam distilled. The distillate is cooled, filtered and titrated against 0.1 N KOH.

(5) Uses

(i) Many oils and fats are used as food material.

(ii) Oils and fats are used for the manufacture of glycerol, fatty acids, soaps, candles, vegetable ghee, margarine, hair oils, etc.

(iii) Oils like linseed oil, tung oil, etc., are used for the manufacture of paints, varnish, etc.

(iv) Castor oil is used as purgative and codliver oil as a source of vitamins *A* and *D*. Almond oil is used in pharmacy. Olive oil is also used as medicine.

(v) Oils are also used as lubricants and illuminants.

| Property | Vegetable oils | Minerals oils |
|--|--|--|
| 1. Composition | These are triesters of glycerol with higher fatty acids. | These are hydrocarbons (saturated). Kerosene oil-Alkanes from C_{12} to C_{16} . |
| 2. Source | Seeds root and fruits of plants. | These occur inside earth in the form of petroleum. |
| 3. Hydrolysis | Undergo hydrolysis with alkali. Form soap and glycerol. | No hydrolysis occurs. |
| 4. On adding <i>NaOH</i> and phenolphthalein | Decolourisation of pink colour occurs. | No effect. |
| 5. Burning | Burns slowly | Burn very readily. |
| 6. Hydrogenation | Hydrogenation occurs in presence of nickel catalyst. Solid glycerides (fats) are formed. | No hydrogenation occurs. |

Table : 28.4 Difference between vegetable oils and Mineral oils

(6) **Soaps :** Soaps are the metallic salts of higher fatty acids such as palmitic, stearic, oleic, etc. The sodium and potassium salts are the common soaps which are soluble in water and used for cleansing purposes. Soaps of other metals such as calcium, magnesium, zinc, chromium, lead, etc., are insoluble in water. These are not used for cleansing purposes but for other purposes (lubricants, driers, adhesives, etc.)

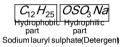
Ordinary soaps (sodium and potassium) are the products of hydrolysis of oils and fats with sodium hydroxide or potassium hydroxide. The oils and fats are mixed glycerides and thus *soaps are mixtures of salts of saturated and unsaturated long chain carboxylic acids containing 12 to 18 carbon atoms.* This process always yields glycerol as a byproduct.

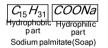
$$\begin{array}{ccc} CH_2OCOR_1 & CH_2OH & R_1COONa \\ | \\ CHOCOR_2 &+ 3NaOH \rightarrow & CHOH &+ \\ | \\ CH_2OCOR_3 & CH_2OH & R_3COONa \\ | \\ Triglyceride & Glycerol & Soap \end{array}$$

There are three methods for manufacture of soaps :

- (i) The cold process
- (ii) The hot process
- (iii) Modern process

(7) **Synthetic Detergents :** The synthetic detergents or **Syndets** are substitutes of soaps. They have cleansing power as good or better than ordinary soaps. Like soap, they contain both hydrophilic (water soluble) and hydrophobic (oil-soluble) parts in the molecule.





Some of the detergents used these days are given below:

(i) *Sodium alkyl sulphates* : These are sodium salts of sulphuric acid esters of long chain aliphatic alcohols containing usually 10 to 15 carbon atoms. The alcohols are obtained from oils or fats by hydrogenolysis.

$$CH_3(CH_2)_{10} CH_2 OH + HO SO_3 H \rightarrow$$
Lauryl alcohol Sūlphūričacid

$$CH_3(CH_2)_{10} CH_2 OSO_2 OH \xrightarrow{NaOH}$$

Lauryl hydrogen sulphate

 $CH_3(CH_2)_{10} CH_2 OSO_2 ONa$ Sodium lauryl sulphate(Detergen)

 ~ 1.1

The other examples are sodium cetyl sulphate, $C_{16}H_{33}OSO_2ONa$ and sodium stearyl sulphate, $CH_3(CH_2)_{16}CH_2OSO_3Na$. Unlike ordinary soaps, they do not produce OH^- ions on hydrolysis and thus can be safely used for woollen garments.

(ii) *Sodium alkyl benzene sulphonates* : Sodium *p*-dodecyl benzene sulphonate (S.D.S.) acts as a good detergent. It is most widely used since 1975.

$$CH_{3}(CH_{2})_{9}CH = CH_{2} + C_{6}H_{6} \xrightarrow{AICI_{3}} CH_{3}(CH_{2})_{9}CH - C_{6}H_{5}$$
1-Dodecene 2-Dodecyl benzene

$$\xrightarrow{(1)H_2SO_4} CH_3 - (CH_2)_9 - \overset{CH_3}{\overset{(i)}{C}H_4 - SO_3Na}$$

These long chain alkyl benzene sulphonate (L.A.S.) are most widely used syndets.

(iii) **Quaternary ammonium salts :** Quaternary ammonium salts with long chain alkyl group have been used as detergents, *e.g.*, trimethyl stearyl ammonium bromide.

$$(CH_3)_3 N \swarrow Br \\ C_{18}H_{37}$$

(iv) Sulphonates with triethanol ammonium ion in place of sodium serve as highly soluble materials for liquid detergents.

$$R - \left\langle \bigcirc \right\rangle - O - \bar{SO}_{2} \left[\overset{\oplus}{N} H (-CH_{2} - CH_{2}OH)_{3} \right]$$

(v) Partially esterified polyhydroxy compounds also acts as detergents.

$$CH_{2}OH$$

$$C_{17}H_{35}COOCH_{2} - C_{-}CH_{2}OH$$

$$CH_{2}OH$$
Pentaerythritol monostearate

Detergents are superior cleansing agents due to following properties.

(i) These can be used both in soft and hard waters as the calcium and magnesium ions present in hard water form soluble salts with detergents. Ordinary soap cannot be used in hard water.

(ii) The aqueous solution of detergents are neutral. Hence these can be used for washing all types of fabrics without any damage. The solution or ordinary soap is alkaline and thus cannot be used to wash delicate fabrics.

(8) **Waxes :** Waxes are the esters of higher fatty acids with higher monohydric alcohols. The acids and alcohols commonly found in waxes are palmitic, cerotic acid ($C_{25}H_{51}COOH$), melissic acid ($C_{30}H_{61}COOH$) and cetyl alcohol ($C_{16}H_{33}OH$), ceryl alcohol ($C_{26}H_{53}OH$), myricyl alcohol ($C_{30}H_{61}OH$), etc.

Waxes are insoluble in water but are readily soluble in benzene, petroleum, carbon disulphide etc. Waxes on hydrolysis with water yields higher fatty acids and higher monohydric alcohols.

$$\begin{array}{c} C_{15}H_{31}COOC_{16}H_{33} + H_2O \rightarrow C_{15}H_{31}COOH + \begin{array}{c} C_{16}H_{33}OH \\ \text{Cetyl palmitate} \end{array} \begin{array}{c} C_{15}H_{31}COOH + \begin{array}{c} C_{16}H_{33}OH \\ \text{Palmiticacid} \end{array} \begin{array}{c} C_{15}H_{31}COOH + \begin{array}{c} C_{16}H_{33}OH \\ \text{Cetyl alcohol} \end{array} \end{array}$$

When hydrolysis is carried with caustic alkalies, soap and higher monohydric alcohols are formed.

$$C_{15}H_{31}COOC_{16}H_{33} + NaOH \rightarrow C_{16}H_{33}OH + C_{15}H_{31}COONa$$

Sodium palmitate(Soap)

The common waxes are:

(i) **Bees wax**, Myricyl palmitate,
$$C_{15}H_{31}COOC_{30}H_{61}$$

(ii) Spermaceti wax, Cetyl palmitate, $C_{15}H_{31}COOC_{16}H_{33}$

(iii) Carnauba wax, Myricyl cerotate, $C_{25}H_{51}COOC_{30}H_{61}$

Waxes are used in the manufacture of candles, polishes, inks, water proof coating and cosmetic preparations.

Waxes obtained from plants and animals are different than

paraffin wax which is a petroleum product and a mixture of higher hydrocarbons (20 to 30 carbon atoms). So paraffin wax is not an ester.

Candles are prepared by mixing paraffin wax (90%) with higher fatty acids like stearic and palmitic. The fatty acids are added to paraffin wax as to give strength to candles. The mixture is melted and poured into metal tubes containing streched threads. On cooling candles are obtained.

Substituted carboxylic acids

The compounds formed by the replacement of one or more hydrogen atoms of the hydrocarbon chain part of the carboxylic acids by atoms or groups such as X (halogen), OH or NH_2 , are referred to as substituted acids. For example,

> $CH_2OHCOOH$: $CH_2CICOOH$; CH₂NH₂COOH Chloroacetc acid Hydroxyacetic acid Aminoacetic acid

The position of the substituents on the carbon chain are indicated by Greek letters or numbers.

$$\begin{array}{c} 6 \\ \mathcal{C}-\\ \mathcal{C}-\\ \mathcal{S} \\ \mathcal{S} \end{array} \begin{array}{c} 4 \\ \mathcal{C}-\\ \mathcal{C}-\\ \mathcal{S} \end{array} \begin{array}{c} 4 \\ \mathcal{C}-\\ \mathcal{C}-\\$$

For example,

CH3CHOHCOOH ; CH3CHOHCH2COOH α -Hydroxypropionic acid 2-Hydroxypropanoic acid

Lactic Acid or α -hydroxy propionic acid or 2-hydroxy propanoic acid

 β -Hydroxybutyric acid 3-Hydroxybutanoic acid

It is the main constituent of sour milk. It is manufactured by fermentation of molasses by the micro-organism (Bacterium acidi lactici-sour milk) in presence of CaCO3.

(1) Method of Preparation

From acetaldehyde :

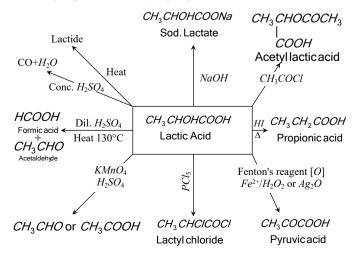
$$\begin{array}{c} CH_3 CHO + HCN \rightarrow CH_3 CH(OH)CN \xrightarrow{H_2O/H^+} \\ \text{Acetaldehyde} & \text{Cyanohydrin} \end{array}$$

(2) Physical Properties

It is a colourless syrupy liquid having a sour taste and smell.

It is hygroscopic and very soluble in water. It is optically active and exists in three distinct forms.

(3) Chemical Properties : It gives reactions of secondary alcoholic group and a carboxylic group.



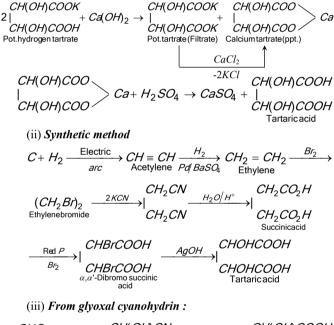
(4) Uses : It is used in medicine as calcium and iron lactates, as mordant in dyeing, as acidulant in beverages and candies, as a solvent (ethyl and butyl lactates) for cellulose nitrate.

Tartaric Acid. Or a.a'-Dihvdroxy succinic acid or 2.3-Dihydroxy-Butane-1,4-Dioic acid

It is found as free or potassium salt in grapes, tamarind, and berries.

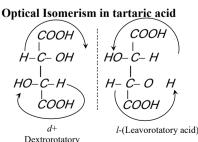
(1) Methods of Preparation

(i) Argol which separates as a crust during fermentation of grape juice is impure potassium hydrogen tartrate. Argol is boiled with limewater. Calcium tartrate is precipitated which is filtered. The solution contains potassium tartrate which is also precipitated by addition of CaCl₂. The calcium salt is then decomposed with calculated quantity of dilute H_2SO_4 . The precipitate (CaSO₄) is filtered and the filtrate on concentration gives the crystals of tartaric acid.



| CHO | HCN CH(OH)CN | H_2O/H^+ CH(OH)COOH |
|--------------------------------|--|--|
| I — — <i>CHO</i> Glyoxal | <i>CH(OH)CN</i> Glyoxal cyanohydrin | <i>CH</i> (<i>OH</i>) <i>COOH</i> Tartaric acid |

(2) Physical Properties : It is a colourless crystalline compound. It is soluble in water and alcohol but insoluble in ether. It contains two asymmetric carbon atoms and thus shows optical isomerism (four forms). Natural tartaric acid is the dextro variety. It contains two secondary alcoholic groups and two carboxylic groups.



Tartaric acid



Meso-Tartaric acid (Optical inactive)

°OOF

СООН

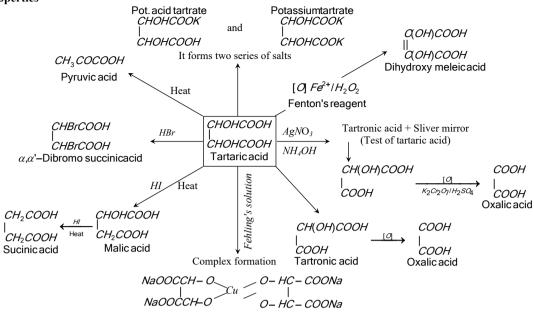
- OH

Optical active

(i) d + Tartaric acid-Dextro-rotatory _____ Optical active

(iii) Meso tartaric acid-optically inactive due to internal compensation.

(3) Chemical Properties



(4) **Uses :** It is used in carbonated beverages and effervescent tablets, in making baking powder (cream of tartar) and mordant in dyeing (potassium hydrogen tartrate), in preparing Fehling's solution (sodium potassium tartrate–Rochelle salt), in medicine as emetic, dyeing and calico-printing (tartar emetic-potassium antimonyl tartrate) and silver mirroring.

(5) Tests

(i) When heated strongly, tartaric acid chars readily giving a smell of burnt sugar to produce free carbon and pyruvic acid.

(ii) *With AgNO*₃ : A neutral solution of tartaric acid gives a white ppt. which is soluble in ammonia. A silver mirror is obtained on warming the ammonical silver nitrate solution (Tollen's reagent).

(iii) *With Fenton's reagent* : $(H_2O_2 \text{ containing a little of ferrous salt)} and caustic soda, It gives a violet colour.$

(iv) With Resorcinol and conc. H₂SO₄ : It gives blue colour.

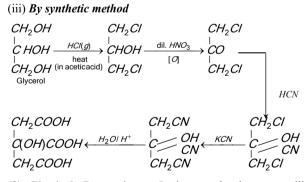
Citric Acid Or 2-Hydroxypropane Or 1,2,3-Tri Carboxylic Acid Or β-Hydroxy Tricarballylic Acid

It occurs in the juice of citrus fruits such as lemon, galgal, orange, lime, etc. Lemon juice contains 6-10% of citric acid.

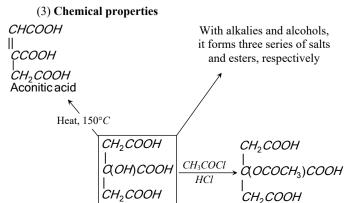
(1) Methods of Preparation

(i) **By Fermentation**: Citric acid is obtained by carrying fermentation of dilute solution of molasses with micro-organism, *Aspergillus nigar*, at 26-28°C for 7 to 10 days. The resulting solution is neutralised with $Ca(OH)_2$ to form insoluble precipitate, calcium citrate. It is decomposed by dilute H_2SO_4 . The $CaSO_4$ is filtered off and the solution is concentrated under vacuum to get crystals of citric acid.

(ii) By Lemon juice : It is also obtained from lemon juice. The juice is boiled to coagulate proteins. From clear solution, citric acid is obtained as calcium salt with $Ca(OH)_2$.



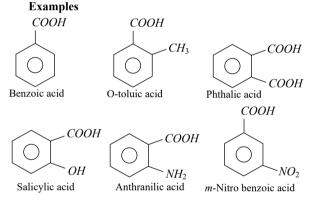
(2) **Physical Properties :** It is a colourless crystalline compound. It possesses one water molecule as water of crystallisation. It is soluble in water and alcohol but less soluble in ether. It is not optically active compound. It is nontoxic in nature. It behaves as an alcohol and tribasic acid.



(4) Uses : It finds use in making lemonades, as acidulant in food and soft drinks and makes the lemon sour, as mordant in dyeing and calico printing. Ferric ammonium citrate, magnesium citrate (as an antacid and laxative), sodium or potassium citrate are used in medicine. Ferric ammonium citrate finds use in making blue prints.

Aromatic Carboxylic Acids

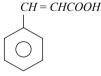
Aromatic acid contain one or more carboxyl group (COOH) attached directly to aromatic nucleus.



Aromatic acid containing-COOH group in the side chain, they are considered as aryl substituted aliphatic acid.

Examples

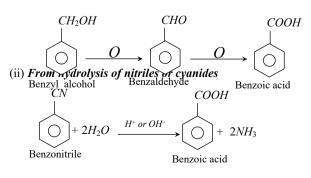


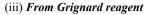


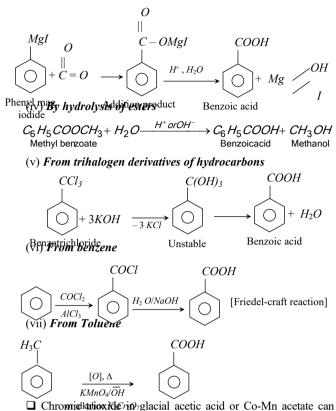
Cinnamic acid

Phenyl acetic acid Benzoic Acid

- (1) Methods of Preparation
- (i) From oxidation of Benzyl alcohol [Laboratory method]

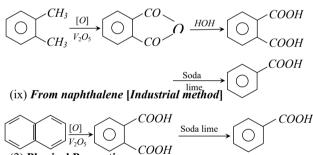






 \Box Chrownekthowfde morglacial acetic acid or Co-Mn acetate callo be taken in place of alkaline $KMnO_4$.

(viii) From o-xylene [Industrial method]



(2) Physical Properties

(i) It is a white crystalline solid.

(ii) It has m.p. 394 K.

(iii) It is sparingly soluble in cold water but fairly soluble in hot water, alcohol and ether.

(iv) It has a faint aromatic odour and readily sublimes and is volatile in steam.

(3) Acidity of Aromatic Carboxylic Acid : Aromatic acid dissociates to give a carboxylate anion and proton.

$$C_6H_5COOH \Rightarrow C_6H_5COO+H^+$$

Since the carboxylate anion (*ArCOO*) is resonance stabilised to a greater extent than the carboxylic acid (*ArCOOH*).

$$\begin{array}{cccc} O & O^{-} & O & O^{-} \\ \parallel & & & \mid & \uparrow & \\ Ar - C - OH \leftrightarrow Ar - C = OH & Ar - C - O^{-} \leftrightarrow Ar - C = O \end{array}$$

| Resonancein carboxylicacid | Resonancein carboxylate anion |
|-------------------------------|--------------------------------|
| Non - equivalent structureand | [Equivalentstructureand hence] |
| hencelessstable | more stable |

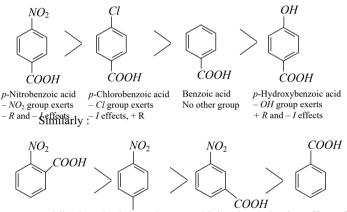
Effect of Substituents on Acidity : The overall influence of a substituent on acidity of substituted benzoic acids is due to two factors.

(i) **Inductive effect :** If the substituent exerts–I effect, it increases the acidity of carboxylic acids, while if it exerts + I effect it decreases the acidity. Inductive effect affects all positions, *i.e.*, o–, m– and p–.

(ii) **Resonance effect :** Like inductive effect, if the resonance producing group exerts minus effect *i.e.*, if it withdraws electrons, it increases the strength of the benzoic acid. Similarly, if the group causes +R effect it decreases the acidity of benzoic acid. However, remember that resonance effect affects only *o*- and *p*- positions. Thus if resonance producing group is present in the *m*-position it will not exert its effect.

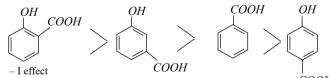
In case resonance and inductive effects both operate in the molecule, resonance effect being stronger overpowers the inductive effect.

Thus on the above basis, the following order of acidity can be explained.



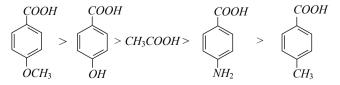
Acidity is only ductoold the model of the $-NO_2$ group (resonance does not affect the *m*-position) while in the *p*-isomer acidity is due to electron withdrawing inductive as well as resonance effect.

The acidity of the three isomers of hydroxybenzoic acids follows the following order.



+ M effection only the M-diffection of th

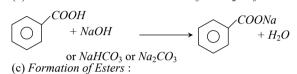
Acidic character among benzoic acids having different electron releasing group.



(4) Chemical Properties :

- (i) Reactions of carboxylic group
- (ii) Reactions of aromatic ring
- (i) Reactions of Carboxylic Group
- (a) Reaction with metals

(b) Reaction with Alkalies Or NaHCO₃ Or Na₂CO₃:

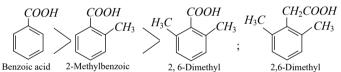


Aromatic acid (benzoic acid) having no group in its ortho positions can be readily esterified with alcohol in presence of a mineral acid.



In presence of ortho substituent the rate of esterification is greatly decreased due to steric effect.

The esterification of the various benzoic acids :



The substituted phenylacetile and for the substituted phenylacetile H^{i} and H^{i} a

The *ortho*-substituted benzoic acids can be easily esterified by treating the silver salt of the acid with alkyl halides, *i.e.*,

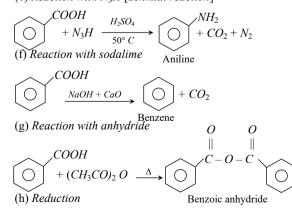
$$H_{3}C \xrightarrow{COOH} H_{3}C \xrightarrow{COOAg} H_{3}C \xrightarrow{COOC_{2}H_{5}} H_{3}C \xrightarrow{CH_{3}} H_{3}C \xrightarrow{CH_{3}}$$

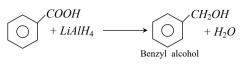
This is due to the fact that in such cases the attack of the alkyl group of the alkyl halides is on the oxygen atom of the -COOH group but not on the sterically hindered carbon atom.

(d) Formation of acid chloride

$$\bigcirc COOH \\ + PCl_5 \text{ or } SOCl_2 \longrightarrow \bigcirc + POCl_3 + HCl_3$$

(e) Reaction with N₃H [Schmidt reaction] Chloride





(i) Decarboxylation

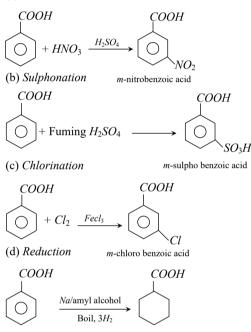
$$\bigcirc +HCOOH \xrightarrow{MnO} \bigcirc +CO_2 + H_2O$$

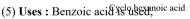
(j) Hunsdiecker reaction :

$$\begin{array}{ccc} C_6 H_5 COOAg & +X_2 & \xrightarrow{\text{ in } CCl_4} & C_6 H_5 - X + CO_2 \uparrow + AgX \\ \text{Silver berzoate } (B_{l_2} \text{ or } Cl_2) & \xrightarrow{\text{ heat}} & \text{ Phenylhalide} \end{array}$$

(ii) Reactions of Aromatic Ring

(a) Nitration





(i) in medicine in the form of its salts especially as urinary antiseptic.

(ii) As sodium benzoate for preservation of food such as fruit juices, tomato ketchup, pickles etc.

(iii) In the preparation of aniline blue.

(iv) In treatment of skin diseases like eczema.

(6) General Tests

(i) Benzoic acid dissolves in hot water but separates out in the form of white shining flakes on cooling.

(ii) It evolves CO_2 with sodium bicarbonate, *i.e.*, it gives effervescence with sodium carbonate.

(iii) Neutral ferric chloride gives a buff coloured precipitate.

(iv) When warmed with ethyl alcohol and a little conc. H_2SO_4 , a fragrant odour of ethyl benzoate is obtained.

(v) When heated strongly with soda lime, benzene vapours are evolved which are inflammable.

Cinnamic Acid [β -Phenyl acrylic acid] CH = CH - COOH

(1) Methods of Preparation

(i) By Perkin's reaction

$$C_6H_5CHO + (CH_3CO)_2O \xrightarrow{CH_3COONa}_{180^{\circ}C}$$

$$C_6H_5CH = CHCOOH + CH_3COOH$$

(ii) By Claisen condensation

 $C_6H_5CHO+CH_3COOC_2H_5$

$$C_{6}H_{5}CH = CHCOOC_{2}H_{5} \xrightarrow{H_{2}O} \xrightarrow{H_{2}O} \xrightarrow{H_{2}O} \xrightarrow{H_{2}O} \xrightarrow{H_{2}O} \xrightarrow{H_{2}OH} \xrightarrow{H_{$$

(iii) By knoevenagel reaction

$$C_{6}H_{5}CHO + CH_{2}(COOH)_{2} \xrightarrow[heat]{} \overset{NH_{3}}{\xrightarrow[heat]{}} \\ C_{6}H_{5}CH = CHCOOH + CO_{2} + H_{2}O$$

(iv) Industrial method

 $\begin{array}{c} C_{6}H_{5}CHCl_{2}+ \\ H_{2}CHCOONa & \xrightarrow{200^{\circ}C} \\ \text{Benzalchloride} & \text{Sodium acetate} \end{array} \xrightarrow{200^{\circ}C} C_{6}H_{5}CH = CHCOOH + \\ NaCl + HCl & \text{Sodium acetate} \end{array}$

(2) Physical Properties

(i) It is a white crystalline solid and its melting point $133^{\circ}C$.

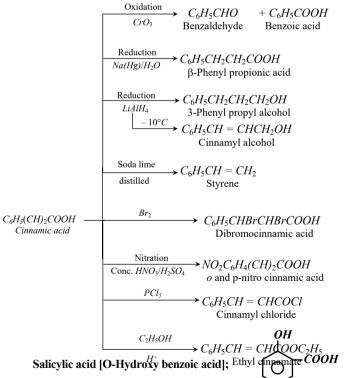
(ii) It is sparingly soluble in water.

(iii) It exhibits geometrical isomerism.

$$\begin{array}{cccc} C_6 H_5 - C - H & C_6 H_5 - C - H \\ \parallel & \parallel \\ H - C - COOH \\ \text{Trans-form} \\ (Cinnamic acid) & (Allo cinnamicacid) \end{array}$$

Cinnamic acid (stable form) occurs in nature both free and as esters in balsams and resins.

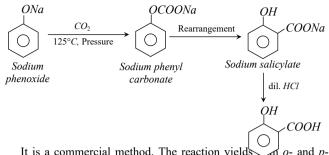
(3) Chemical properties



Salicylic acid is present in many essential oils in the form of esters. Oil of winter green is a methyl ester of salicylic acid.

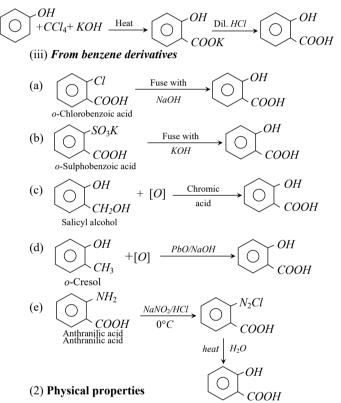
(1) Methods of preparation

(i) Kolbe Schmidt reaction



It is a commercial method. The reaction yield of *o*- and *p*isomers. Salicylic acid is more volatile and separated and steam distillation.

(ii) Reimer-Tiemann reaction



(i) It is a colourless needle shaped crystalline compound.

(ii) Its *m.p.* is 156°*C*.

(iii) It is sparingly soluble in cold water but readily soluble in hot water, alcohol, ether and chloroform.

(iv) It is steam volatile.

(v) It is poisonous in nature. However, its derivative used in medicine internally and externally as antipyretic and antiseptic.

(3) Chemical properties

(i) Reaction with Na₂CO₃, NaHCO₃ or NaOH O = O = OH O = OHSalicylic acid $Aq. Na_2CO_3 = OH$ Mono|sodium salicylate Aq. NaOH OHCOONa

ONa Disodium salicylate

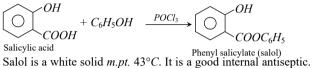
(ii) Reaction with alcohols or phenols

$$\bigcirc OH + CH_3OH \xrightarrow{HCl(gas)} OH + H_2O$$

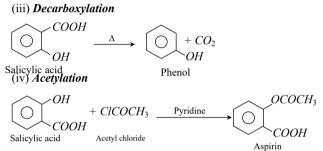
$$\bigcirc OH + CH_3OH \xrightarrow{HCl(gas)} OH + H_2O$$

$$\bigcirc OH + CH_3OH \xrightarrow{HCl(gas)} OH + H_2O$$

Methyl^{ic}saffcylate is an oily liquid (*Methyl* salicylate of *winter green*) with pleasant material. It is also used in medicine in the treatment of rheumatic pain and as a remedy for aches, sprains and bruises. It is used in perfumery and as a flavouring. It is used for making of iodex.

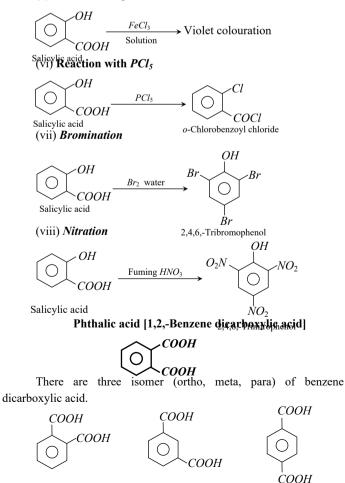


Salol is a white solid *m.pt.* $43^{\circ}C$. It is a good internal antiseptic. It is used in making of toothpastes. Salol absorbs ultraviolet light and its main use now is sun-screening agent and stabiliser of plastics.



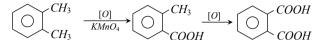
□ Aspirin is a white solid, melting point 135°®! ∰itsliused) as antipyretic and pain killer (analgesic action).

(v) Reaction with ferric chloride solution

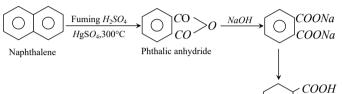


(1) Methods of preparation

(i) By the oxidation of o-xylene :



^{o-Xylenc} (ii) **From naphthalene** (Toluic acid (iii) **From naphthalene** (Industrial method) : It is known as aerial oxidation.



СООН

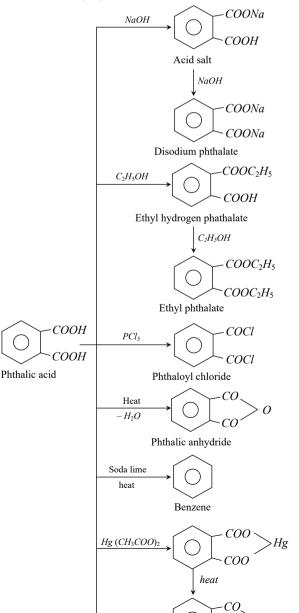
(2) Physical properties

(i) It is colourless crystalline compound.

(ii) Its melting point is not sharp (195–213°C).

(iii) It is sparingly soluble in cold water but soluble in hot water, alcohol, ether, benzene etc.

(3) Chemical properties



Benzene-1,2-dicarboxylic acid (Phthalic acid)

Benzene-1,3-dicarboxylic acid (Isophthalic acid)

Benzene-1,4-dicarboxylic acid (Terphthalic acid)

(4) Uses : It is used in the manufacture of plastics, dyes and other compounds such as phthalic anhydride, phthalimide, anthraquinone and fluorescein etc.

Acid derivatives

The compounds which are obtained by replacing the -OH of the carboxylic group by other atoms or groups such as X^- , $-NH_2$, -OR and O-C-R are known as acid derivatives.

• R - C- group is common to all the derivatives and is known

as acyl group and these derivatives are termed as acyl compound.

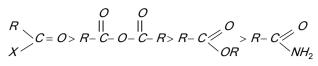
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|-----------------------------------|----------------|-----------------|
| The important | derivatives an | e given helow · |
| • The important | ucrivatives ar | |

| Group replacing – <i>OH</i> | Name | Structure |
|--|-------------|--|
| (<i>X</i> = <i>F</i> , <i>Cl</i> , <i>Br</i> , <i>l</i>) | Acyl halide | $\begin{array}{c} O \\ \parallel \\ R - C - X \end{array}$ |
| - <i>NH</i> ₂ | Amide | 0 R- C- NH ₂ |
| - <i>OR</i> | ester | 0 <i>R - C - OR</i> (<i>R</i> may be <i>R</i>) |
| - <i>00CR</i> | anhydride | 0 0 R-C-O-C-R |

Reactivity

Acyl derivatives are characterised by nucleophilic substitution reactions.

The relative reactivities of various acyl compounds have been found to be in the following order:



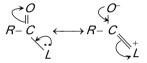
Out of acid halides, the acid chlorides are more important ones.

The overall order of reactivity can be accounted for in terms of the following three factors:

(i) Basicity of the leaving group (ii) Resonance effects and (iii) Inductive effects.

(i) **Basicity of the leaving group :** Weaker bases are good leaving groups. Hence, the acyl derivatives with weaker bases as leaving groups are more reactive. Chloride ion is the weakest base while – NH_2 is the strongest base. Thus, acyl chlorides are most reactive and amides are least reactive.

(ii) *Resonance effect*: The leaving group in each case has an atom with lone pair of electrons adjacent to the carbonyl group. The compound exists, therefore, as a resonance hybrid.



This makes the molecule more stable. The greater the stabilization, the smaller is the reactivity of the acyl compound.

However, acyl chlorides are least affected by resonance. Due to lower stabilization, the acid chlorides are more reactive as the loss of -CI is easier. Greater stabilization is achieved by resonance in esters and amides and thus, they are less reactive.

(iii) *Inductive effect*: Higher the -I effect, more reactive is the acyl compound. Inductive effect of oxygen in ester is greater than nitrogen in amide, hence ester is more reactive than an amide.

Acyl Halides
$$R - C = C$$

where *R* may be alkyl or aryl group.

(1) Methods of Preparation

(i) From carboxylic acid :

 $RCOOH + PCl_5 \rightarrow RCOCI + POCl_3 + HCl_3$

 $3RCOOH + PCI_3 \rightarrow 3RCOCI + H_3PO_3$

(ii) Industrial method : By distilling anhydrous sodium acetate

 $3CH_3COONa + PCI_3 \xrightarrow{heat} 3CH_3COCI + Na_3PO_3$

 $\begin{array}{ccc} 2CH_3 COONa + POCI_3 & \xrightarrow{\text{heat}} 2CH_3 COCI + NaPO_3 + NaCI \\ \text{Sodium acetate} & \text{Acetylchloride} \end{array}$

$$\begin{array}{ccc} (CH_3 COO)_2 Ca+ SO_2 CI_2 & \xrightarrow{\text{heat}} 2CH_3 COCI+ CaSO_2 & \xrightarrow{\text{heat}$$

(iii) With thionyl chloride :

$$RCOOH + SOCI_2 \rightarrow RCOCI + SO_2 + HCI$$

This is the best method because SO_2 and HCl are gases and easily escape leaving behind acyl chloride.

(2) **Physical properties :** The lower acyl chloride are mobile, colourless liquid while the higher members are coloured solids.

Acyl chloride have very pungent, irritating order and are strong lachrymators (tears gases)

They fume in air due to the formation of hydrochloric acid by hydrolysis.

They are readily soluble in most of the organic solvent. Acyl chloride don't form intermolecular hydrogen bonding. Therefore, their boiling points are lower than those of their parent acids.

(3) Chemical properties

 $\begin{array}{c} \textit{CI}^- + \textit{H}^+ \rightarrow \textit{HCI} \\ (i) \textit{Hydrolysis}: & \textit{CH}_3\textit{COCI} + \textit{HOH} \rightarrow \textit{CH}_3\textit{COOH} + \textit{HCI} \\ & \text{Acetylchloride} & \text{Aceticacid} \\ \end{array}$

 $\begin{array}{c} C_6 \ H_5 \ COCl + \ H_2 \ O \rightarrow \ C_6 \ H_5 \ COOH + \ H_2 \ O \\ \text{Benzoylchloride} & \text{Benzoicacid} \end{array}$

(ii) Reaction with alcohols (alcoholysis) $CH_3COCI + CH_3CH_2OH \rightarrow CH_3COOCH_2CH_3 + HCI$ Ethyl acetate

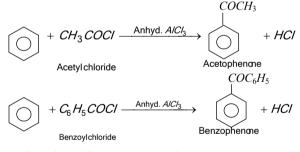
 $\begin{array}{c} C_{6}H_{5}COCI + C_{2}H_{5}OH & \xrightarrow{\text{aq NaOHor}} \\ \text{Benzoylchloride Ethyl alcohol} & \xrightarrow{\text{Pyridine}} & C_{6}H_{5}COOC_{2}H_{5} + HCI \\ & \text{Ethyl benzoate} \end{array}$

This reaction is called Schotten Baumann reaction.

(iii) Reaction with salts of carboxylic acid

$$CH_3COCI + CH_3COO^- Na^+ \xrightarrow{\text{Pyridine}} CH_3 \stackrel{O}{\underset{C}{\overset{\parallel}{\longrightarrow}}} CH_3 \stackrel{O}{\underset{C}{\overset{\parallel}{\longrightarrow}}} CH_3 \stackrel{O}{\underset{Aceticanhydride}{\overset{\parallel}{\longrightarrow}}} CH_3 \stackrel{O}{\underset{\Lambda}{\longrightarrow}} CH_3 \stackrel{O}{\underset{\Lambda}{\longrightarrow} CH_3 \stackrel{O}{\underset{\Lambda}{\longrightarrow}} CH_3 \stackrel{O$$

(iv) *Reaction with benzene* (acylation) : This reaction is called friedel craft reaction.



However, acyl chlorides react with amines to form substituted amides.

$$O$$

$$||$$

$$CH_3COCI + H_2NC_2H_5 \rightarrow CH_3 C - NH - C_2H_5$$
N-Ethyl acetamide
$$CH_3COCI + (C_2H_5)_2 NH \rightarrow CH_3CON(C_2H_5)_2 + HCI$$
N, N-Diethyl acetamide
(vi) Reduction :

$$CH_3COCI \xrightarrow{LiAIH_4 \text{ or}} CH_3CH_2OH$$

$$\xrightarrow{NaBH_4} Ethanol(Primary alcohol)$$

$$CH_3COCI + H_2 \xrightarrow{Pd/BaSO_4} CH_3CHO + HCI$$

This reaction is called **Rosenmund reaction**.

(vii) *Reaction with organocadmium compounds* (formation of ketones)

$$\begin{array}{c} 2\textit{CH}_{3}\textit{COCl} + (\textit{CH}_{3})_{2}\textit{Cd} \rightarrow 2\textit{CH}_{3}\textit{COCH}_{3} + \textit{CdCl}_{2}\\ & \text{Dimethyl} & \text{Acetone} \\ & \text{Cadmium} \end{array}$$

$$2C_6H_5COCI + (CH_3)_2Cd \rightarrow 2C_6H_5COCH_3 + CdCI_2$$

Acetophenane

(viii) Reaction with diazomethane

$$CH_{3} - C - CI + 2CH_{2} - N = N \rightarrow CH_{3} - C - CH - N = N$$
Diazomethane
$$O$$

$$H_{2}O$$

$$H_{2}O$$

$$H_{3}CH_{2}C - OH$$

(ix) Reaction with water

$$CH_3COCI \xrightarrow{AgNO_3 \mid H_2O} CH_3COOH + AgCI + HNO_3$$

(x) Reaction with chlorine

$$CH_3COCI + CI_2 \xrightarrow{\text{Red }P} CI - CH_2 - CO - CI + HCI \\ Mono-\alpha - chloroaceyl chloride$$

(xi) Reaction with Grignard reagent

$$\begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ Methyl \ \overline{magnesiumiodide} \qquad Acetone \qquad \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \swarrow \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \H \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \H \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \H \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \H \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \H \\ \end{array} \\ \begin{array}{c} CH_3CO[\overline{C/+} \ \overline{IMg}]CH_3 \rightarrow CH_3COCH_3 + \ Mg \H \\ \end{array} \\ \end{array}$$
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(xii) Reaction with KCN

$$\begin{array}{c} CH_3COCI+\ KCN \rightarrow CH_3COCN \xrightarrow{H_2O} CH_3COCOOH\\ Acetyl cyanide \end{array} \begin{array}{c} Pyruvic acid \end{array}$$

(xiii) Reaction with Salicylic acid

$$\bigcirc OH \\ COOH^{+} CIOCCH_{3} \rightarrow \bigcirc OOCCH_{3} \\ COOH^{+} HCl$$
Salicylic
acid Acetyl salicylicacid
(Aspirin)

(xiv) Reaction with ether

ŝ

$$\frac{CH_3COCI + C_2H_5OC_2H_5}{\text{Diethyl ether}} - \frac{2nCI_2}{\text{anhy.}}$$

 $\begin{array}{c} CH_3COOC_2H_5 + C_2H_5CI\\ \text{Ethyl acetate} & \text{Ethyl chloride} \end{array}$

(xv) *Reaction with sodium peroxide* (Peroxide formation)

$$\begin{array}{c} O & O \\ \parallel \\ 2CH_3 - C - CI + \begin{array}{c} + & - & - + \\ NaO - ONa \rightarrow \begin{array}{c} O \\ H_3 C - O - O - C - C - CI \\ Acetyl chloride \end{array} + 2 NaCI \\ Acetyl peroxide \end{array}$$

$$\begin{array}{ccc} CH_3COCI+ & H_2NOH \rightarrow & CH_3CONHOH + HCH \\ & & & & \\ Hydroxyl & & & \\ amine & & & (hydroxylamine \\ & & & (hydroxamic acid) \end{array}$$

 $\begin{array}{c} {\it CH_3COCl+} {\it H_2NNH_2} \rightarrow {\it CH_3CONHNH_2+} {\it HCl} \\ {\it Hydrazine} & {\it Acetyl hydrazine} \end{array}$

(4) Uses

(i) As an acetylating agent.

(ii) In the estimation and determination of number of hydroxyl and amino groups.

(iii) In the preparation of acetaldehyde, acetic anhydride, acetamide, acetanilide, aspirin, acetophenone etc.

Acid Amides
$$R - C = 0$$

where, $R = -CH_3, -CH_2CH_3, -C_6H_5$

(1) Methods of preparation

(i) Ammonolysis of acid derivatives

$$\begin{array}{c} CH_3 COCI + 2NH_3 \rightarrow CH_3 CONH_2 + NH_4 CI \\ Acetamide \end{array}$$

$$(CH_3CO)_2O + 2NH_3 \rightarrow CH_3CONH_2 + CH_3COONH_4$$

Acetamide Amm. acetate

 $\begin{array}{c} C_6 \ H_5 \ COCl + \ NH_3 \rightarrow C_6 \ H_5 \ CONH_2 + \ HCl \\ \text{Benzoylchloride} & \text{Benzamide} \end{array}$

(ii) From ammonium salts of carboxylic acids (Laboratory Method)

$$CH_3COONH_4 \xrightarrow{\text{Heat}} CH_3CONH_2 + H_2O$$
Acetamide

 \Box Ammonium acetate is always heated in presence of glacial acetic acid to avoid the side product (*CH*₃*COOH*).

(iii) By partial hydrolysis of alkyl cyanide :

 $CH_3C \equiv N \xrightarrow{\text{Conc.}HCl} CH_3CONH_2$ $\xrightarrow{H_2O/OH} CH_3CONH_2$ Acetamide

(iv) By heating carboxylic acid and urea

$$\begin{array}{ccc} H_2N-C-NH_2+R-C-OH \xrightarrow{heat} R-C-NH_2+CO_2+NH_3\\ \parallel & & \parallel \\ O & O \\ Amide \end{array}$$

(2) Physical properties

(i) *Physical state* : Formamide is a liquid while all other amides are solids.

(ii) **Boiling points :** Amides have high boiling points than the corresponding acids.

| Acetamide | Boiling points 494 K |
|--------------|----------------------|
| Acetic Acid | Boiling points 391 K |
| Benzamide | Boiling points 563 K |
| Benzoic acid | Boiling points 522 K |

The higher boiling points of amides is because of intermolecular hydrogen bonding

(iii) **Solubility** : The lower members of amide family are soluble in water due to the formation of hydrogen bonds with water.

(3) Chemical properties

(i) *Hydrolysis*

$$\begin{array}{l} CH_{3}CONH_{2} + H_{2}O \xrightarrow{\text{Slowly}} CH_{3}COOH + NH_{3} \\ \\ CH_{3}CONH_{2} + H_{2}O + HCI \xrightarrow{\text{Rapidly}} CH_{3}COOH + NH_{4}CI \\ \\ CH_{3}CONH_{2} + NaOH \xrightarrow{\text{Farmore rapidly}} CH_{3}COONa + NH_{3} \end{array}$$

(ii) Amphoteric nature (Salt formation)

It shows feebly acidic as well as basic nature.

$$\begin{array}{c} {\it CH}_3{\it CONH}_2 + {\it HCl}({\it conc}) \rightarrow {\it CH}_3{\it CONH}_2.{\it HCl}\\ {\it Acetamidehydrochloide}\\ {\it (only stable in aqueoussolution)} \end{array}$$

$$\begin{array}{c} 2CH_3CONH_2 + \begin{array}{c} HgO \\ \text{Acetamide} \end{array} \rightarrow (CH_3CONH)_2Hg + H_2O \\ \begin{array}{c} \text{Mercuric} \\ \text{Oxide} \end{array} \end{array}$$

$$CH_3CONH_2 + Na \xrightarrow{\text{Ether}} CH_3CONHNa + \frac{1}{2}H_2$$

Sodium acetamide

(iii) **Reduction**

$$\begin{array}{c} CH_3CONH_2 + 4[H] \xrightarrow{\text{LiAIH}_4} CH_3CH_2NH_2 + H_2O \\ \text{Acetamide} & \text{Ethylamine} \end{array}$$

(iv) **Dehydration**

$$\begin{array}{c} CH_3CONH_2 \xrightarrow{P_2O_5} CH_3C \equiv N + H_2O \\ \text{Acetamide} & \text{Methyl cyanide} \end{array}$$

$$\begin{array}{c} C_{6}H_{5}CONH_{2} \xrightarrow{P_{2}O_{5}} & C_{6}H_{5}C \equiv N + H_{2}O \\ \text{Benzamide} & \text{heat} & \text{Phenyl cyanide} \end{array}$$

$$C_6 H_5 CONH_2 \xrightarrow{SOCI_2} C_6 H_5 C \equiv N$$
Phenyl cyanide

(v) Reaction with nitrous acid

$$CH_{3}CONH_{2} + HONO \xrightarrow{NaNO_{2} \mid HCl} CH_{3}COOH + N_{2}$$
Aceticacid

$$C_{6}H_{5}CONH_{2} + HONO \xrightarrow{NaNO_{2} \mid HCl} C_{6}H_{5}COOH$$
Benzoic acid

 $+N_{2}+H_{2}O$

 $+H_2O$

(vi) Hofmann bromamide reaction or Hofmann degradation : This is an important reaction for reducing a carbon atom from a compound, *i.e.*, $-CONH_2$ is changed to $-NH_2$ group.

$$\begin{array}{c} CH_3 CONH_2 \xrightarrow{Br_2} CH_3 NH_2 \\ Acetamide \end{array} \xrightarrow{AOHor KOH} CH_3 NH_2 \\ \text{Methyl amine(p-)} \end{array}$$

This reaction occurs is three steps:

$$CH_{3} - C - NH_{2} + Br_{2} + KOH \rightarrow CH_{3}CONHBr + KBr + H_{2}O$$

$$CH_{3} - C - NHBr + KOH \rightarrow CH_{3}NCO + KBr + H_{2}O$$

$$CH_{3}NCO + 2KOH \rightarrow CH_{3}NH_{2} + K_{2}CO_{3}$$

$$Methylamine$$

$$\overline{CH_3CONH_2 + Br_2 + 4KOH} \rightarrow CH_3NH_2 + 2KBr + K_2CO_3 + 2H_2O_3$$

 \Box In this reaction a number of intermediates have been isolated; *N*-bromamides, *RCONHBr*, salts of these bromamides

[*RCONBr*⁻] K^+ ; Isocyanates, *RNCO*.

□ Nitrene rearranges to form isocyanate. (vii) *Action with alcohol* :

$$CH_{3}CONH_{2} + CH_{3}OH \xrightarrow{HCI} CH_{3}COOCH_{3} + NH_{4}CI_{3}COOCH_{3} + N$$

(viii) Reaction with grignard reagent

$$CH_{3} - Mg - Br + CH_{3} - CONH_{2} \rightarrow CH_{4} + CH_{3} - CONH - MgBr$$
$$\bigcup CH_{3}MgBr$$

$$\begin{bmatrix} OH \\ CH_3 - C - NH_2 \\ CH_3 \\ Unstable \end{bmatrix} \xrightarrow{H_2O/H^+} CH_3 - CH_3 - CH_3 \\ Hydrolysis \\ CH_3 \\ CH_3 \\ CH_3 - C - CH_3 \\ Acetone \end{bmatrix}$$

(4) Uses

(i) In organic synthesis. The compounds like methyl cyanide, Methylamine and ethylamine can be prepared.

(ii) In leather tanning and paper industry.

(iii) As a wetting agent and as soldering flux.

Amides such as dimethyl formamide (DMF), dimethyl acetamide (DMA) are used as solvents for organic and inorganic compounds.

These are the most important class of acid derivatives and are widely distributed in nature in plants, fruits and flowers.

(1) Methods of preparation

(i) From carboxylic acid [Esterification] : Laboratory method.

$$\begin{array}{c} O & O \\ R - C - \left[\overline{OH} + \overline{H} \right] OR \xrightarrow{H^+} R - C - OR + H_2 O \\ \hline Ester \\ CH_3 COOH + CH_2 N_2 & \xrightarrow{\text{Ether}} CH_3 COOCH_3 + N_2 \\ Aceticacid & Diazomethane & Methyl acetate \\ \end{array}$$

Ether $\rightarrow C_6 H_5 COOCH_3 + N_2$ Methyl berzoate $C_6H_5COOH+CH_2N_2$ Benzoicacid Diazomethane

With diazomethane is the best method.

From acid chloride or acid anhydrides (ii) $\begin{array}{c} \mathcal{CH}_3\mathcal{CO}_{\boxed{}}\overrightarrow{\mathcal{CI}_{+}}\overrightarrow{H_{}}\mathcal{OC}_2\mathcal{H}_5 \rightarrow \mathcal{CH}_3\mathcal{COOC}_2\mathcal{H}_5 + \mathcal{HCI}\\ \text{Acetyl chloride} \quad \text{Ethyl alcohol} \quad \text{Ethyl acetate} \end{array}$

CH3CO~ $> O+ CH_3CH_2OH \rightarrow CH_3COOCH_2CH_3+ CH_3COOH$ Ethyl acetate CH₃CO Aceticanhydride Ethyl alcohol

$$\begin{array}{ccc} C_6 \hspace{0.5mm} H_5 \hspace{0.5mm} CO[\overline{CI} + \overline{H}] \hspace{0.5mm} OC_2 \hspace{0.5mm} H_5 \hspace{0.5mm} \rightarrow \hspace{0.5mm} C_6 \hspace{0.5mm} H_5 \hspace{0.5mm} COOC_2 \hspace{0.5mm} H_5 \hspace{0.5mm} + \hspace{0.5mm} HCI \hspace{0.5mm} \\ \text{Benzoyl chloride} \hspace{0.5mm} \text{Ethyl alcohol} \hspace{0.5mm} \rightarrow \hspace{0.5mm} C_6 \hspace{0.5mm} H_5 \hspace{0.5mm} COOC_2 \hspace{0.5mm} H_5 \hspace{0.5mm} + \hspace{0.5mm} HCI \hspace{0.5mm} \\ \text{Ethyl benzoate} \hspace{0.5mm} \end{array}$$

(iii) From alkyl halide :

$$\begin{array}{c} C_2H_5Br+CH_3COOAg \rightarrow CH_3COOC_2H_5+AgBr\\ \text{Ethyl bromide} & \text{Silveracetate} & \text{Ethyl acetate} \end{array}$$

$$\begin{array}{c} CH_{3}-O-CH_{3}+CO & \xrightarrow{BF_{3}} \\ \text{Methoxy methane} & 350 \ \text{K} \end{array} \xrightarrow{CH_{3}COOCH_{3}} \\ \text{Methyl acetate} \end{array}$$

(v) From Tischenko reaction :

$$\begin{array}{c} CH_3 - \underbrace{C-H}_{H} + O = \underbrace{C-CH_3}_{H} \xrightarrow{Al(OC_2H_5)_3} O CH_3 - \underbrace{C-OC_2H_5}_{H} \\ O \end{array}$$

(2) Physical properties

(i) Physical state and smell : Esters are colourless liquids (or solids) with characteristic fruity smell. Flavours of some of the esters are listed below :

| | Ester | Flavour | Ester | Flavour |
|--|-------|---------|-------|---------|
|--|-------|---------|-------|---------|

| Amyl acetate | Banana | Isobutyl formate | Raspberry |
|----------------|---------|------------------|-----------|
| Benzyl acetate | Jasmine | Ethyl butyrate | Pineapple |
| Amyl butyrate | Apricot | Octyl acetate | Orange |

(ii) Solubility : They are sparingly soluble in water but readily soluble in organic solvents such as alcohol, ether etc.

(iii) Boiling points : Their boiling points are lower than the corresponding acids because of the absence of hydrogen bonding. *i.e.*, ethyl acetate = $77.5^{\circ}C$.

(3) Chemical properties

(i) *Hydrolysis* :

$$\begin{array}{c} CH_3COOC_2H_5 + H_2O \xrightarrow{\text{dil. acid}} CH_3COOH + C_2H_5OH\\ \text{Ethyl acetate} & \text{Aceticacid} & \text{Ethyl alcohol} \end{array}$$

$$\begin{array}{c} CH_3COOC_2H_5 + NaOH \longrightarrow CH_3COONa + C_2H_5OH\\ \text{Ethyl acetate} & \text{Sod. acetate} & \text{Ethyl alcohol} \end{array}$$

Hydrolysis of ester by alkalies (NaOH) is known as saponification and leads to the formation of soaps

□ This reaction (saponification) is irreversible because a resonance stabilized carboxylate (acetate) ion is formed.

The acid hydrolysis of esters is reversible.

(ii) Reaction with ammonia (ammonolysis):

$$CH_3CO[OC_2H_5 + H] NH_2 \rightarrow CH_3CONH_2 + C_2H_5OH$$

Ethyl acetate Acetamide
(iii) Reduction

$$CH_{3}COOC_{2}H_{5} + 4[H] \xrightarrow[\text{or Nal } C_{2}H_{5}OH] 2C_{2}H_{5}OH$$

$$COOC_{2}H_{5} + 4H \xrightarrow[\text{or Nal } C_{2}H_{5}OH] + C_{2}H_{5}OH$$

Ethyl benzoate

Benzylalcohol

 \Box Reduction in presence of Nal C₂H₅OH is known as Bouveault Blanc reduction.

□ The catalytic hydrogenation of ester is not easy and requires high temperature and pressure. The catalyst most commonly used is a mixture of oxides known as copper chromate $(CuO.CuCr_2O_4)$.

$$R - \overset{||}{C} - OR + 2H_2 \xrightarrow{CuO. CuC_{P_2}O_4} RCH_2OH + ROH$$

(iv) Reaction with PCl₅ or SOCl₂

 $CH_3COOC_2H_5 + PCI_5 \rightarrow CH_3COCI + C_2H_5CI + POCI_3$

$$\begin{array}{c} {\it CH}_3{\it COOC}_2{\it H}_5 + {\it SOCI}_2 \rightarrow {\it CH}_3{\it COOCI} + {\it C}_2{\it H}_5{\it CI} + {\it SO}_2 \\ {\it Acetyl \, chloride} & {\it Ethyl \, chloride} \end{array}$$

$$\begin{array}{c} \mathcal{C}_{6}\mathcal{H}_{5}\mathcal{C}\mathcal{OOC}_{2}\mathcal{H}_{5} + \mathcal{PCI}_{5} \rightarrow \mathcal{C}_{6}\mathcal{H}_{5}\mathcal{C}\mathcal{OCI} + \mathcal{POCI}_{3} + \mathcal{C}_{2}\mathcal{H}_{5}\mathcal{CI}_{5} \\ \text{Ethyl benzoate} \end{array}$$

(v)Reaction with alcohols : On refluxing ester undergoes exchange of alcohols residues.

$$R - C \swarrow O \\ OR + (Excess) \swarrow R - C \swarrow O \\ OR' + ROH$$
$$CH_3COOC_2H_5 + CH_3OH \rightarrow CH_3COOCH_3 + C_2H_5OH$$
$$Ethyl acetate Methyl acetate$$

□ This reaction is known as alcoholysis or trans esterification.

(vi) Reaction with Grignard reagents

$$O \\ CH_{3} - C - OC_{2}H_{5} + CH_{3}MgBr \rightarrow \begin{bmatrix} OMgBr \\ H_{3} - C - OC_{2}H_{5} \\ CH_{3} - C - OC_{2}H_{5} \\ H_{3} \end{bmatrix} \xrightarrow{I_{3}} CH_{3} \\ OMgBr \\ CH_{3} - C - CH_{3} \xleftarrow{CH_{3}MgBr} CH_{3} - C - CH_{3} \xleftarrow{I_{3}} CH_{3} \\ H^{+} H_{2}O \\ OH \\ CH_{3} - C - CH_{3} \\ H^{+} H_{2}O \\ OH \\ CH_{3} - C - CH_{3} \\ CH_{3} \\ GH_{3} \\ 3^{\circ} \text{ alcohol} \end{bmatrix}$$

(vii) Claisen condensation

$$CH_{3} - C - \underbrace{O}_{C_{2}} \underbrace{H_{5}}_{H_{5}} + H_{-} + CH_{2}COOC_{2}H_{5} \xrightarrow{C_{2}H_{5}O^{-}Na^{+}} \\ Ethyl acetate(2 \text{ molecules}) \\ O \\ \parallel \\ O \\ \blacksquare \\$$

$$\begin{matrix} \text{II} \\ \mathcal{CH}_3 - \mathcal{C} - \mathcal{CH}_2 \mathcal{COOC}_2 \mathcal{H}_5 + \mathcal{C}_2 \mathcal{H}_5 \mathcal{OH} \\ \text{Ethyl acetoacetate} \\ (\beta\text{-ketoester}) \end{matrix}$$

(viii) Reaction with hydroxyl amine

$$\begin{array}{c} O \\ H \\ CH_3 - C - OC_2 H_5 + H \\ Hylacetate \end{array} | HNOH \xrightarrow{base} CH_3 - C - NHOH + C_2H_5OH \\ Hydroxyl amine \end{array}$$

(ix) Reaction with hydrazine $CH_3COOC_2H_5 + H_2NNH_2 \rightarrow CH_3CONHNH_2 + C_2H_5OH$

Hydrazine Acid hydrazide

(x) *Halogenation*

 $CH_3COOC_2H_5 + Br_2 \xrightarrow{\text{Red P}} CH_2BrCOOC_2H_5 + HBr$ $\alpha \text{-Bromoethyl acetate}$

(xi) *Reaction with HI*

$$\begin{array}{c} CH_3COOC_2H_5 + HI \rightarrow CH_3COOH + C_2H_5OH \\ \text{Aceticacid} & \text{Ethyl alcohol} \end{array}$$

(4) Uses

(i) As a solvent for oils, fats, cellulose, resins etc.

(ii) In making artificial flavours and essences.

(iii) In the preparation of ethyl acetoacetate.

(5) General Tests

(i) It has sweet smell

(ii) It is neutral towards litmus

(iii) A pink colour is developed when one or two drops of phenolphthalein are added to dilute sodium hydroxide solution. The pink colour is discharged when shaken or warmed with ethyl acetate.

(iv) Ethyl acetate on hydrolysis with caustic soda solution forms two compounds, sodium acetate and ethyl alcohol.

$$CH_3COOC_2H_5 + NaOH \rightarrow CH_3COONa + C_2H_5OH$$

Acid Anhydride
$$\begin{array}{c} CH_3CO > 0 \\ CH_3CO > 0 \end{array}$$
 or $(CH_3CO)_2O$

(1) Method of preparation

(i) From carboxylic acid

$$C_{6}H_{5}CO[OH+H]O-C-R \xrightarrow{Quartztube}{Porcelain} R \xrightarrow{C} C_{-}O-C-R_{+}H_{2}O$$

$$C_{6}H_{5}CO[OH+H]OOCC_{6}H_{5} \xrightarrow{P_{4}O_{10}}{heat}$$

$$C_{6}H_{5}-C-O-C-C_{6}H_{5}+H_{2}O$$
Benzoicanhydride

(ii) From carboxylic acid salt and acyl chloride [Laboratory method]

$$CH_{3}COONa + CH_{3}COCI \xrightarrow{Py} CH_{3}COOCOCH_{3} + NaCI$$
Aceticanhydride
$$C_{6}H_{5}COONa + C_{6}H_{5}COCI \xrightarrow{Py} C_{6}H_{5}COOCOC_{6}H_{5}$$
Benzoicanhydride

+ NaCl

(iii) From acetylene

$$CH + 2CH_{3}COOH \xrightarrow{HgSO_{4}} | CH_{3} \xrightarrow{CH_{3}} \xrightarrow{Distill} \stackrel{heat}{\longrightarrow} CH(OOCCH_{3})_{2} \xrightarrow{Distill} \xrightarrow{heat} CH_{3}CHO + \frac{CH_{3}CO}{CH_{3}CO} \xrightarrow{O}_{Aceticanhydride}$$

(iv) From acetaldehyde :

$$CH_3CHO + O_2 \xrightarrow[\text{acetate}]{\text{cobalt}} 2CH_3 - C - O - O - H$$

 $\rightarrow (CH_3CO)_2O + H_2O$

(2) Physical properties

(i) *Physical state*: Lower aliphatic anhydrides are colourless liquids with sharp irritating smell. The higher members of the family as well as the aromatic acid anhydrides are solids in nature.

(ii) *Solubility* : They are generally insoluble in water but are soluble in the organic solvents such as ether, acetone, alcohol, etc.

(iii) *Boiling points*: The boiling points of acid anhydrides are higher than those of carboxylic acids because of the greater molecular size.

(3) Chemical Properties

(i) *Hydrolysis* :

$$\begin{array}{c} O & O \\ \parallel & \parallel \\ CH_3 - C - O - C - C - CH_3 + H_2O \rightarrow 2CH_3COOH \\ \text{Aceticanhydride} & \text{Aceticacid} \end{array}$$

(ii) Action with ammonia

$$(CH_3CO_{12}O + 2NH_3 \rightarrow CH_3CONH_2 + CH_3COONH_4$$

Acetamide Armm. acetate

(iii) *Acetylation* : Acetic anhydride react with compound having active hydrogen.

$$(CH_3CO)_2O + \underset{\text{Ethyl alcohol}}{C_2H_5OH} \rightarrow CH_3COOC_2H_5 + CH_3COOH$$

$$(CH_3CO)_2O + \underset{\text{Ethyl amine}}{H_2NC_2H_5} \rightarrow CH_3CONHC_2H_5 + CH_3COOH$$

 $(CH_3CO)_2O + HN(C_2H_5)_2 \rightarrow CH_3CON(C_2H_5)_2 + CH_3COOH$ Diethylamine N, N-Diethyl acetamide

$$(CH_3CO)_2O + H_2NC_6H_5 \rightarrow CH_3CONHC_6H_5 + CH_3COOH_{Acetanilide}$$

$$(CH_3CO_{2}O + \bigcirc OH \\ COOH \rightarrow \bigcirc OOCCH_3 + CH_3COOH$$

Salicylic acid Acetyl salicylic acid (Aspiriin)

 $(CH_3CO)_2O + HCI \rightarrow CH_3COCI + CH_3COOH$

(v) *Reaction with chlorine*

 $(CH_3CO)_2O + CI_2 \rightarrow CH_3COCI + CH_2CICOOH \\ \text{Acetyl chloride} \qquad \begin{array}{c} CH_2CICOOH \\ \text{Monochloroacetic} \\ \text{acid} \end{array}$

(vi) Reaction with PCl₅

 $(CH_3CO)_2O + PCI_5 \rightarrow 2CH_3COCI + POCI_3$

(vii) Friedel craft's reaction

$$(CH_3CO)_2O + C_6H_6 \xrightarrow{AICI_3} C_6H_5COCH_3 + CH_3COOH$$

Benzene Acetophenone

(viii) Reaction with acetaldehyde

$$\begin{array}{c} (CH_3CO)_2O+ CH_3CHO \rightarrow CH_3CH(OOCCH_3)_2\\ \text{Acetaldehyde} \qquad \text{Ethylideneacetate} \end{array}$$

(ix) Reduction

$$(CH_3CO)_2O \xrightarrow{LiAIH_4} CH_3CH_2OH$$

Ether Ethyl alcohol

(x) Action with ether :

$$CH_{3}CO[O.COCH_{3} + C_{2}H_{5}] O - C_{2}H_{5} \rightarrow 2CH_{3}COOC_{2}H_{5}$$

Ethyl acetate Ethyl acetate

(xi) Action with N₂O₅

$$CH_3COOCOCH_3 + N_2O_5 \rightarrow CH_3 - C - O - N \subset C$$

(4) Uses : Acetic anhydride is used

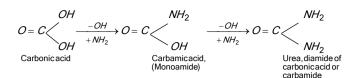
(i) as an acetylating agent.

(ii) For the detection and estimation of hydroxyl and amino group.

(iii) in the manufacture of cellulose acetate, aspirin, phenacetin, acetamide, acetophenone, etc.

Urea or Carbamide
$$O = \bigcirc \frac{NH_2}{NH_2}$$

Urea may be considered as diamide of an unstable and dibasic carbonic acid from which both the hydroxyl groups have been replaced by $-NH_2$ groups.



□ First time isolated from urine in 1773 by Roulle and hence the name urea was given.

□ It was the first organic compound synthesised in the laboratory from inorganic material (by heating a mixture of ammonium sulphate and potassium cyanate) by Wohler in 1828.

□ This preparation gave a death blow to Vital force theory.

□ It is the final decomposition product of protein's metabolism in man and mammals and is excreted along with urine.

Adults excrete about 30 grams of urea per day in the urine.

(1) Method of preparation

(i) *From urine* : Urine is treated with conc. nitric acid where crystals of urea nitrate $CO(NH_2)_2$. HNO₃ are obtained.

$$\begin{array}{c} 2\textit{CO}(\textit{NH}_2)_2.\textit{HNO}_3 + \textit{BaCO}_3 \rightarrow 2\textit{CO}(\textit{NH}_2)_2 + \textit{Ba}(\textit{NO}_3)_2 + \textit{H}_2\textit{O} + \textit{CO}_2 \\ \text{Urea nitrate} & \text{Urea} \end{array}$$

(ii) *Laboratory preparation*

(a) Wohler synthesis

$$\begin{array}{c} 2 \textit{KCNO} + (\textit{NH}_4)_2 \textit{SO}_4 \rightarrow 2 \textit{NH}_4 \textit{CNO} + \textit{K}_2 \textit{SO}_4 \\ \text{Potassium cyanate} & \text{Ammonium sulphate} \end{array} \rightarrow \begin{array}{c} 2 \textit{NH}_4 \textit{CNO} + \textit{K}_2 \textit{SO}_4 \\ \text{Ammonium cyanate} \end{array}$$

$$\begin{array}{c} NH_4 CNO & \xrightarrow{\text{Isomericchange}} NH_2 CONH_2 \\ \text{Ammonium cyanate} & On heating \\ \end{array}$$

☐ The solid residue is extracted with alcohol and the extract evaporated when the crystals of urea are obtained. It can be recrystalised from water.

(b) From phosgene or alkyl carbonate

$$O = C < \underbrace{CI}_{CI} + 2 NH_3 \rightarrow O = C < \underbrace{NH_2}_{NH_2} + 2HCI$$
Carbonyl chloride
(Phosgene)
Urea

$$O = C < \underbrace{\begin{array}{c} OC_2H_5\\ OC_2H_5 \end{array}}_{\text{Ctyle}} + 2NH_3 \rightarrow O = C < \underbrace{\begin{array}{c} NH_2\\ NH_2 \end{array}}_{\text{Urea}} + 2C_2H_5OH \\ \text{Urea} \end{aligned}$$

(iii) Industrial method

(a) By partial hydrolysis of calcium cyanide

$$\begin{array}{ccc} CaC_2 + N_2 & \xrightarrow{\text{heat}} & CaCN_2 + C\\ Calcium & Calcium\\ Carbide & cvanamide \end{array}$$

The cyanamide is treated with dilute sulphuric acid at $40^{\circ}C$ where partial hydrolysis occurs with the formation of urea.

$$CaCN_{2} \xrightarrow{H_{2}SO_{4}} H_{2}NCN \xrightarrow{H_{2}O} H_{2}NCONH_{2}$$

$$\xrightarrow{-CaSO_{4}} Cyanamide \xrightarrow{(H_{2}O_{2})} (Urea)$$

or

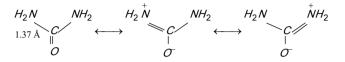
$$CaCN_2 + H_2O + H_2SO_4 \xrightarrow{40^{o}C} NH_2CONH_2 + CaSO_2$$

(b) From carbon dioxide and ammonia

$$CO_{2} + 2NH_{3} \xrightarrow{150-200^{\circ}C} NH_{2}COONH_{4}$$
Ammonium carbamate
$$\xrightarrow{\text{heat}(140^{\circ}C)} NH_{2}CONH_{2}$$
Urea
Urea

(2) **Physical properties** : Urea is a colourless, odourless crystalline solid. It melts at $132^{\circ}C$. It is very soluble in water, less soluble in alcohol but insoluble in ether, chloroform and benzene.

Crystal structure: In solid urea, both nitrogen atoms are identical.



This indicates that C - N bond in urea has some double bond character.

(3) Chemical Properties

(i) *Basic nature (Salt formation)*: It behaves as a weak monoacid base $(K_b = 1.5 \times 10^{-14})$. It forms solt with strong acid.

$$\begin{split} & \textit{NH}_2\textit{CONH}_2 + \textit{HNO}_3(\textit{conc}) \rightarrow \textit{NH}_2\textit{CONH}_2.\textit{HNO}_3 \\ & \text{Urea nitrate} \\ & 2\textit{NH}_2\textit{CONH}_2 + \textit{H}_2\textit{C}_2\textit{O}_4 \rightarrow (\textit{NH}_2\textit{CONH}_2)_2\textit{H}_2\textit{C}_2\textit{O}_4 \\ & \text{Oxalic acid} & \text{Urea oxalate} \end{split}$$

Urea is a stronger base than ordinary amide. It is due to the resonance stabilization of cation, the negatively charged oxygen atom is capable of coordination with one proton.

□ An aqueous solution of urea is neutral.

(ii) Hydrolysis

$$O = C \xrightarrow{[\overline{N}H_2 + H]} OH \xrightarrow{Aq. alkali or} O = C \xrightarrow{H} 2NH_3$$

$$\xrightarrow{[\overline{N}H_2 + H]} OH \xrightarrow{Aq. alkali or} O = C \xrightarrow{H} 2NH_3$$

Carbonic acid
$$CO_2 + H_2O$$

 $NH_2CONH_2 + 2NaOH \rightarrow 2NH_3 + Na_2CO_3$

An enzyme, urease, present in soyabean and soil also brings hydrolysis .

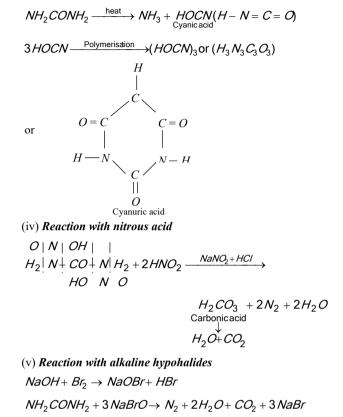
$$\begin{array}{c} \textit{NH}_2\textit{CONH}_2 + \textit{2H}_2\textit{O} \rightarrow (\textit{NH}_4)_2\textit{CO}_3 \rightarrow \textit{2NH}_3 + \textit{CO}_2 + \textit{H}_2\textit{O}_3 \\ \text{Ammonium carbonate} \end{array}$$

(iii) Action of heat

$$\begin{array}{c} \textit{NH}_2\textit{CO}[\overline{\textit{NH}_2} + H] \\ (\texttt{Two molecules} \texttt{f urea}) \end{array} \xrightarrow{\text{heat}} \textit{NH}_2\textit{CONHCONH}_2 + \textit{NH}_3 \\ \overrightarrow{\text{Biuret}} \end{array}$$

Urea is identified by the test known as biuret test. The biuret residue is dissolved in water and made alkaline with a few drops of NaOH. When a drop of copper sulphate solution is added to the alkaline solution of biuret, a violet colouration is produced.

when heated rapidly at $170^{\circ}C$, polymerisation takes place:



 $NH_2CONH_2 + (CH_3CO)_2O \rightarrow NH_2CONHCOCH_3$ Acetyl urea

> + CH₃COOH Aceticacid

(vii) Reaction with hydrazine

$$\underset{\text{Urea}}{\text{NH}_2CONH}_2 + \underset{\text{Hydrazine}}{H_2N.NH_2} \xrightarrow{100^{\circ}C} \text{NH}_2CONH.NH_2 + \text{NH}_3$$

Semicarbażde

(viii) Reaction with ethanol

$$H_2 NCQ \boxed{NH_2} + H_1 OC_2 H_5 \xrightarrow{\text{heat}} H_2 NCOOC_2 H_5 + NH_3$$

Ethanol Urethane

(ix) Reaction with chlorine water

$$O = O + 2Cl_2 \rightarrow O = O + 2HCl$$

$$VH_2 \qquad VHCl$$
Urea Dichloro urea

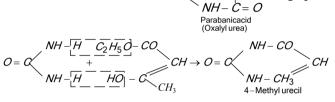
(x) **Dehydration**

$$NH_2CONH_2 + SOCI_2 \rightarrow H_2N - C \equiv N + SO_2 + 2HCI + H_2O$$

$$NH_2CONH_2 + \underbrace{H_2SO_4 + SO_3}_{Oleum} \rightarrow \underbrace{2NH_2SO_3H_+ CO_2}_{sulphamicacid}$$

(xii) Formation of cyclic ureides

$$\langle \prod_{i=1}^{n} \prod_{i=1}^{n} \prod_{j=1}^{n} \prod_{i=1}^{n} \prod_{j=1}^{n} \prod_$$



(xiii) Reaction with formaldehyde

 $\begin{array}{c} CH_2 = O + NH_2CONH_2 \xrightarrow{HCI} CH_2(OH)NHCONH_2 \xrightarrow{CH_2=O} \\ Formaldehyde & Monomethylol urea \\ \hline CH_2(OH)NHCONH(OH)CH_2 \xrightarrow{heat} Resin \\ Dimethylol urea & (Urea-Formaldehyde) \end{array}$

(4) Uses

(i) Mainly as a nitrogen fertilizer. It has 46.4% nitrogen.

(ii) In the manufacture of formaldehyde-urea plastic and semicarbazide.

(iii) As animal feed.

(iv) For making barbiturates and other drugs.

(v) As a stabilizer for nitrocellulose explosives.

(5) General Tests

(i) When heated with sodium hydroxide, ammonia is evolved.

(ii) When heated gently, it forms biuret which gives violet colouration with sodium hydroxide and a drop of copper sulphate solution.

(iii) Its aqueous solution with concentrated nitric acid gives a white precipitate.

(iv) On adding sodium nitrite solution and dil. HCI (*i.e.*, HNO_2) to urea solution, nitrogen gas is evolved and gives effervescence due to carbon dioxide.



 \mathcal{K} Oxidation of 1° alcohols and aldehyde cannot be carried out with alkaline $K_2Cr_2O_7$ since under these condition K_2CrO_4 is formed which does not act as an oxidising agent.

 \mathcal{L} During oxidation of alkyl benzenes with alkaline $KMnO_4$ or acidified $K_2Cr_2O_7$, the aromatic nucleus remain intact but each side chain oxidised to -COOH group irrespective of its length. The ease of oxidation of alkyl benzenes follows the order.

Toluene > Ethyl benzene > Isopropyl benzene.

Tert butyl benzene however does not undergo oxidation to give benzoic acid since it does not have any benzylic *H*-atom.

Carboxylic acids are stronger acids than phenols since carboxylate ion is better stabilized by resonance than phenoxide ion.

E The melting points of carboxylic acids show oscillation or alternation effect, *i.e.* melting point of an acid containing even number of carbon atom is higher than the next lower or higher homologue containing odd number of carbon atom, due to greater symmetry and close packing of molecules in the crystal lattice.

Carboxylic acid do not give the characteristic reaction of the carbonyl group. The reason being that due to resonance, the double bond character of the C = O bond in carboxylic acid is greatly reduced as compared to that in aldehyde and ketone.

✗ The boiling point of acid derivatives follow the order.

 $RCONH_2 > (RCO)_2O > RCOOH > RCOOR' > RCOCl.$

 \mathcal{E} The boiling point of acid chloride and ester are lower than those of their parent acid due to absence of *H*-bonding in their molecule.

 \mathcal{L} The boiling points of acid anhydrides are higher than those of the acids from which they are derived because of stronger Vander Waal's forces of attraction owing to the larger size of their molecules.

 \mathcal{K} The melting points and boiling points of acid amide are much higher than those of the acids from which they are derived due to strong intermolecular *H*-bonding even though their molecular masses are almost identical.

 \mathcal{K} Aromatic acid chlorides are less reactive than aliphatic acid chlorides primarily due to greater electron donating effect of the benzene ring over alkyl group which tends to reduce the electron deficiency of aromatic acyl carbon.

 \checkmark Phthalimide and succinimide on treatment with Br_2 -KOH undergo Hofmann bromamide reaction to form anthranilic acid and β -aminopropionic acid respectively.

✗ Urea acts as a monoacidic base.

\not{\mathbb{Z}} Malonic acid on heating with P_2O_5 gives carbon suboxide (C_3O_2) .

Z Tamarind contain tartaric acid which does not exist in nature.

E Baking powder is a mixture of sodium bicarbonate and cream of tartar *i.e.* acid potassium hydrogen sulphate.

Smell of ammonia in public urinals is due to hydrolysis of urea present in urine by the enzyme urease present in atmosphere.

 \bigstar Tartar emetic (*i.e.* potassium antimony D(+) tartrate is used to cause nausea and vomiting during treatment of poisoning.

✗ Magnesium citrate is used as an antacid.

- Succinic acid was prepared by the distillation of amber.
- Malic acid is found in apples, grapes etc.